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(54) **EFFICIENT REVERSIBLE ELECTRODES FOR SOLID OXIDE ELECTROLYZER CELLS**

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(52) **U.S. Cl.** **204/252**; 204/263; 204/266; 204/291

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See application file for complete search history.

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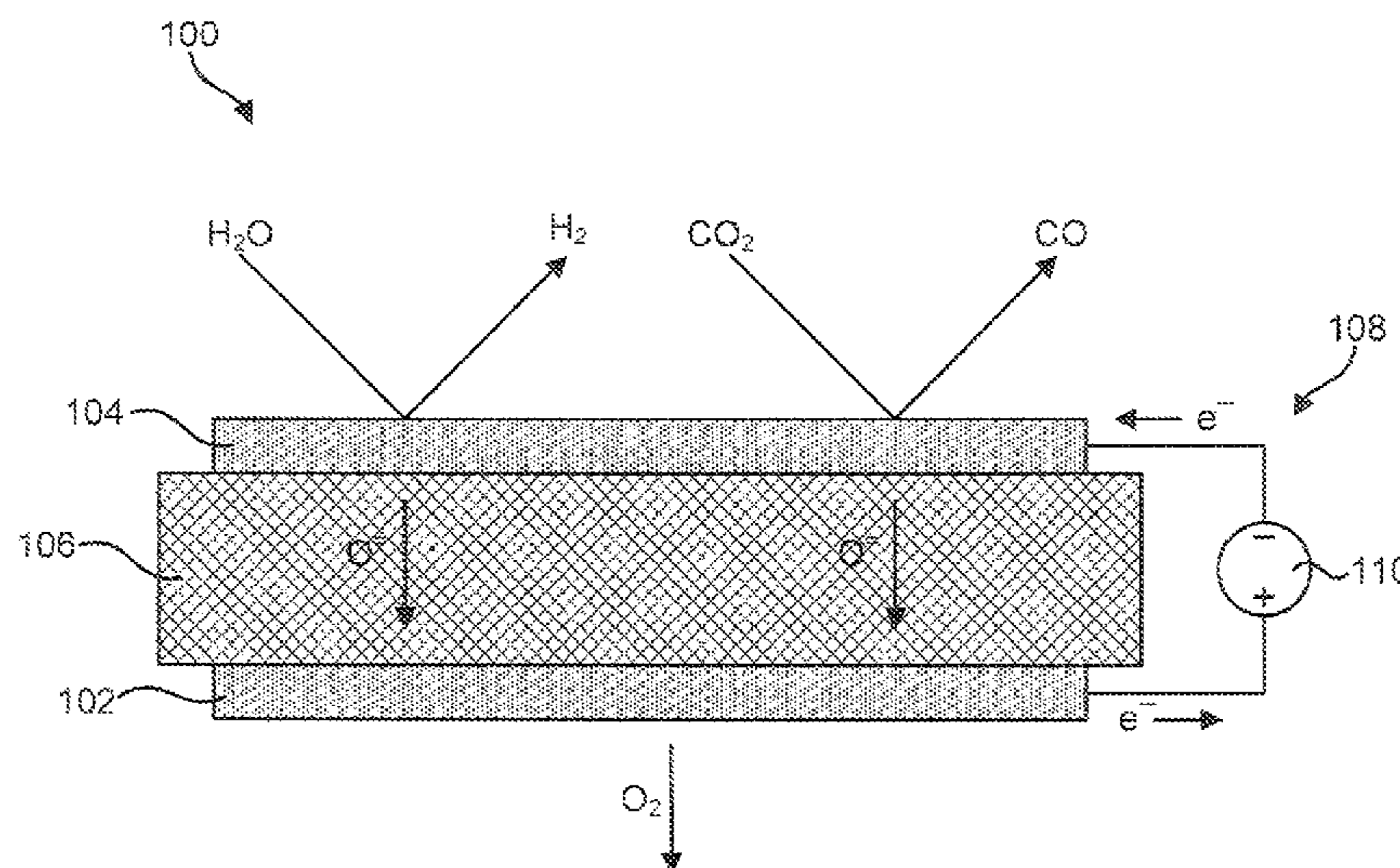
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(57) **ABSTRACT**

An electrolyzer cell is disclosed which includes a cathode to reduce an oxygen-containing molecule, such as H₂O, CO₂, or a combination thereof, to produce an oxygen ion and a fuel molecule, such as H₂, CO, or a combination thereof. An electrolyte is coupled to the cathode to transport the oxygen ion to an anode. The anode is coupled to the electrolyte to receive the oxygen ion and produce oxygen gas therewith. In one embodiment, the anode may be fabricated to include an electron-conducting phase having a perovskite crystalline structure or structure similar thereto. This perovskite may have a chemical formula of substantially (Pr(1-x)Lax)(z-y)A'yBO(3-δ), wherein 0 ≤ x ≤ 0.5, 0 ≤ y ≤ 0.5, and 0.8 ≤ z ≤ 1.1. In another embodiment, the cathode includes an electron-conducting phase that contains nickel oxide intermixed with magnesium oxide.

11 Claims, 3 Drawing Sheets



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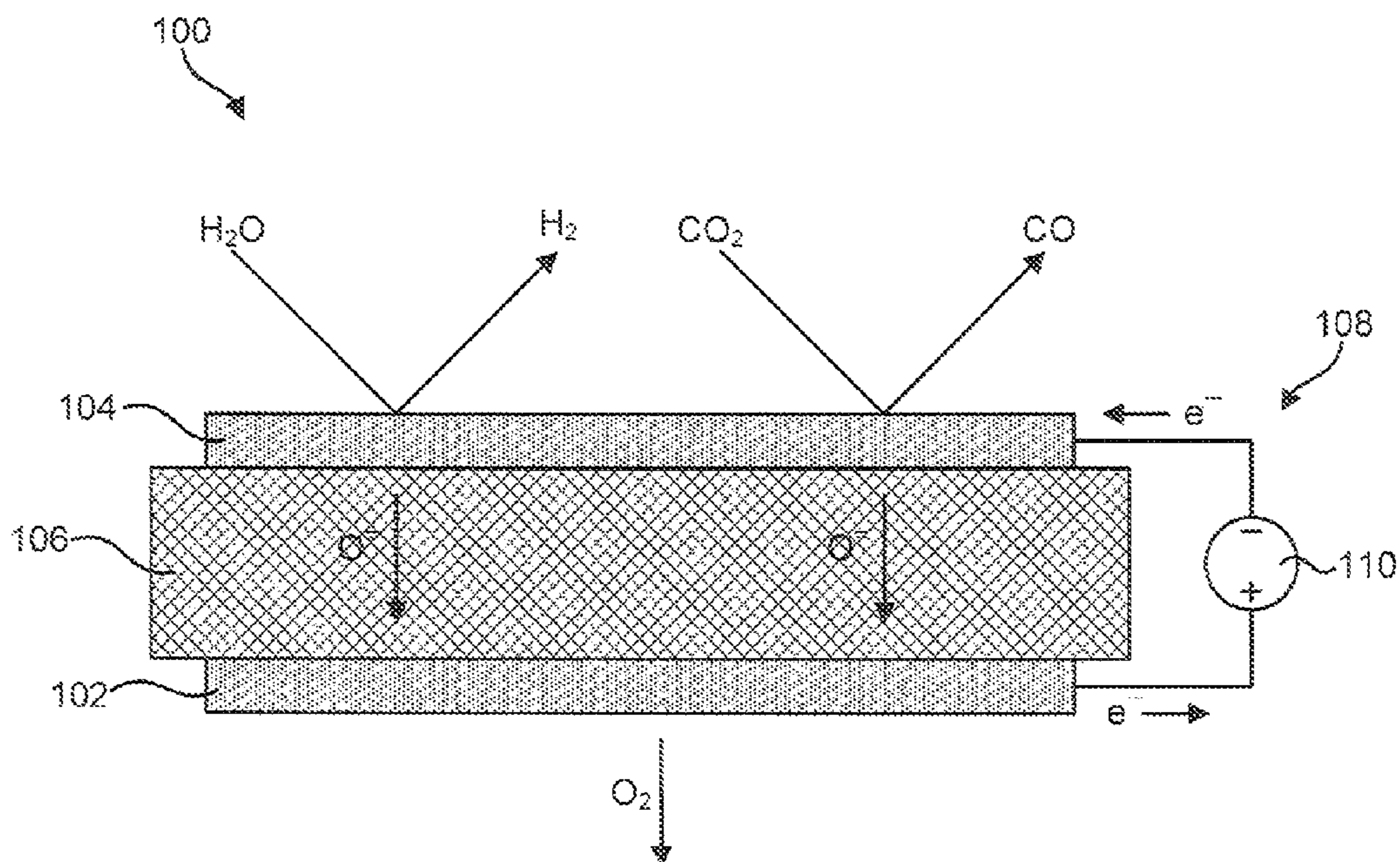


Fig. 1

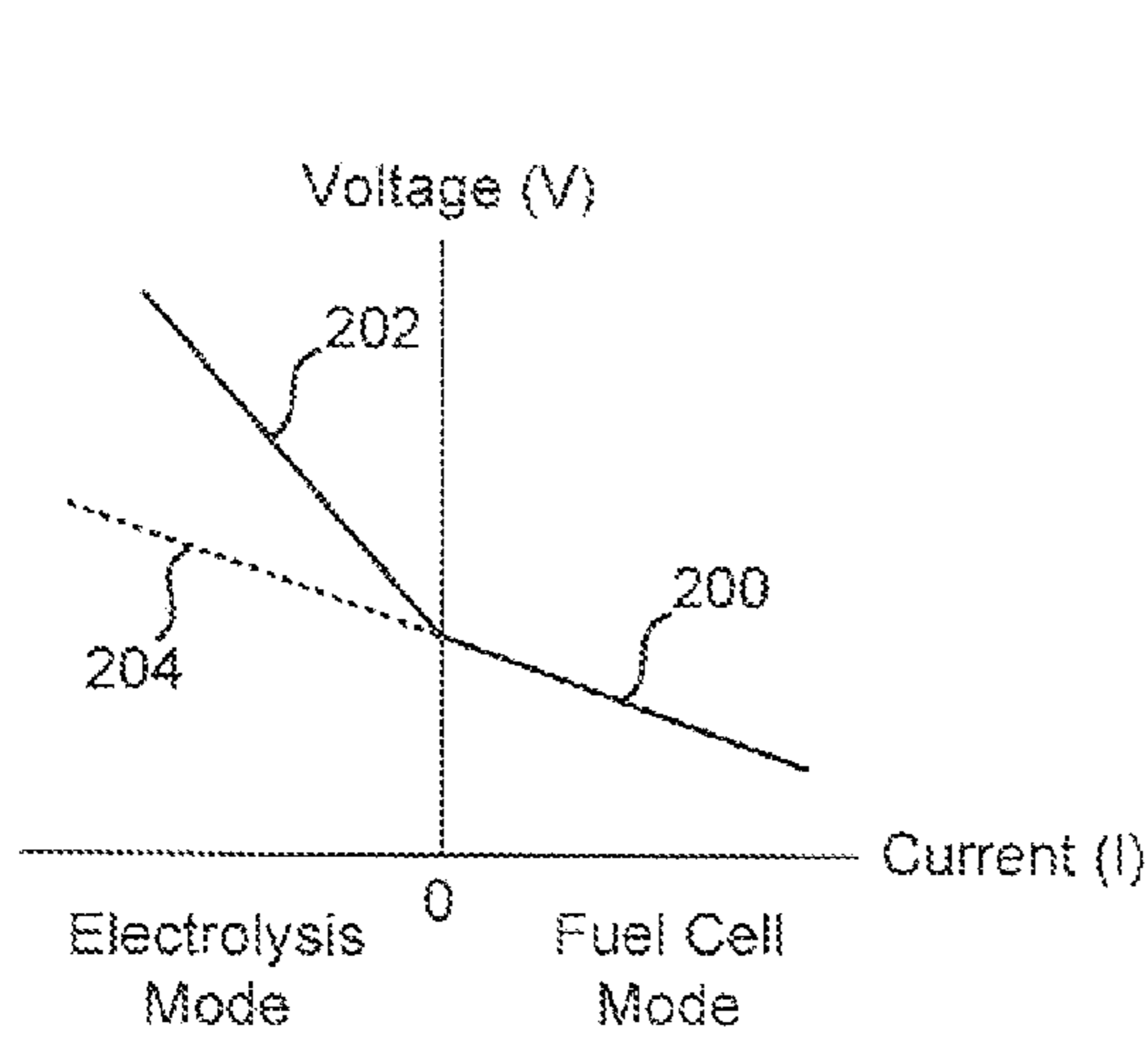


Fig. 2A

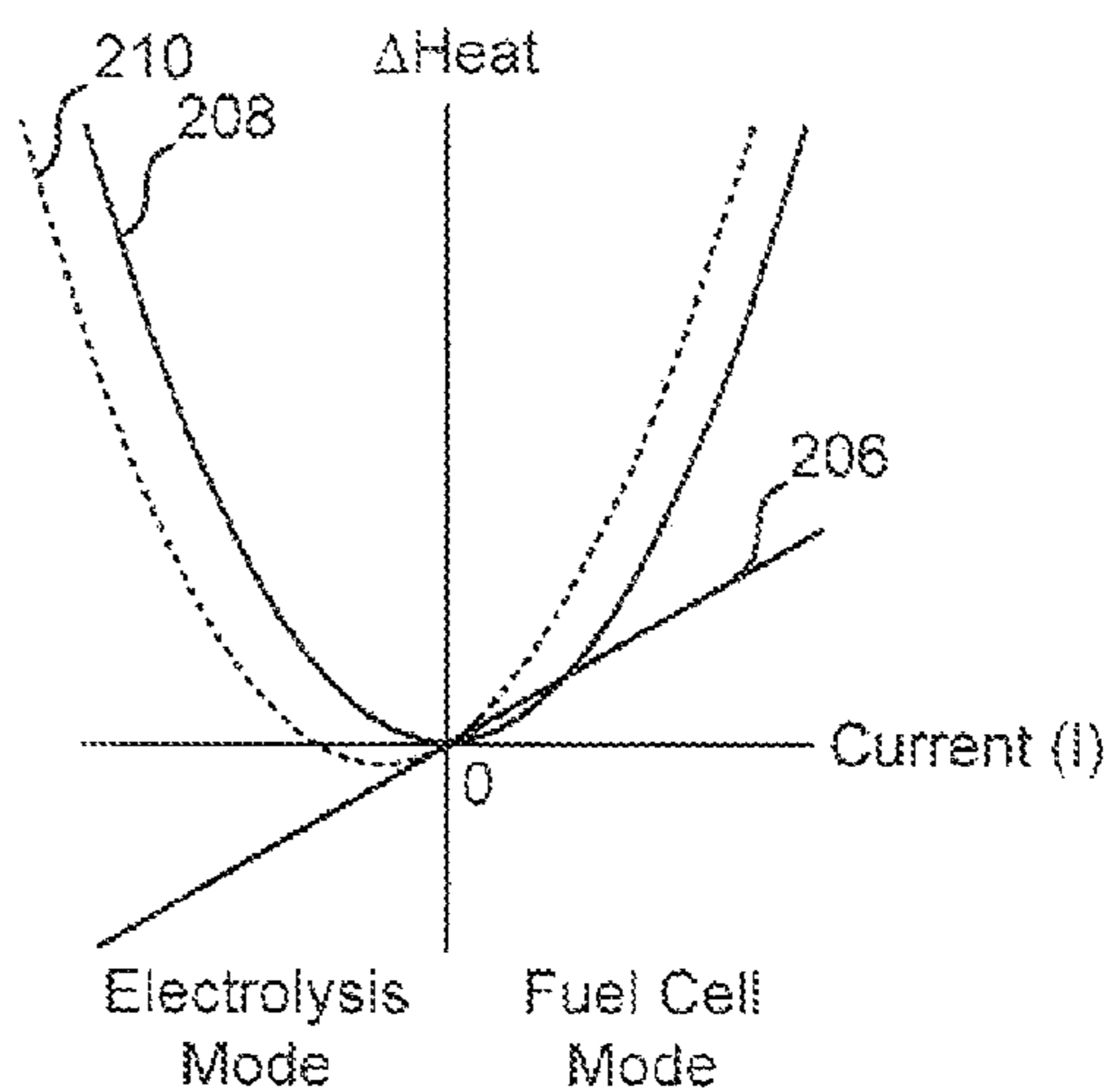


Fig. 2B

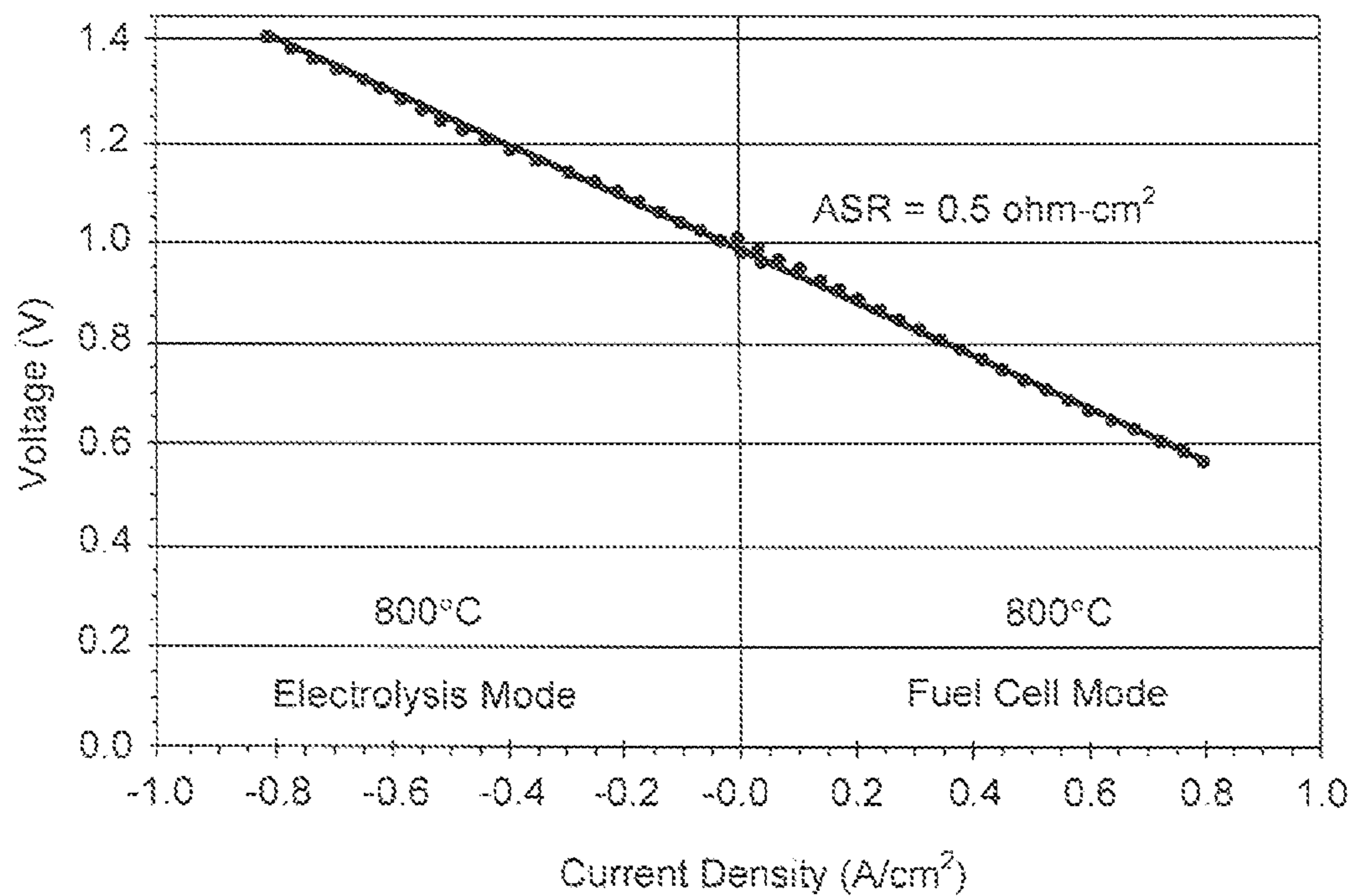


Fig. 3

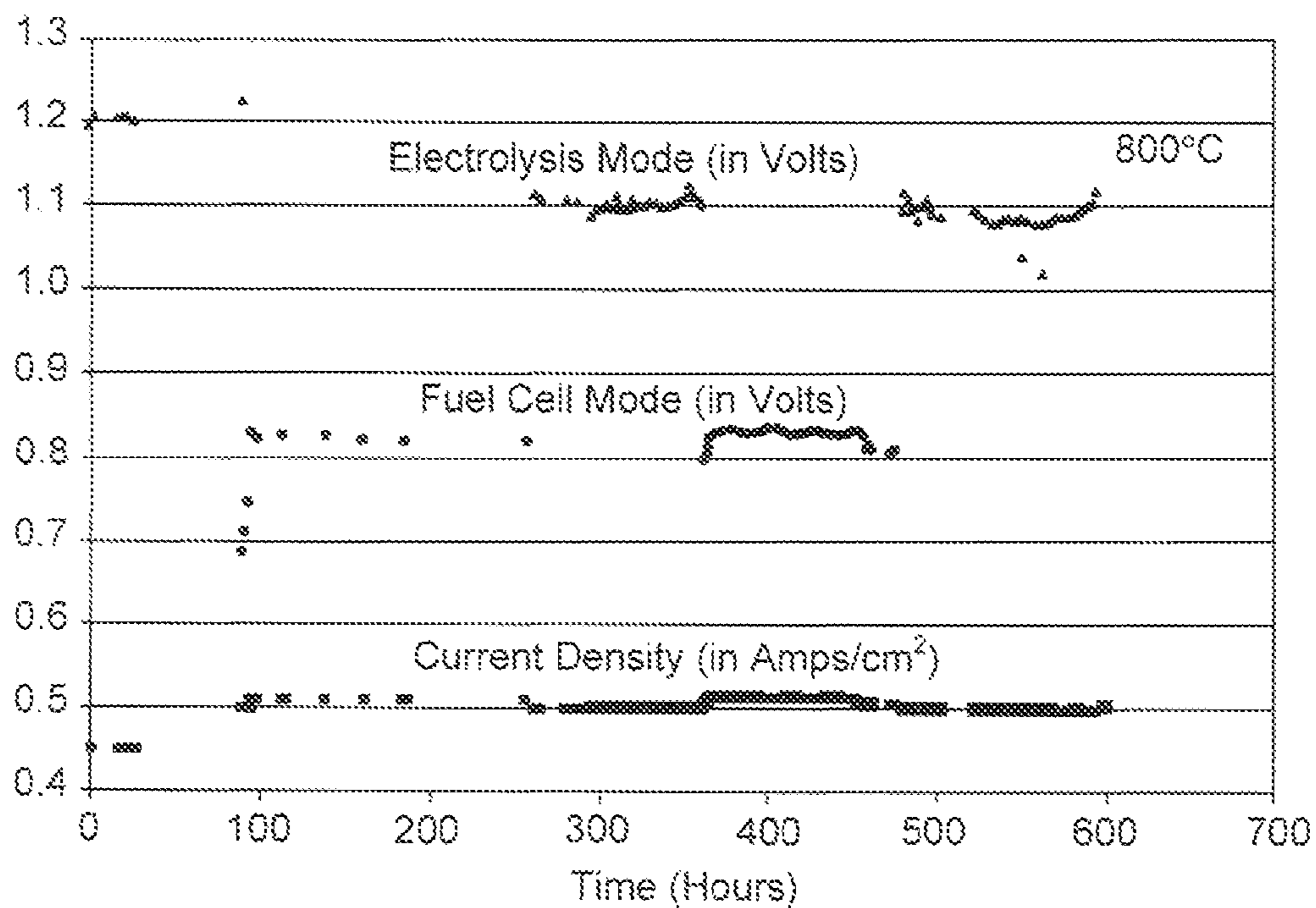


Fig. 4

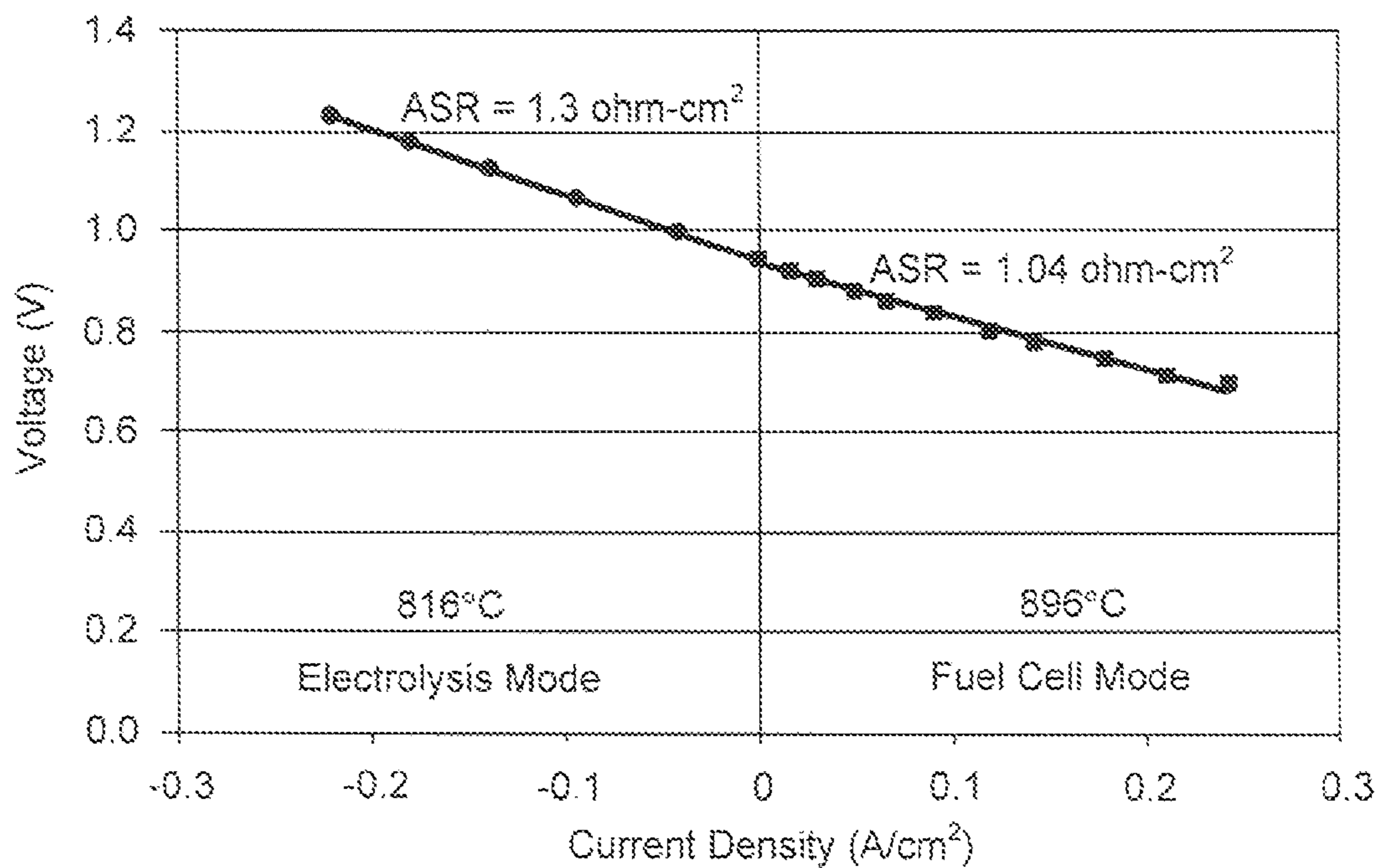


Fig. 5

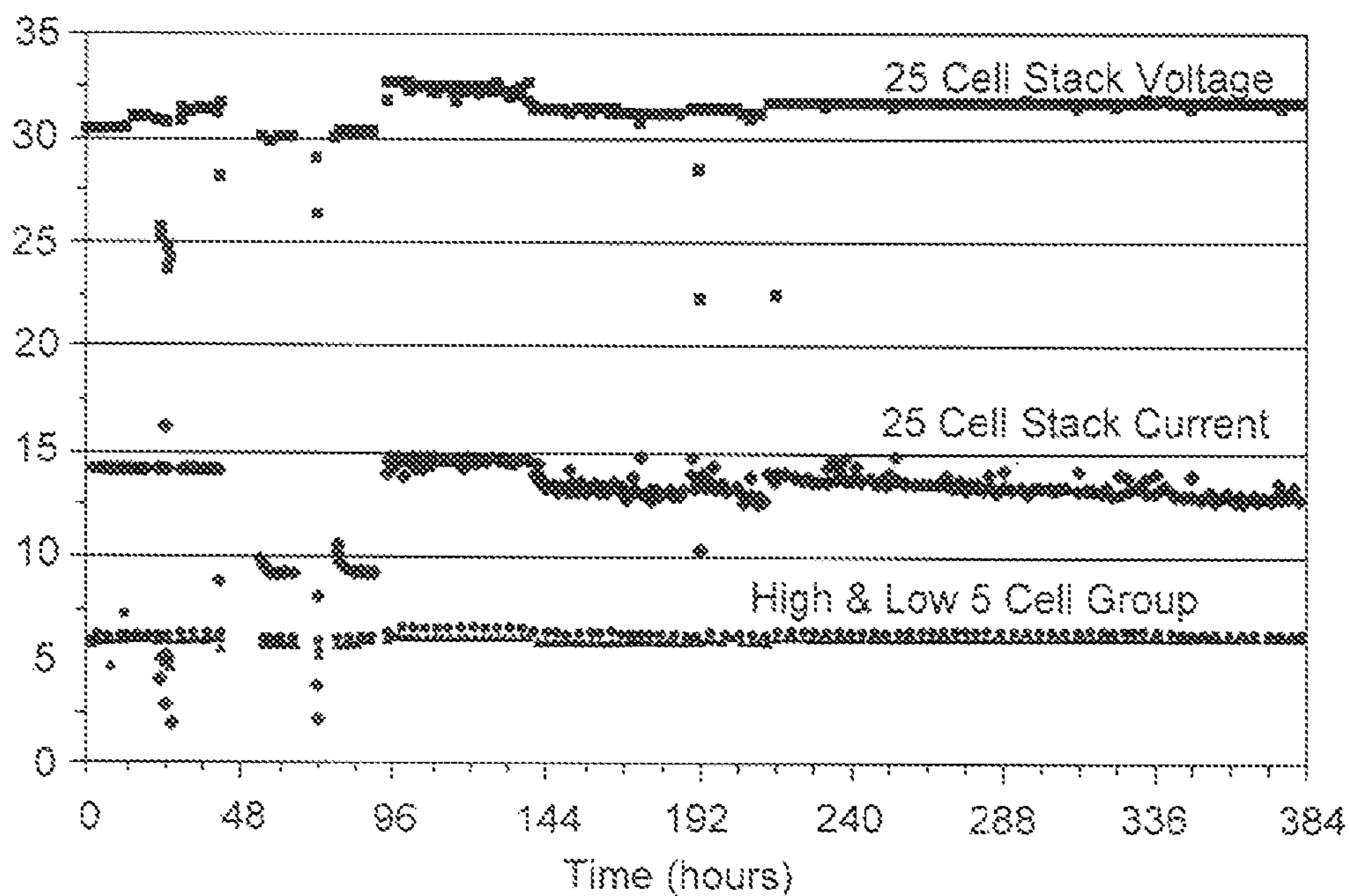


Fig. 6

EFFICIENT REVERSIBLE ELECTRODES FOR SOLID OXIDE ELECTROLYZER CELLS

RELATED APPLICATIONS

This application claims priority to U.S. Provisional Patent No. 60/820,103 filed on Jul. 22, 2006 and entitled ELECTRODES FOR SOLID OXIDE ELECTROLYZER CELLS.

GOVERNMENT RIGHTS

This invention was made in part with government support under grant number DE-AC07-05ID14517 awarded by the United States Department of Energy. The Government has certain rights in the invention.

BACKGROUND OF THE INVENTION

1. Field of the Invention

This invention relates to electrolyzer cells and more particularly to reversible electrodes for solid oxide electrolyzer cells.

2. Description of the Related Art

Fuel cells are expected to play an important role in ensuring our energy security and may be an important component to establishing a "hydrogen economy." Recent efforts have been dedicated to developing and implementing a commercially viable hydrogen fuel cell technology to power cars, trucks, homes, businesses, and the like, in order to reduce dependence on foreign sources of oil. Fuel cells also have the potential to reduce or eliminate harmful emissions generated by conventional power sources such as internal combustion engines.

Unlike many hydrocarbons, molecular hydrogen is not available in centralized natural reservoirs. Moreover, the present infrastructure for production, storage, and delivery of hydrogen is currently vastly inadequate to support a hydrogen economy. Transition to hydrogen use would require altering current industrial and transportation practices on an enormous scale.

Due to the expense and problems associated with hydrogen storage and transportation, some have advocated a more distributed approach to hydrogen production. For example, small regional plants, local filling stations, or even home-based generators could be used to produce hydrogen using energy provided through the electrical distribution grid. While this may lower generation efficiency compared to a centralized generation approach, it may increase overall efficiency when considering all the costs and energy requirements needed to deliver hydrogen to the end user.

Currently, a "reversible" fuel cell offers one potential solution for generating both electricity and hydrogen (or synthesis gas) using a single device. A reversible fuel cell may be used to generate electricity, when operated in fuel cell mode, and hydrogen (or synthesis gas) when operated in electrolysis mode. A reversible fuel cell can be used to produce electricity as needed but may also utilize excess capacity of the electrical grid during off-peak hours to produce hydrogen fuel. This fuel may be used at a later time during periods of high electrical demand or to power a vehicle or other device. A reversible fuel cell also has the potential to reduce costs significantly by converting electricity to hydrogen and hydrogen to electricity using a single device. Nevertheless, to achieve commercial success, the reversible cell must produce hydrogen with enough efficiency to be competitive with other means of production.

While various studies have shown that reversible fuel cells are feasible, it has also been shown that electrodes that perform well in fuel cell mode typically do not perform as well in electrolysis mode, and vice versa. For example, one study demonstrated that a solid oxide fuel cell (SOFC) comprising a negative electrode containing Ni and YSZ and a positive electrode containing LSM significantly outperformed a solid oxide electrolyzer cell (SOEC) using the same electrode materials. In contrast, an SOEC with a platinum negative electrode and an LSCo-containing positive electrode showed lower polarization losses than an SOFC using the same electrode materials. Yet another study demonstrated that, using typical SOFC materials, polarization losses during electrolysis operation were greater than those for fuel cell operation for both electrodes. This study also demonstrated that the polarization increase at reduced temperature was greater in electrolysis mode than in fuel cell mode.

In view of the foregoing, what are needed are electrodes that perform equally well in both fuel cell and electrolysis modes. Such electrodes may be used to provide a reversible fuel/electrolyzer cell which is efficient in either mode of operation. Ideally, the electrodes would exhibit similar polarization and other characteristics in both fuel cell and electrolysis modes of operation.

SUMMARY OF THE INVENTION

The present invention has been developed in response to the present state of the art, and in particular, in response to the problems and needs in the art that have not yet been fully solved by currently available electrolyzer electrodes.

Consistent with the foregoing and in accordance with the invention as embodied and broadly described herein, an electrolyzer cell is disclosed in one embodiment of the invention as including a cathode to reduce an oxygen-containing molecule, such as H₂O, CO₂, or a combination thereof, to produce an oxygen ion and a fuel molecule, such as H₂, CO, or a combination thereof. An electrolyte is coupled to the cathode to transport the oxygen ion to an anode. The anode is coupled to the electrolyte to receive the oxygen ion and produce oxygen gas therewith. The anode may be fabricated to include an electron-conducting phase having a perovskite crystalline structure or structure similar thereto. This perovskite may have a chemical formula of substantially (Pr_(1-x)La_x)_(z-y)A'_yBO_(3-ε), wherein 0 ≤ x ≤ 0.5, 0 ≤ y ≤ 0.5, and 0.8 ≤ z ≤ 1.1.

In selected embodiments, the anode further includes an ion-conducting phase intermingled with the electron-conducting phase. The ion-conducting phase may include one or more oxides. For example, in selected embodiments, the oxide may include zirconia doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, cerium oxide, or the like. In other embodiments, the oxide may include ceria doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, zirconium oxide, or the like.

In certain embodiments, the anode may be designed with a selected porosity to allow gas to pass therethrough. To further improve the electrochemical efficiency, the pores of the anode may be infiltrated with an electrocatalyst, such as praseodymium, cobalt, cerium, europium, other rare earth elements, and combinations thereof.

In the chemical equation above, "A" may include one or more alkaline earth metal, such as calcium, strontium, or the like. Similarly, "B" may include one or more transition metals, such as manganese, cobalt, iron, or the like.

In another aspect of the invention, an electrolyzer cell in accordance with the invention may include a cathode to

reduce an oxygen-containing molecule, such as H₂O, CO₂, or combinations thereof, to produce an oxygen ion and a fuel molecule, such as H₂, CO, or combinations thereof. The cathode may include an electron-conducting phase that contains nickel oxide intermixed with magnesium oxide. This nickel oxide may be reduced to nickel upon operating the electrolyzer cell. The magnesium oxide reduces coarsening of the nickel oxide to maintain the nickel's surface area and to maintain conductivity with adjacent particles. An electrolyte is coupled to the cathode to transport the oxygen ion from the cathode. An anode is coupled to the electrolyte to receive the oxygen ion and produce oxygen gas therewith.

In selected embodiments, the molar ratio of nickel oxide to magnesium oxide is greater than or equal to three. The cathode may further include an ion-conducting phase intermixed with the electron-conducting phase. This ion-conducting phase may include one or more oxides. For example, the oxide may include zirconia doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, or the like. In other embodiments, the oxide may include ceria doped with one or more of samarium oxide, gadolinium oxide, yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, or the like.

In certain embodiments, the cathode may be designed to have a selected porosity to allow gas to pass therethrough. To improve the electrochemical efficiency, the pores of the cathode may be infiltrated with an electrocatalyst, such as praseodymium, cobalt, cerium, europium, other rare earth elements, and combinations thereof.

The present invention relates to improved electrodes for reversible fuel/electrolyzer cells. The features and advantages of the present invention will become more fully apparent from the following description and appended claims, or may be learned by practice of the invention as set forth hereinafter.

BRIEF DESCRIPTION OF THE DRAWINGS

In order that the advantages of the invention will be readily understood, a more particular description of the invention briefly described above will be rendered by reference to specific embodiments illustrated in the appended drawings. Understanding that these drawings depict only typical embodiments of the invention and are not therefore to be considered limiting of its scope, the invention will be described and explained with additional specificity and detail through use of the accompanying drawings in which:

FIG. 1 is a high-level block diagram of one embodiment of an electrolyzer cell in accordance with the invention;

FIG. 2A is a graph showing desired voltage/current characteristics of a cell when operated in fuel cell and electrolysis modes;

FIG. 2B is a graph showing the heat generated by a cell when operated in fuel cell and electrolysis modes;

FIG. 3 is a graph showing the current/voltage characteristics of a single button cell in accordance with the invention, when operated in fuel cell and electrolysis modes;

FIG. 4 is a graph showing the long term performance of a single button cell when operated in fuel cell and electrolysis modes;

FIG. 5 is a graph showing the current/voltage characteristics of a stack of twenty-five cells when operated in fuel cell and electrolysis modes; and

FIG. 6 is a graph showing the long term performance of a stack of twenty-five cells when operated in electrolysis mode.

DETAILED DESCRIPTION OF THE INVENTION

It will be readily understood that the components of the present invention, as generally described and illustrated in the

Figures herein, could be arranged and designed in a wide variety of different configurations. Thus, the following more detailed description of the embodiments of the invention, as represented in the Figures, is not intended to limit the scope of the invention, as claimed, but is merely representative of certain examples of presently contemplated embodiments in accordance with the invention. The presently described embodiments will be best understood by reference to the drawings, wherein like parts may be designated by like numerals throughout.

Referring to FIG. 1, in general, a solid oxide electrolyzer cell 100 (which may also, in selected embodiments, function as a fuel cell 100 when the current direction is reversed) may include an anode 102, a cathode 104, and an electrolyte layer 106. Each of the layers 102, 104, 106 may, in selected embodiments, be composed of solid-state ceramic materials.

In general, the anode 102 and cathode 104 may be composed of a material that is both electronically and ionically conductive. The electrolyte, on the other hand, may be composed of a material that is ionically conductive but electronically insulating. Thus, electrons flowing between the anode 102 and cathode 104 may need to flow through an external path 108 or circuit 108 in order to bypass the electrolyte 106. The electrolyte layer 106 may be composed of one or more known ionically-conductive but electrically-insulating ceramic materials such as, for example, yttria stabilized zirconia (YSZ), scandium stabilized zirconia (ScSZ), or lanthanum strontium gallate magnesite (LS GM).

When operating in electrolysis mode, oxygen-containing molecules, such as H₂O or CO₂, may be received at the cathode 104, where they may react to generate a fuel, such as H₂, CO, or a mixture thereof. Electrons flowing from an external power source 110 may be consumed by this reaction. Similarly, negatively charged oxygen ions may be liberated by this reaction. These oxygen ions may be conducted through the electrolyte layer 106 to the anode 102, where they may react to form oxygen molecules. The reaction at the anode 102 may release electrons, which may in turn flow through the pathway 108 or circuit 108.

As mentioned, electrodes that perform well in fuel cell mode typically do not perform as well in electrolysis mode, and vice versa. For example, electrodes used in many fuel cells exhibit unacceptably high polarization losses when operated in electrolysis mode. This reduces the desirability of using the cell in electrolysis mode because it is inefficient and may not be competitive with other means of producing hydrogen (or carbon monoxide). Thus, it would be desirable to provide electrodes that perform equally well in both fuel cell and electrolysis modes.

For example, as illustrated in FIG. 2A, in fuel cell mode, the voltage across a typical fuel cell will drop as current conducted through the fuel cell increases. This may be observed from the downward slope of the trace 200, which is an indicator of the cell's internal resistance, or polarization. When operated in electrolysis mode, however, a disproportionately larger voltage is needed to produce a roughly equal increase in current through the cell. This may be observed from the steeper slope of the trace 202, which is indicative of the higher resistance, or polarization, of many conventional fuel cells when operated in electrolysis mode. Ideally, the resistance of the cell (and thus the slope of the traces 200, 202) is minimized in both fuel cell and electrolysis modes, with the resistance in electrolysis mode ideally being roughly equivalent to the resistance in fuel cell mode. Thus, the slope of the trace 202 in electrolysis mode will ideally be roughly equal to

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the slope of the trace **200** in fuel cell mode, as indicated by the dotted line **204**, with each slope being minimized as much as possible.

Referring to FIG. **2B**, determining whether electrodes **102**, **104** behave in a manner consistent with the ideal slope **204** illustrated in FIG. **2A** may be challenging because the resistance of a cell may vary with temperature. This problem may be compounded by the fact that conversion of H₂ and/or CO to H₂O and/or CO₂ (fuel cell mode) is exothermic, whereas conversion of H₂O and/or CO₂ to H₂ and/or CO (electrolysis mode) is endothermic. Thus, the fuel cell reaction generates heat whereas the electrolysis reaction absorbs heat, as represented by the linear curve **206**.

On the other hand, heat generated by resistive losses (i.e., I²R) due to the flow of current may be roughly equal for current flowing in either direction through the cell, as represented by the curve **208**. The two curves **206**, **208** may be added together to provide a resultant curve **210**. As shown by the resultant curve **210**, a cell operating in electrolysis mode will initially cool but then begin to heat up once the heat from resistive losses is greater than the heat absorbed by the reaction. This concept becomes important when interpreting the test data illustrated in FIG. **5**.

Referring again to FIG. **1**, in selected embodiments, an electrolyzer cell **100** that operates with substantially equal efficiency in both fuel cell mode and electrolyzer mode may include an anode **102** containing an electron-conducting phase intermingled with an ion-conducting phase. The electron-conducting phase may include a perovskite generally represented by the chemical formula ABO₃, wherein "A" and "B" represent A- and B-side cations. The A-side cations may be represented generally by the formula (Pr_(1-x)La_x)_(z-y)A'_y, wherein 0 ≤ x ≤ 0.5, 0 ≤ y ≤ 0.5, and 0.8 ≤ z ≤ 1.1, and A' represents one or more alkaline earth metals, such as calcium and strontium. The alkaline earth metals may create defects in the crystalline structure of the perovskite to improve its electronic conductivity.

The B-side cations may include one or more transition metals, such as manganese, cobalt, iron, or combinations thereof. The ∅ may refer to the oxygen non-stoichiometry which may depend on the crystalline chemistry and electro-neutrality conditions. Suitable materials for the electron-conducting phase may be represented generally by the formula (Pr_(1-x)La_x)_(z-y)A'_yBO_(3-∅) and may include, for example, (Pr_(1-x)La_x)_(z-y)Sr_yMnO_(3-∅), (Pr_(1-x)La_x)_(z-y)Sr_yFeO_(3-∅), and (Pr_(1-x)La_x)_(z-y)Sr_yCoO_(3-∅).

The ion-conducting phase may include one or more oxides. For example, the ion-conducting phase may include zirconia doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, cerium oxide, or the like. In other embodiments, the ion-conducting phase may include ceria by itself or doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, zirconium oxide, or the like. One advantage of ceria is that it is both electronically and ionically conductive, with much better ionic conductivity than zirconia. The enhanced ionic conductivity is at least partly due to the fact that oxygen is less firmly attached to the ceria molecule than it is with zirconia, providing improved oxygen conductivity and better electrochemical performance.

As shown by the chemical formula above, the molar concentration of praseodymium (Pr) may equal, exceed, or completely replace lanthanum (La), a common material in conventional electrodes. Lanthanum and zirconia are known to react to form the insulating product lanthanum zirconate, which may accumulate between the ion and electron-conducting phases and increase the resistance of the electrode

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over time. To reduce this problem, praseodymium, an element which does not react with zirconia, may be used to replace all or part of the lanthanum.

In selected embodiments, the anode **102** may also be infiltrated with an electrocatalyst such as praseodymium, cobalt, cerium, europium, or other rare earth elements or combinations thereof. The anode **102** may be fabricated to have a desired porosity, such as, for example, between about ten and twenty percent porosity by volume. These pores may be infiltrated with the electrocatalyst to increase the catalytic activity of the anode **102**.

In certain embodiments, the anode **102** may be infiltrated with the electrocatalyst by saturating the anode **102** with an electrocatalyst-containing solution. For example, nanoparticles of praseodymium cobalt nitrate suspended in solution may be used to saturate the pores of the anode **102**. The nitrate may then be burned off to leave the catalyst finely deposited on the porous surface of the anode **102**.

In selected embodiments, the cathode **104** of the electrolyzer cell **100** may also include an electron-conducting phase intermingled with an ion-conducting phase. The electron-conducting phase may, in selected embodiments, include a solid solution of nickel oxide and magnesium oxide. In selected embodiments, the molar ratio of nickel oxide to magnesium oxide is greater than or equal to three.

Upon operating the electrolyzer cell **100**, the nickel oxide may be reduced to nickel in the presence of a reducing gas, such as hydrogen, to provide the electronically conducting phase. The magnesium oxide, on the other hand, will remain in oxide form because it is a more stable oxide. This magnesium oxide will remain finely dispersed through the nickel.

The magnesium oxide improves the performance of the cathode **104** by reducing the tendency of the nickel to coarsen over time as it is subjected to high temperatures. Otherwise, very fine particles of nickel will tend to sinter and grow together when subjected to high temperatures. This reduces electrochemical performance at least partly because it reduces the surface area of the nickel. In some cases, the courser nickel particles may pull apart from adjacent particles and lose conductivity therewith. The distribution of magnesium oxide particles through the nickel tends to reduce or at least slow down the coarsening of the nickel, thereby improving electrochemical performance.

The ion-conducting phase of the cathode **104** may also include one or more oxides, such as oxides of zirconia, ceria, or mixtures thereof. For example, the ion-conducting phase may include zirconia doped with one or more of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, or the like. In other embodiments, the ion-conducting phase may include ceria doped with one or more of samarium oxide, gadolinium oxide, yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, or the like. As mentioned above, ceria may provide better electronic and ionic conductivity than zirconia. The enhanced ionic conductivity and oxygen availability of ceria may be used to further improve the electrochemical performance.

Like the anode **102**, the cathode **104** may also be infiltrated with an electrocatalyst such as praseodymium, cobalt, cerium, europium, other rare earth elements, or combinations thereof. This electrocatalyst provides additional activation energy to break the bonds of H₂O and CO₂ molecules supplied to the cathode **104**, further improving electrochemical performance.

One notable characteristic of the anode and cathode materials described above is that they exhibit significantly improved performance when electrolyzing CO₂ to generate

CO. With many conventional electrode materials, it has been observed that the kinetics of the CO₂ to CO electrolysis reaction are far worse than the kinetics of the H₂O to H₂ reaction. Stated otherwise, conventional electrodes exhibit far greater inherent resistance when converting CO₂ to CO than they do when converting H₂O to H₂. Consequently, a good indicator of the overall performance of electrodes is how well they will convert CO₂ to CO. It has been noted that the electrodes **102**, **104** described herein exhibit significantly improved electrocatalytic performance for the CO₂ to CO reaction when compared to historical electrode materials.

It should be recognized that various different processes and techniques may be used to fabricate the electrolyzer cell **100** and electrodes discussed herein. For example, the electrode materials discussed herein may be used to produce an ink which may be screen printed on the electrolyte layer **106**. In other embodiments, the electrodes may be sprayed or tape cast, laminated, and sintered onto the electrolyte layer **106**. Thus, electrodes in accordance with the invention are not limited any one method of fabrication, but rather encompass all processes and techniques used to fabricate electrodes **102**, **104** which incorporate the novel materials disclosed herein.

The following paragraphs provide one non-limiting example of a process for producing an electrolyzer cell **100** in accordance with the invention.

EXAMPLE

The material for the anode **102** was produced by initially synthesizing a perovskite composition having a formula of approximately Pr_{0.8}Sr_{0.2}MnO_(3-ε). This composition was synthesized using a liquid nitrate mix process. To produce this composition, stoichiometric proportions of constituent nitrates were mixed and charred over a hot plate at about 150° C. The resulting char was then mixed and calcined at approximately 1000° C. to form a perovskite crystalline phase. The synthesized perovskite powder was then milled in a container with partially stabilized zirconia media to reduce it to appropriate particle size distribution and specific surface area suitable for making screen-printable ink. An ink was then produced using a commercial organic vehicle (Heraeus V-006) and terpineol.

The material for the cathode **104** was also synthesized using a liquid nitrate mix process. Specifically, a mixture of nitrates of nickel, magnesium, cerium, and calcium powders were combined such that the Ni:Mg molar ratio was about 90:10. The final mixture contained about fifty to seventy percent by weight Ni(Mg)O (solid solution of nickel oxide and magnesium oxide) with the remainder being Ce(Ca)O_(2-ε) (ceria doped with calcium oxide). This mixture was then calcined at about 1000° C. and milled to achieve a desired particle size distribution and specific surface area. An ink was then produced using a process similar to that used for the anode **102**.

A batch of scandium-doped zirconia (Daiichi, Japan) was tape cast and sintered to provide electrolyte sheets of about 0.18 to 0.20 millimeters thick. The anode and cathode inks were then screen printed on opposite sides of the electrolyte sheet and sintered at about 1300° C. In some cases, the cathode ink was first printed and sintered at about 1400° C. The anode ink was then printed and sintered at a temperature between about 1250° C. and 1300° C.

The above-described process was used to create two different sizes of electrolyzer cells. The first size included button cells containing an electrolyte layer with a diameter of about forty millimeters. The button cells were used primarily for electrode characterization. The second size included square

cells measuring about ten centimeters along each edge. The square cells were used primarily for stack evaluation. Platinum mesh was attached to the electrodes of the button cells for current collection. A layer of scandium-doped lanthanum cobaltite ink was printed on the anode **102** and nickel ink was printed on the cathode **104** to provide current collection for the square cells. A stainless steel layer with gas flow passages and cell-to-cell interconnects were then placed in contact with the anode **102** and cathode **104** of the square cells.

Prior to testing the button cells and stacks of square cells, the anodes **102** and cathodes **104** were treated and infiltrated with a catalyst material. This was accomplished by saturating the porous electrodes **102**, **104** with a praseodymium cobalt nitrate solution. The molar ratio of praseodymium to cobalt in the solution was approximately one to one. The button cells and stacks of square cells were then heated for testing. In some cases, multiple infiltrations were conducted on the electrodes **102**, **104** along with heat treatments at various intermediate temperatures at or around 600° C.

FIGS. **3** and **4** show the performance of the button cells when operated in fuel cell and electrolysis modes. FIGS. **5** and **6** show the performance of the square cells in twenty-five cell stacks when operated in fuel cell and electrolysis modes.

Referring to FIG. **3**, as shown by the graph, the button cell exhibited an area specific resistance (ASR) of approximately 0.5 ohm-cm² in both fuel cell and electrolysis modes, showing the button cell performed equally well in either mode. This resistance was measured while feeding the button cell a mixture containing approximately fifty percent steam and fifty percent hydrogen at an operating temperature of around 800° C. The button cell's small size and environment allowed it to either absorb heat from or radiate heat to its surroundings. Thus, the button cell was substantially isothermal at a temperature of around 800° C. while operating in both fuel cell and electrolysis modes.

Referring to FIG. **4**, the long term performance of the button cell was observed over approximately 600 hours of operation. During that period, the button cell was switched between fuel cell and electrolysis modes of operation. As shown by the relationship between cell voltage and current density in both fuel cell and electrolysis modes, the ASR of the cell remained reasonably stable during the approximately 600 hours of operation.

Referring to FIG. **5**, as shown by the graph, a stack of twenty-five square cells exhibited an ASR of approximately 1.04 ohm-cm² in fuel cell mode and an ASR of approximately 1.3 ohm-cm² in electrolysis mode. In electrolysis mode, the stack had a core temperature of approximately 816° C. with steam utilization of approximately 49.4 percent. In fuel cell mode, the stack had a core temperature of approximately 896° C. with hydrogen utilization of approximately 72.4 percent. The ratio of hydrogen to steam was approximately 3:4.

The higher ASR of the cell stack compared to the button cell, as demonstrated by FIG. **3**, may be attributed to the longer current path and additional components (e.g., interconnects, etc.) that are used to implement the cell stack architecture. Furthermore, the variation in the ASR of the cell stack in fuel cell and electrolysis modes may be attributed primarily to the difference in core operating temperature for each mode, as was described in association with FIG. **2B**. That is, the approximately 80° C. difference in the stack core temperature in fuel cell and electrolysis modes causes a significant difference in the measured ASR. Thus, the button cell, as described in association with FIG. **3**, may provide a better picture of actual electrode performance since the cell was substantially isothermal in both fuel cell and electrolysis modes.

Referring to FIG. 6, the long term performance of the cell stack was observed over approximately 400 hours of operation. During that period, the cell stack was operated in electrolysis mode. As shown by the relationship between the cell stack voltage and current, the ASR of the cell stack remained reasonably stable during the approximately 400 hours of operation. The cell stack voltage was also monitored and measured for the high and low groups of five cells, as indicated by the lower data points of the graph.

The present invention may be embodied in other specific forms without departing from its spirit or essential characteristics. The described embodiments are to be considered in all respects only as illustrative and not restrictive. The scope of the invention is, therefore, indicated by the appended claims rather than by the foregoing description. All changes which come within the meaning and range of equivalency of the claims are to be embraced within their scope.

What is claimed is:

1. An electrolyzer cell comprising:
 - a cathode to reduce an oxygen-containing molecule to produce an oxygen ion and a fuel molecule, wherein the cathode comprises an electron conducting phase comprising nickel oxide intermixed with magnesium oxide; an electrolyte coupled to the cathode to transport the oxygen ion;
 - an anode coupled to the electrolyte to receive the oxygen ion to produce oxygen gas therewith, the anode comprising an electron-conducting phase comprising a perovskite having a chemical formula of substantially $(\text{Pr}_{(1-x)}\text{La}_x)_{(z-y)}\text{A}'_y\text{BO}_{(3-\delta)}$, wherein $0 \leq x \leq 0.5$, $0 \leq y \leq 0.5$, and $0.8 \leq z \leq 1.1$, and wherein A' comprises at least one alkaline earth metal and B comprises at least one transition metal.
2. The electrolyzer cell of claim 1, wherein the anode further comprises an ion-conducting phase intermixed with the electron-conducting phase, the ion-conducting phase comprising an oxide.
3. The electrolyzer cell of claim 2, wherein the oxide comprises zirconia doped with at least one of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, and cerium oxide.

4. The electrolyzer cell of claim 2, wherein the oxide comprises ceria doped with at least one of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, and zirconium oxide.

5. The electrolyzer cell of claim 1, wherein the anode is porous, the porosity thereof being infiltrated with an electrocatalyst.

6. The electrolyzer cell of claim 5, wherein the electrocatalyst comprises at least one of praseodymium, cobalt, cerium, europium, and other rare earth elements.

7. The electrolyzer cell of claim 1, wherein A' is an alkaline earth metal.

8. The electrolyzer cell of claim 7, wherein the alkaline earth metal comprises at least one of calcium and strontium.

9. The electrolyzer cell of claim 1, wherein B is a transition metal.

10. The electrolyzer cell of claim 9, wherein the transition metal comprises at least one of manganese, cobalt, and iron.

11. An electrolyzer cell comprising:

- a cathode to reduce an oxygen-containing molecule to produce an oxygen ion and a fuel molecule;
- an electrolyte coupled to the cathode to transport the oxygen ion;

- an anode coupled to the electrolyte to receive the oxygen ion to produce oxygen gas therewith, the anode comprising an electron-conducting phase comprising a perovskite having a chemical formula of substantially $(\text{Pr}_{(1-x)}\text{La}_x)_{(z-y)}\text{A}'_y\text{BO}_{(3-\delta)}$, wherein $0 \leq x \leq 0.5$, $0 \leq y \leq 0.5$, and $0.8 \leq z \leq 1.1$, and wherein A' comprises at least one alkaline earth metal and B comprises at least one transition metal; and

wherein the anode further comprises an ion-conducting phase intermixed with the electron-conducting phase, the ion-conducting phase comprising an oxide comprising ceria doped with at least one of yttrium oxide, ytterbium oxide, calcium oxide, magnesium oxide, scandium oxide, and zirconium oxide.

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