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# United States Patent [19]

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[54] **NON-CHROMATE CORROSION INHIBITORS FOR ALUMINUM ALLOYS**

4,615,918 10/1986 Reichert et al. .... 427/385.5

(List continued on next page.)

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FOREIGN PATENT DOCUMENTS

[73] Assignee: **PRC-DeSoto International, Inc.**, Burbank, Calif.

0296540 12/1988 European Pat. Off. .  
0337075 10/1989 European Pat. Off. .  
0639656 2/1995 European Pat. Off. .  
2259920 3/1993 United Kingdom .  
2302 092 8/1997 United Kingdom .

[\*] Notice: This patent is subject to a terminal disclaimer.

OTHER PUBLICATIONS

[21] Appl. No.: **09/329,613**

Charles R. Hegedus, Anthony T. Eng and Donald J. Hirst; Corrosion 92, The NACE Annual Conference and Corrosion Show; Self-Priming Topcoats for Steel Applications; (undated); Paper No. 466.

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CIBA Additives, CIBA-Geigy Corporation; IRGACOR 252 in Wash Primers Based on Beckopox Em 460/MOWITAL (Automotive Gray Primer): Jan. 1994.

**Related U.S. Application Data**

[63] Continuation-in-part of application No. 08/731,066, Oct. 9, 1996, Pat. No. 5,951,747.

Wilcox, G.D. et al., "The Role of Molybdates in Corrosion Prevention," Corrosion Reviews, vol. 6, No. 4, pp. 327-365 (1986) (no month).

[60] Provisional application No. 60/005,027, Oct. 10, 1995, and provisional application No. 60/005,028, Oct. 10, 1995.

Holmes, H., "Hazardous Waste Compliance is (Almost) Easy," Metal Finishing, vol. 87, pp. 65-66 (1989) (no month).

[51] **Int. Cl.**<sup>7</sup> ..... **C09D 5/08**; C08L 75/00

Albanese, J., "Regulatory Update—TCLP Rule Promulgated by EPA," Plating and Surface Finishing, vol. 77, pp. 120-121, 126 (1990) (no month).

[52] **U.S. Cl.** ..... **106/14.44**; 106/14.05; 106/14.11; 106/14.2; 106/14.13; 106/14.14; 106/14.15; 106/14.16; 106/14.17; 106/14.21; 106/14.34; 106/14.35; 106/14.37; 106/14.38; 106/14.39; 106/14.41; 106/14.42; 106/14.43; 106/14.45; 524/115; 524/175; 524/404; 524/413; 524/414; 524/415; 524/418; 524/442; 524/443

Carpenter, C.J., "The Air Force Hazardous Waste Minimization Program," Plating and Surface Finishing, vol. 77, pp. 35-36 (1990) (no month).

[58] **Field of Search** ..... 106/14.05, 14.11, 106/14.12, 14.13, 14.14, 14.15, 14.16, 14.17, 14.21, 14.34, 14.35, 14.37, 14.38, 14.39, 14.41, 14.42, 14.43, 14.44, 14.45; 524/115, 175, 404, 413, 414, 415, 418, 442, 443

Evans, A.T. et al., "The Adhesion and Corrosion of Chromate Containing Coatings on Aluminum," University of Manchester, Institute of Science and Technology, pp. 266-273 (1994) (no month).

[56] **References Cited**

U.S. PATENT DOCUMENTS

1,927,842 1/1933 McDermott ..... 568/701  
2,147,395 2/1939 Bayes ..... 252/72  
2,147,409 2/1939 Lamprey ..... 252/72  
2,624,708 1/1953 Langer et al. .... 508/276  
3,732,115 5/1973 Lankard et al. .... 106/19  
3,846,148 11/1974 Nordyke et al. .... 106/462  
3,852,087 12/1974 Nordyke et al. .... 106/287.17  
3,964,936 6/1976 Das ..... 148/273  
4,037,019 7/1977 Steger ..... 428/469  
4,110,129 8/1978 Matsushima et al. .... 148/6.15  
4,145,462 3/1979 Kuwabara et al. .... 427/160  
4,148,670 4/1979 Kelly ..... 148/247  
4,174,980 11/1979 Howell et al. .... 148/251  
4,264,378 4/1981 Oppen et al. .... 148/261  
4,298,404 11/1981 Greene ..... 148/247  
4,303,568 12/1981 Nay et al. .... 252/389.2  
4,313,769 2/1982 Frelin et al. .... 148/247  
4,324,684 4/1982 Geiger et al. .... 252/389 A  
4,338,140 7/1982 Reghi ..... 148/247  
4,370,177 1/1983 Frelin et al. .... 148/247  
4,387,027 6/1983 May et al. .... 210/697  
4,411,865 10/1983 Geiger et al. .... 422/16  
4,451,304 5/1984 Batiuk ..... 148/268  
4,457,790 7/1984 Lindert et al. .... 148/247  
4,470,853 9/1984 Das et al. .... 148/247  
4,539,051 9/1985 Hacias ..... 204/488  
4,608,191 8/1986 Wu ..... 252/391  
4,612,049 9/1986 Berner et al. .... 106/14.13

Rohrig, H., "Angriffvermindernde Zusätze zu das Aluminium angreifende Lösungen," Aluminium, vol. 17, pp. 559-562 (1935) (no month).

Foster, T. et al., "Development of a Chromate-Free Primer For Aluminum and Steel to Meet CGSB Specifications," Journal of Coatings Technology, vol. 63, No. 801 (1991) (no month).

(List continued on next page.)

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[57] **ABSTRACT**

A non-chromate, corrosion-inhibiting coating composition for metal surfaces includes at least one inhibitor selected from the group consisting of phosphates, phosphosilicates, silicates, and mixtures thereof, at least one inhibitor selected from the group consisting of titanates, zinc salts, and mixtures thereof, and a carrier for these inhibitors, the carrier capable of placing the inhibitors in proximity with the metal surface. In a preferred embodiment, the coating composition further includes a borate such as boric acid, and a succinate. A preferred phosphate includes calcium dihydrogen phosphate, and sodium titanium oxide is a preferred titanate. The zinc salt can include zinc phosphate or zinc cyanamide.

**23 Claims, No Drawings**



## U.S. PATENT DOCUMENTS

4,659,481	4/1987	Chen	210/697
4,696,763	9/1987	Bentley et al.	252/391
4,710,404	12/1987	Reichert et al.	427/386
4,740,393	4/1988	Cody et al.	427/388.1
4,812,174	3/1989	Kimura	148/251
4,921,552	5/1990	Sander et al.	148/247
4,992,116	2/1991	Hallman	148/247
5,026,440	6/1991	Finnenthal et al.	148/247
5,059,640	10/1991	Hegedus et al.	523/451
5,077,332	12/1991	Blatter et al.	524/389
5,089,064	2/1992	Reghi	148/247
5,089,551	2/1992	Hegedus et al.	524/396
5,125,989	6/1992	Hallman	148/247
5,126,074	6/1992	Bittner	252/387
5,129,967	7/1992	Sander et al.	148/247
5,135,681	8/1992	Carter et al.	252/389.62
5,135,999	8/1992	Wu et al.	526/224
5,147,472	9/1992	Reichgott	148/255
5,158,622	10/1992	Reighgott et al.	148/247
5,169,458	12/1992	Shulman et al.	148/248
5,192,374	3/1993	Kindler	148/272
5,202,367	4/1993	Hegedus et al.	524/204
5,236,983	8/1993	Hegedus et al.	524/202
5,244,512	9/1993	Kawasaki et al.	148/260
5,260,357	11/1993	Sachdeva	523/414
5,262,464	11/1993	Koevenig et al.	524/413
5,290,839	3/1994	Hegedus et al.	524/210
5,298,092	3/1994	Schriever	148/275
5,308,903	5/1994	Hegedus et al.	524/204
5,314,532	5/1994	Hughes et al.	106/253
5,328,525	7/1994	Musingo et al.	148/247
5,338,347	8/1994	Rohr et al.	106/14.44
5,342,456	8/1994	Dolan	148/247
5,344,505	9/1994	Ouyang et al.	148/261
5,358,623	10/1994	Bibber	205/201
5,378,293	1/1995	Schriever	148/275
5,389,405	2/1995	Purnell et al.	427/387
5,401,333	3/1995	Ouyang et al.	148/241
5,401,337	3/1995	Carlson et al.	148/257
5,403,880	4/1995	Hegedus et al.	524/204
5,411,606	5/1995	Schriever	148/240
5,412,011	5/1995	Morris et al.	524/261
5,415,687	5/1995	Schriever	106/1.27
5,951,747	9/1999	Lewis et al.	106/14.44

## OTHER PUBLICATIONS

- Bohner, V.H., "Ueber die angriffsvermindernde Wirkung von Wasserglaszusätzen zu Brom- und Chlorwasser bzw. Chlorkalk-Lösungen," *Aluminium*, vol. 13, pp. 347-348 (1931) (no month).
- Keohan, F. et al., "Process Parameters for the APK-2000 Adhesive Pretreatment System," *Moving Forward With 50 Years of Leadership in Advanced Materials Conference*, vol. 39, pp. 1130-1142 (1994) (no month).
- Rodgers, T., "Evaluation of Chromate Conversion Coatings on Aluminum by Electrochemical Impedance Spectroscopy," 2nd AESF Analytical Methods Symposium, 8 pages (1988) (no month).
- Mansfield, F., Wang, Y. and Shih, H., "The Ce-Mo Process for the Development of a Stainless Aluminum," *Electrochimica Acta.*, 37, 2277 (1992) (no month).
- Aklian, J.H., Lewis, K.J. and Zook, J.D., "Quantitative Methods of Predicting Relative Effectiveness of Corrosion Inhibitive Coatings," 1996 (no month).
- Harvey, G.J., et al., "Production and Performance of Pre-painted 55% Al-Zn Alloy Coated Steel Sheet in Australia," *International Conference on Zinc and Zinc Alloy Coated Steel Sheet-Galvatech*, pp. 367-373 (1989) (no month).
- Clearfield, A., "Crystalline Hydrous Zirconia," *Inorg. Chem.*, vol. 3, pp. 146-148 (no month).
- Mears, R.B. et al., "The Use of Inhibitors for Aluminum Chemical Equipment," *Trans. Electrochem. Soc.*, vol. 83, pp. 403-417 (1943) (no month).
- Thomas, J.F.J., "Inhibition of Corrosion of Aluminum and Other Metals in Soda Ash Solutions," *Can. J. Research*, vol. 21B, pp. 43-53 (1943) (no month).
- Weiler, W. "Aluminum als Werkstoff in der Textilveredelungsindustrie," *Deut. Farber-Ztg.*, vol. 74, p. 27 (1938) (no month).
- Robinson, E. A., "Silicates in Detergents . . .," *Soap and Sanitary Chem.*, vol. 28, pp. 34-36 (1952) (no month).
- Wilcox, G.D., "Passivation studies using group VIA anions," *Br. Corros. J.*, vol. 22, No. 4, pp. 254-258 (1987) (no month).
- Chakrabarty, C., Singh, M.M. and Agarwal, C.V., "1-Aryl-Substituted-3-Formamidino Thiocarbamide—An Efficient Corrosion Inhibitor for Aluminum Alloys in Nitric Acid," *Corrosion Inhibition, Proceedings of the International Conference on Corrosion Inhibition, May 1983, NACE-7*, 113 (1983).
- Reinhard, G., Simon, P. and Rammelt, U., "Application of Corrosion Inhibitors in Water-Borne Coatings," *Progress in Organic Coatings* 20, 383 (1992) (no month).
- Agarwala, V.S., "Corrosion Inhibition by Phenanthrolines," *Corrosion*, 46, 376 (May 1990).
- DeBerry, D.W., "Organic Inhibitors for Pitting Corrosion," *Reviews on Corrosion Inhibitor Science and Technology, NACE, II-19-1* (1989) (no month).
- S. Szklarska-Smialowska, "Inhibition of Pitting Corrosion," *Ohio State University, Dept. of Materials Science & Eng.* (undated).
- Natishan, P.M., McCafferty, E. and Hubler, G.K., "Surface Change Consideration in the Pitting of Ion-Implanted Aluminum," *J. Electrochem. Soc.*, 135, 321 (Feb. 1988).
- G. Parks, *The Isoelectric Points of Solid Oxides, Solid Hydroxides, and Aqueous Hydroxo Complex Systems*, *Chemical Reviews*, 65, 177 (1965) (no month).
- Cartledge, G.H., "Recent Studies of the Action of Inorganic Inhibitors," *Corrosion*, 18, 316t (1962) (no month).
- Hinton, B.R.W., "New Approaches to Corrosion Inhibition with Rare Earth Metal Salts," *Reviews on Corrosion Inhibitor Science and Technology, NACE, I-11-1* (1989) (no month).
- Foster, T., Blenkinsop, G.N., Blattler, P. and Szandorowski, M., "Search for a Chromate-Free Wash Primer," *J. Coatings Technology*, 63, 91 (1991) (no month).
- Service Bulletin for Borogard ZB (form of zinc borate-2ZnO-3B2O3·3.5H2O, U.S. Borax & Chemical Corp., Los Angeles, CA (1991) (no month).
- Service Bulletins for Busan 11M1 and Butrols 9100 & 9104 (barium metaborate, calcium metaborate and calcium pyroborate), *Buckman Laboratories, Memphis, TN* (1983, 1992 & 1994) (no month).
- Urquidi, M. and Macdonald, D.D., "Solute-Vacancy Interaction Model and the Effect of Minor Alloying Elements on the Initiation of Pitting Corrosion," *J. Electrochem. Soc.*, 132, 555 (1985) (no month).
- Cartledge, G.H., "Passivation and Activation of Iron in the Presence of Molybdate Ions," *Corrosion* 24, 223 (Aug. 1968).

- Kendig, M., Ryang, H.S., Liao, T., Cunningham, M. and Jeanjaquet, S., "Corrosion Induced Adhesion Loss—Low Volatile Organic Content (VOC) Coatings," Final Report, Contract No. N00014-90-C-0110 for Office of Naval Research, Rockwell International Science Center (Mar. 1993).
- Shaw, B.A., Davis, G.D., Fritz, T.L. and Olver, K.A., "A Molybdate Treatment for Enhancing the Passivity of Aluminum in Chloride-Containing Environments," *J. Electrochem. Soc.*, 137, 359 (Jan. 1990).
- Macdonald, D.D., "The Point Defect Model for the Passive State," *J. Electrochem. Soc.*, 139, 3434 (Dec. 1992).
- Sato, N., "Toward a More Fundamental Understanding of Corrosion Processes," *Corrosion*, 45, 355 (May 1989).
- Fontana, M.G., *Corrosion Engineering*, McGraw-Hill, 284 (1986) (no month).
- Gui, J.Y., Stern, D.A., Frank, D.G., Lu, F., Zapfen, D.C. and Hubbard, A.T., "Adsorption and Surface Structural Chemistry of Thiophenol, Benzyl Mercaptan, and Alkyl Mercaptans, Comparative Studies at Ag (111) and Pt (111) Electrodes by Means of Auger Spectroscopy, Low Energy Electron Diffraction, and Electrochemistry," *Langmuir*, 7, 955 (1991) (no month).
- McCafferty, E., "A Competitive Adsorption Model for the Inhibition of Crevice Corrosion and Pitting," *J. Electrochem. Soc.*, 137, 3731 (Dec. 1990).
- Crompton, J.S., Andrews, P.R. and McAlpine, E., "Characteristics of a Conversion Coating on Aluminum," *Surface and Interface Analysis*, 13, 160 (1988) (no month).
- Hawkins, J., Thompson, G.E., Wood, G.C. and Isaacs, H.S., "A Study of the Inhibition of Pitting of Aluminum by Chromate," *Proceedings of the 179th Electrochemical Society Meeting*, May 1991.
- Stewart, R., *Oxidation Mechanisms*, W.A. Benjamin, Inc., 33 (1964) (no month).
- Sugerman, G. and Monte, S.J., "Alkoxy Titanates and Zirconates As Corrosion Inhibitors," *Modern Paints and Coatings*, Jun. 1988.
- Hinton, B.R.W., "Corrosion Prevention and Chromate, the End of an Era?" *Metal Finishing*, Sep., 55 and Oct. 15 (Sep. 1991).
- Davis, G.D., Moshier, W.C., Fritz, T.L. and Cote, G.O., "Evolution of the Chemistry of Passive Films of Sputter-Deposited, Supersaturated Al Alloys," *J. Electrochem. Soc.*, 137, 422 (Feb. 1990).
- Chao, C.Y., Lin, L.F. and Macdonald, D.D., "A Point Defect Model for Anodic Passive Films, Parts I, II and III," *J. Electrochem. Soc.*, 128, 1187, 1194 (1981) and 129, 1874 (1982) (no month).



## NON-CHROMATE CORROSION INHIBITORS FOR ALUMINUM ALLOYS

### RELATED APPLICATIONS

This application is a continuation-in-part of U.S. patent application 08/731,066 filed Oct. 9, 1996, U.S. Pat. No. 5,951,747, which is based on provisional applications Ser. Nos. 60/005,027 and 60/005,028 filed Oct. 10, 1995, all of which are hereby incorporated by reference.

### FIELD OF THE INVENTION

This invention relates generally to methods and compositions for inhibiting corrosion, and in particular, to methods and compositions for inhibiting corrosion of metals, especially aluminum and other light metal alloys used in aircraft. Most specifically, the present invention relates to a non-chromate containing, corrosion-inhibiting coating composition capable of protecting a wide variety of metal surfaces.

### BACKGROUND OF THE INVENTION

The U.S. Environmental Protection Agency has identified the metal finishing industry as one of the one most significant contributors to environmental pollution in the United States and, in all likelihood, throughout the rest of the world. This is because the materials currently most used in metal finishing include chromium, cadmium, zinc, lead, copper, nickel, chromates, and many toxic or polluting volatile organics.

The aircraft industry, being one of the largest of the industrial finishers, provides an example of the environmental impact of these processes. In a 1990 report, Tinker Air Force Base (Oklahoma) reported producing and treating 1.4 million gallons of industrial waste water per day mostly from metal finishing. The principal contaminants were chromium, nickel, copper, cadmium, lead, zinc, tartrates, EDTA, phosphate, and ammonia. It is estimated that the cost of disposing of these wastes is approximately \$220 per ton, which equates to several thousand dollars per day for this one site.

The chromate ion, which is an excellent corrosion inhibitor, has been one of the most widely used for almost a hundred years. It is generally used as a pigment in corrosion inhibitive paints, wash primers, sealants and caulks. It is also used in chromate conversion coatings, etching solutions, and in sealing anodized and phosphate coatings.

For the past ten years, however, chromate has been recognized as toxic and carcinogenic, and because of its health risks, has become highly regulated. With pressure for elimination being exerted by government regulations, continued use of chromate will incur ever increasing economic penalties. Hence, there is urgent need for non-toxic substitutes, both from economic and environmental standpoints.

Currently, the most widely used inhibitors for passivating aerospace aluminum alloys and other light metals are the alkaline earth and zinc salts of hexavalent chromium. They vary mostly in their degree of water solubility (in the order  $Mg > Ca > Sr \geq Zn$ ) and to a much smaller extent in their pH ( $Sr \geq Ca > Mg$ ). The chromate anion is the active species, reliably performing four necessary functions to be more fully described below. In addition, all of the above-described hexavalent chromium salts enhance adhesion in many paint and sealing systems.

Chromate performs four functions, thereby making it a desirable inhibitor. These functions include:

1. Rapid exit from a carrier matrix, such as paint or sealant.
2. Adsorption of the chromate anion on the bare metal or metal oxide. This alters space charge distributions at the interface, lowering the isoelectric point of the protective anodic metal oxide layer naturally forming on active metals. This repels chloride attack, and/or shifts the corrosion potential of that metal and/or its pitting potential in the noble direction.
3. Instead of oxygen reduction at cathode sites, reduction of chromium (VI) anion occurs to form an acid-insoluble ("persisting"), chromium (III) oxide layer at cathode sites. This fills oxide voids over cathode sites and blocks further corrosion reaction.
4. Buffering the pH or neutralizing increasing acidity at the metal/electrolyte interface which comes from metal oxidization in the absence of air. Increasing acidity accelerates corrosion exponentially.

In addition to these desirable inhibitive functions, chromate salts have the advantages of: (1) promoting adhesion at the metal/resin interface under a coating or sealant compound; (2) working well on a wide variety of metal and alloy substrates because they passivate both anodically and cathodically; (3) being relatively neutral in pH; and (4) being strong oxidizers only in acid conditions, and thus not destroying or strongly reacting with the resin matrix in which they are placed.

The prior art discloses a number of non-chromate species which have some inhibitive capabilities. For example, U.S. Pat. No. 5,126,074 discloses "hydrogen phosphate" anions as exhibiting corrosion inhibitive activity on aluminum. The patent further discloses the use in coatings of alkaline earth monohydrogen phosphates, together with a carbonate of the same alkaline earth metal and an additive of alkaline fluorosilicate or fluoroborate or alkali or alkaline fluoride. This combination is said to prevent filiform corrosion on aerospace aluminum alloys.

Other references cite the dihydrogen phosphate anion as adsorbing on alumina and lowering its isoelectric point (IEP) from pH=9 to pH=5. Lowering the IEP of aluminum oxide on aluminum metal has been shown to increase its resistance to pitting. Since this species also exhibits buffering capability, it performs functions 2 and 4.

U.S. Pat. No. 2,624,708 discloses carcinogenic mercaptobenzothiazole (MBT) as an inhibitor for aluminum and steel. Sulphur and mercapto groups which are "soft bases" are known to have a high affinity for noble and other bare metal ("soft acid") surfaces. They are effective inhibitor structures under acidic conditions where no oxide is present. By itself, this species performs only function 2.

U.S. Pat. Nos. 4,457,790 and 5,125,989 disclose the use of Mannich adducts of vegetable tannin or polyalkenyl phenols to "conversion coat" aluminum. A titanium ion or compound such as fluorotitanic acid, among others, is claimed as a co-reactant. U.S. Pat. No. 5,129,967 discloses minute catalytic amounts of dihydrohexafluorotitanic acid and hydrofluoric acid used with much larger amounts of dihydrohexafluorozirconic acid and polyacrylic acid. These patents refer to usages on aluminum and/or aluminum alloys.

U.S. Pat. No. 5,314,532 discloses zinc, cobalt, nickel and lead cyanamide pigments as exhibiting corrosion inhibitive effects on silver and thin mirror coatings. Bare, oxide-free copper would be expected to show adsorption characteristics somewhat analogous to silver, especially in an acidic, crevice environment. As with silver groups, the availability of electrons on the cyano group act as a "soft base" on bare metal "soft acid" surfaces, performing function 2.



The present invention provides a corrosion-inhibiting coating composition which performs many, if not all of the same functions as a chromate-containing composition, but without the need for the harmful chromate species. The problem solved by this invention is the elimination of toxic hexavalent chromium salts which are known to be human carcinogens, as corrosion inhibitors from treatment solutions, coatings, and sealants used on aluminum and other metal alloys. The present invention provides for the synergistic combinations of non-chromate inhibitors for aerospace aluminum alloys and other metal surfaces which can be incorporated into both curable and non-curing sealants and into curable primer and uncoat systems. This synergistic combination of inhibitors can also be incorporated into water-containing or water-absorbing fluids that might cause corrosion when used in the proximity of metal such as de-icing liquids and coolants. These and other advantages of the present invention will be readily apparent from the description, discussion and examples which follow.

### SUMMARY OF THE INVENTION

There is disclosed herein a non-chromate, corrosion-inhibiting coating composition for metal surfaces. As used herein, "coating" is used to mean any composition which can cover a substrate, or which can place the inhibitors in proximity with a substrate. The composition comprises at least one inhibitor selected from the group consisting of phosphates, phosphosilicates, silicates, and mixtures thereof, at least one inhibitor selected from the group consisting of titanates, zinc salts, and mixtures thereof, and a carrier for these inhibitors, the carrier being capable of placing the inhibitors in proximity with the metal surface.

In particular embodiments, the coating composition of the present invention can further comprise a boric acid salt or boric acid, and/or a sulfur-containing succinate such as (2-benzothiazolylthio)succinic acid or amine salts thereof. A preferred phosphate is calcium dihydrogen phosphate. A preferred phosphosilicate is calcium, strontium zinc phosphosilicate. Sodium titanium oxide is a preferred titanate. Zinc phosphate, zinc cyanamide and zinc aliphatic carboxylates are the preferred zinc salts. The carrier comprises a solution or polymer matrix which adheres well to metal substrates and is capable of placing the inhibitors in close proximity with the metal surface. The coating composition of the present invention may also include ancillary ingredients such as pigments, rheological agents, and other performance additives.

### DETAILED DESCRIPTION OF THE INVENTION

The present invention is directed to the synergistic combinations of two to six individual corrosion inhibitors contributing two to five separate functions to the inhibition of corrosion on metals. The coating composition is particularly well suited for protecting light metal alloys, such as the aluminum alloys used in aircraft.

In the broadest sense, the present invention comprises at least one inhibitor selected from the group consisting of phosphates, phosphosilicates, silicates, and mixtures thereof, at least one inhibitor selected from the group consisting of titanates, zinc salts, and mixtures thereof, and a carrier for the inhibitors, the carrier capable of placing the inhibitors in proximity with a metal surface. The coating composition of the present invention can further comprise a boric acid salt or boric acid, and/or a succinate, preferably a sulfur-containing succinate. The succinate comprises a com-

pound selected from the group consisting of (2-benzothiazolylthio)succinic acid; an amine salt of (2-benzothiazolylthio)succinic acid, including volatile amines at standard temperature and pressure such as trimethylamine and  $\text{RNH}_3^+$  where R is a  $\text{C}_6\text{--C}_{24}$  alkyl; and mixtures thereof. In a preferred embodiment, the inhibitors are loaded into the carrier to achieve 3–40% by volume in a dried film, more preferably 5–25% by volume in a dried film, and most preferably 10–20% by volume in a dried film.

In a preferred embodiment, the phosphate comprises a dihydrogen phosphate, most preferably a compound selected from the group consisting of calcium dihydrogen phosphate, potassium dihydrogen phosphate, ammonium dihydrogen phosphate, sodium dihydrogen phosphate and mixtures thereof. The phosphate can also comprise a pyrophosphate, preferably sodium pyrophosphate, or a monohydrogen phosphate, preferably dipotassium monohydrogen phosphate.

The phosphosilicate of the present invention preferably comprises a compound selected from the group consisting of calcium, strontium phosphosilicate, calcium, strontium, zinc phosphosilicate, and mixtures thereof. The silicate of the present invention preferably comprises an orthosilicate, most preferably tetrasodium orthosilicate.

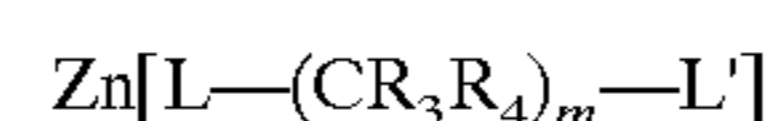
In a preferred embodiment of the present invention, the titanate comprises a titanium oxide, more preferably sodium titanium oxide.

The zinc salt of the present invention includes a variety of inorganic zinc compounds, illustratively including zinc borate, zinc bromate, zinc carbonate, zinc chlorate, zinc chloride, zinc cyanamide, zinc fluoride, zinc fluoroborate, zinc hydroxide, zinc iodate, zinc molybdate, zinc nitrate, zinc nitrite, zinc oxide, zinc permanganate, zinc phosphate, zinc dihydrogen phosphate, zinc polyphosphate, zinc pyrophosphate, zinc selenate, zinc orthosilicate, zinc fluorosilicates, zinc sulfate, zinc sulfite, zinc thiocyanate, hydrates thereof and mixtures thereof. The inorganic zinc compound preferably comprises a compound selected from the group consisting of zinc phosphate, zinc cyanamide, and mixtures thereof.

A zinc salt of the present invention also includes a variety of organo-zinc compounds illustratively including: salts of the formula:



where L is a moiety complexing with zinc selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfinio and sulfamyl;  $\text{R}_1$  and  $\text{R}_2$  are each independently hydrogen,  $\text{C}_1\text{--C}_4$  alkyl,  $\text{C}_1\text{--C}_4$  heteroatom containing alkyl, hydroxyl, carboxyl, and amino; n is an integer between 0 and 18 inclusive; mono-unsaturated analogs thereof where n is at least 2; di-unsaturated analogs thereof where n is at least 4; salts of the formula:



where L and L' are each independently carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; m is an integer between 1 and 8 inclusive;  $\text{R}_3$  and  $\text{R}_4$  are each independently selected from a group consisting of: hydrogen,  $\text{C}_1\text{--C}_4$  alkyl,  $\text{C}_1\text{--C}_4$  heteroatom containing alkyl, carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; zinc ethylene diamine tet-



raacetic acid (EDTA) complexes; and Zn containing polymers including polyacrylates, polyacetates, polymethacrylates, polyglycols and polyalcohols. The organo-zinc compound preferably comprises zinc unsubstituted mono and di-carboxylates. Still more preferably, n or m is an even integer. Preferably, the polymer has between 100 and 1,000,000 monomeric units. More preferably, the polymer has between 500 and 10,000 monomeric units. It is appreciated that the degree of zinc ion activity within the present invention is dictated by factors including zinc salt solubility in the coating composition, composition pH, zinc salt electronegativity, and steric effects associated with organic ligands. Owing to the variety of commercially available aliphatic complexing ligands available, the present invention affords greater control over the exit rate from a carrier matrix, zinc redox potential and pH buffering, as compared to the prior art use of zinc benzoate. It is further appreciated that the combination of two or more zinc salts including both inorganic zinc and organo-zinc compound is operative in controlling the availability of zinc as an inhibitor of the metal substrate useful lifetime.

Salts of boric acid are optionally included in the present invention. Boric acid acts as an electron acceptor and as such is a very weak acid having a pKa of 9.2 and is equally operative in the corrosion inhibitor compositions of the present invention. Thus, it is appreciated that while boric acid is detailed herein as an additive, a variety of salts thereof function in a similar manner. Boric acid salts operative in the present invention illustratively include alkali metal borates and tetraborates,  $M_3^+B_3O_6$  and  $M_2^+B_4O_7$ , respectively, where  $M^+$  is lithium, sodium and potassium; alkali earth borates  $M_3^{2+}(BO_3)_2$  where  $M^{2+}$  is magnesium, calcium, strontium and barium; borax; and hydrates thereof. Preferably, the boric acid salt is sodium tetraborate decahydrate or borax. More preferably, boric acid is used.

When used in the coating composition of the present invention, boric acid or a salt thereof should be loaded into the carrier so that it occupies up to 10% by volume in a dried film, more preferably 0.3–5.0% by volume in a dried film, and most preferably 0.5–2% by volume in a dried film. Likewise, when used as an inhibitor, calcium dihydrogen phosphate can be added in an amount up to 10% by volume in a dried film, more preferably 0.3–7% by volume, and most preferably 0.5–5% by volume in a dried film. When calcium, strontium phosphosilicate or calcium, strontium, zinc phosphosilicate is used as an inhibitor, it can be added in an amount up to 25% by volume in a dried film, more preferably 0.3–10% by volume, and most preferably 1–5% by volume in a dried film. If (2-benzothiazolythio)succinic acid or the fatty amine salt of (2-benzothiazolythio)succinic acid is used as an inhibitor in the coating composition of the present invention, it can be loaded into the carrier in an amount up to 10% by volume in a dried film more preferably 0.3–5% by volume, and most preferably 0.5–3% by volume in a dried film. Sodium titanium oxide can be added in an amount up to 10% by volume in a dried film, more preferably 1–10% by volume, and most preferably 3–6% by volume in a dried film. Zinc salts, when used in the present invention can be added in an amount up to 15% by volume in a dried film, more preferably 1–10% by volume in a dried film, and most preferably 3–6% by volume.

The carrier of the present invention is any compound which is capable of adhesion to a metal surface, and is also capable of placing the combination of inhibitors in proximity with the metal surface. Suitable carriers include both curable and noncuring sealants, as well as curable primer and uncoat systems. The combination of inhibitors of the

present invention can also be incorporated into water-containing or water-absorbing fluids that might cause corrosion when used in the proximity of metal such as de-icing liquids and coolants which are primarily comprised of glycol and water. Other carriers include, but are not limited to, aqueous solutions, amine-cured epoxy coatings, polyester and urethane coatings, sealant matrices such as those based on epoxy-cured polythioether polymers,  $MnO_2$  cured polysulfide sealants, non-curing end-capped polysulfide sealants, and other carriers as are known in the art.

Additional materials can be added to the coating composition of the present invention such as pigments, rheological agents, adhesion promoters and other additives, as are known in the art. For example, an oily, hydrophobic additive can be used with some of the inhibitor combinations of the present invention to achieve more acceptable barrier performance of sealants in water. Additionally, since the inhibitor ingredients of this invention do not necessarily enhance adhesion as do oxidizing inhibitors such as chromate, a porous conversion coating or anodized layer on the substrate is required in very wet environments for good mechanical adhesion of a coating or sealant to the metal.

As has been discussed above, chromate has been identified as an excellent inhibitor because it performs the following four necessary functions:

1. Rapid polymer exit and short term metal passivation.
2. Passivation by metal surface adsorption and double layer space charge altering.
3. Forming a water insoluble passivation layer which persists or remains insoluble in neutral, alkaline and acidic environments.
4. Passivation by control of pH or neutralization of acid at the metal electrolyte interface.

Therefore, in order to replace chromate, any combination of inhibitors must perform at least some, if not all of these functions in order to operate as a successful corrosion inhibitor. Additionally, transport of itself and other inhibitors out of a solid, e.g. polymeric, carrier matrix onto adjacent exposed bare metal areas in a moist but not liquid immersed environment is a performance parameter exhibited by one of the inhibitors (boric acid or salt thereof) of this invention that is not exhibited by chromate.

In general, the first and second inhibitive functions are performed by one or more moderately water-soluble, fast but reversibly adsorbing ingredients which exit readily from a coating or sealant, to give rapid, short-term passivation of a metal surface. Boric acid or a salt thereof appears to have the predominant fast or first-response effect out of a polymer matrix, and appears to also aid in releasing other, less soluble, more permanent passivators from the polymer.

The third inhibitive function is performed by a slower responding combination of two ingredients to form a water-insoluble, acid-resistant, "persisting" passivation layer. By "persisting" is meant remaining on a metal surface that was exposed to inhibitor-containing solution after that inhibitor-containing solution is removed and replaced with a corrosive solution such as aqueous sodium chloride containing no inhibitor. This layer is believed to form on the bare metal as well as on the oxide. The combination of a zinc salt and (2-benzothiazolythio)succinic acid or its fatty amine salt perform this function.

The fourth inhibitive function of these systems consists of buffering the pH, or controlling (neutralizing) the acidification of the metal-electrolyte interface environment in anaerobic or crevice conditions. Acid conditions dissolve the protective oxide and do not allow new insoluble oxide to form. In the formulations described by the present invention,



carboxylates, dihydrogen phosphates, monohydrogen phosphates, pyrophosphate, orthosilicate, sulfates, titanate, phosphosilicates and cyanamide can perform this function.

A variety of different quantitative and qualitative test methods have been employed to identify and corroborate passivation behavior on aircraft aluminum. These include:

1. Galvanic Current Measurement (on solutions, coatings and sealants)
  - a. Uses a titanium cathode electrically shorted to the active metal, usually Al alloy, immersed in aqueous NaCl solution. Stainless steel, Cd plated steel and carbon composite cathodes are also used.
  - b. Current between anode and cathode (Al and Ti or other, respectively) is measured at regular intervals and plotted versus time.
  - c. Configuration accelerates development of acidic crevice conditions in a narrow, air-deficient gap between the parallel, spaced-off anode and cathode metals.
  - d. Used to quantitatively test functions 1, 2 and/or 4.
  - e. We define good performance as 2–5 times the initial (i.e., within 24 hours) galvanic current reduction compared to uninhibited systems, with no rise over time out to 5–6 weeks. (Uninhibited coating and sealant initial currents are approximately 0.5–1 microamps per square centimeter on bare 2024 alloy with a Ti cathode, with a visual rise in current of 2–5 times within 1–2 weeks. Chromate gives approximately 0.2–0.3 microamps per square centimeter initially, with no rise over time. The corresponding values are approximately 3–5 times higher for bare 7075 alloy with a Ti cathode.)
2. Electrochemical Impedance Spectroscopy (EIS) (on solutions, coatings and sealants)
  - a. Uses an active metal (e.g. Al alloy) working electrode, passivated stainless steel counterelectrode, calomel reference electrode and Schlumberger potentiostat and frequency response analyzer.
  - b. Run in one or more of the following configurations:
    - i. Neutral or pH-adjusted, aerated NaCl solution, containing inhibitor(s).
    - ii. Active working electrode covered with inhibitor-containing, bonded coating or sealant film exposed to open, aerated NaCl solution on the upper film side. Measures barrier properties or corrosive-environment-penetration resistance ( $R_{pore}$ ) of bonded films.
    - iii. Active working electrode covered with inhibitor-containing, non-bonded, “free film” of coating or sealant, exposed to open, aerated aqueous NaCl solution on the upper side. In this configuration, when the film is penetrated with liquid, the whole air-deficient electrolyte/active metal interface area is wetted and therefore known dimensionally to calculate accurate charge transfer resistances ( $R_{ct}$ )
  - c. Used quantitatively to test functions 2, 3 and/or 4.
  - d. We define good electrochemical performance as  $R_{ct}$ 's of greater than  $10^6$  ohm-square centimeters, persisting over time. We define good barrier performance as exhibiting high (greater than  $10^8$  ohm-square centimeters), essentially immeasurable  $R_{pore}$ 's for as long as possible. We judge by making comparisons at equal film thicknesses. Low  $R_{pore}$  can be compensated for to a great extent by high  $R_{ct}$  as is the case with chromate.

3. Filiform Testing (on coatings only)
  - a. Uses brief (usually one hour) HCl vapor exposure to initiate acidic filiform (underfilm crevice) corrosion conditions at cut edges of coatings. Subsequent 80+% humidity exposure propagates further filiform corrosion.
  - b. Relative effectiveness of inhibitors are compared on the same substrates, as this test mostly differentiates on the basis of adhesion. Differences in corrosiveness and bondability of Al substrates (various alloys, Alclad, conversion-coated) typically give more variation for a given coating than the differences between inhibitors, including none, in the same coating, on the same substrate.
  - c. Use to qualitatively test function 4 and combined effects of functions 1–3.
  - d. Good performance for a coating or coated system is usually defined as the development from a cut edge of filiforms no greater than 3–6 millimeters in length, with the majority less than 3 millimeters, after 1000 or more hours.
4. Salt Spray Testing (on coatings and sealants)
  - a. Uses condensing corrosive (5%) salt solution to cause electrochemical activity of soluble inhibitors to be a factor, along with barrier properties, but the effects cannot be separately determined.
  - b. By leaving some uncoated area exposed, ability of inhibitors to move from the resin matrix to uncovered or damaged areas and to protect them is determined.
  - c. As in filiform testing, relative effectiveness of inhibitors are compared on the same substrates.
  - d. Used to qualitatively test function 1 and combined effects of functions 2–4.
  - e. Good performance is usually defined as no corrosion blister formation away from cut edges and minimal undercutting or corrosion blister formation at cut edges. Judgments are usually based on visual comparisons over time with uninhibited and chromate-inhibited controls.
5. pH Range Immersion Testing (on solutions only—solutions containing inhibitors, buffered from pH 3–10)
  - a. Found to be a predictor of “persisting” (insoluble) passivation layer formation and of resistance to prolonged-exposure, acidic-condition crevice corrosion.
  - b. Used to test function 3.
  - c. We define good performance as resistance to visual corrosion on flat surfaces and cut edges of bare metal over a wide pH range for as long as possible; on more than one alloy, if possible. Judgments are based on visual comparisons over time with uninhibited and chromate inhibited controls.

In order to work at all, effective inhibitors must exhibit at least function 2, corrosion current suppression. Functions 1, 3 and 4 increase their effectiveness. In addition, whether they perform on one or both of the most common aerospace aluminum substrates (AA2024-T3 or AA7075-T6) (function 5), and whether their addition significantly lowers the water resistance/adhesion of a polymer matrix (function 6) is considered.

The following table summarizes some of the preferred inhibitor combinations, rated according to how many of the above six functions they perform. This invention, however, is not limited to the combinations below. These abbreviations have been used to identify the various inhibitors in which:





-continued

FORMULATIONS (as volume % of non-volatiles)							
Water-borne Primer Coatings							
Components	BCINZ	BCINY	BCINW	BCNZW	BCNZ	BCNW	CINZ
Non-curing polysulfide resin	—	—	—	—	—	—	—
Epoxy-cured polythioether resin	—	—	—	—	—	—	—
Amine-cured epoxy resin	69.1	69.1	69.1	69.1	69.1	69.1	69.1
Adhesion promoters, dispersants, flow agents	1.7	1.7	1.7	1.7	1.7	1.7	1.7
Inert fillers	9.9	9.9	9.9	11.5	14.3	15.9	13.2
H	—	—	—	—	—	—	—
B	3.3	3.3	3.3	3.3	3.3	3.3	—
C	3.3	3.3	3.3	3.3	3.3	1.7	3.3
I	1.7	1.7	1.7	—	—	—	1.7
N	5.5	5.5	5.5	5.5	5.5	2.8	5.5
Z	5.5	—	—	2.8	2.8	—	5.5
Y	—	5.5	—	—	—	—	—
W	—	—	5.5	2.8	—	5.5	—
Solvent	(+28)	(+28)	(+28)	(+28)	(+28)	(+28)	(+28)
Water	(+200)	(+200)	(+200)	(+200)	(+200)	(+200)	(+200)

It will be appreciated from the foregoing, that a non-chromate containing coating composition having excellent properties for inhibiting corrosion of a variety of metal surfaces may be prepared from the synergistic combination of two to six individual corrosion inhibitors contributing two to five separate functions to the inhibition of metals, such as light metal alloys. The specific components of the composition will depend upon particular applications and factors such as the metal alloy substrate, the particular polymer or solution matrix in which the inhibitors are carried, and the range of exposure conditions the material will see in its particular location (e.g. aircraft interior, exterior, fuel tank, skin coating, lap seam, etc.).

The foregoing discussion and examples are merely meant to illustrate particular embodiments of the invention, and are not meant to be limitations upon the practice thereof. It is the following claims, including all equivalents, which define the scope of the invention.

We claim:

1. A nonchromate, corrosion-inhibiting coating composition for a metal surface, said composition containing inhibitors, said inhibitors comprising:

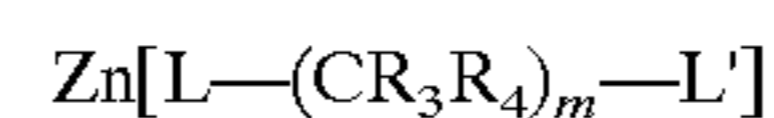
a synergistic amount of at least one compound selected from the group consisting of boric acid and salts thereof; dipotassium monohydrogen phosphate; calcium dihydrogen phosphate; ammonium dihydrogen phosphate; sodium dihydrogen phosphate; potassium dihydrogen phosphate; sodium pyrophosphate; calcium, strontium phosphosilicate; calcium, strontium, zinc phosphosilicate; tetrasodium orthosilicate; and mixtures thereof;

a synergistic amount of at least one compound selected from the group consisting of: sodium titanium oxide, an inorganic zinc compound and an organo-zinc compound wherein the organo-zinc compound is selected from the group consisting of:



wherein L is selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; R<sub>1</sub>

and R<sub>2</sub> are each independently hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, hydroxyl, carboxyl, and amino; n is an integer between 0 and 18 inclusive, mono-unsaturated analogs thereof, di-unsaturated analogs thereof



wherein L and L' are each independently selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; m is an integer between 1 and 8 inclusive; R<sub>3</sub> and R<sub>4</sub> are each independently selected from a group consisting of: hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl, zinc (EDTA) complexes, and zinc containing polymer; and an effective film-forming amount of a carrier for said inhibitors, said carrier capable of placing said inhibitors in proximity with said metal surface.

2. The coating composition of claim 1 and further comprising at least one compound selected from the group consisting of (2-benzothiazolylthio)succinic acid, an amine salt of (2-benzothiazolylthio)succinic acid, and mixtures thereof.

3. The coating composition of claim 1 wherein said inhibitors comprise 3-40 percent by volume in a dried film.

4. The coating composition of claim 1 wherein the inorganic zinc compound is selected from the group consisting of zinc borate, zinc bromate, zinc carbonate, zinc chlorate, zinc chloride, zinc cyanamide, zinc fluoride, zinc fluoroborate, zinc hydroxide, zinc iodate, zinc molybdate, zinc nitrate, zinc nitrite, zinc oxide, zinc permanganate, zinc dihydrogen phosphate, zinc phosphate, zinc polyphosphate, zinc pyrophosphate, zinc selenate, zinc orthosilicate, zinc fluorosilicates, zinc sulfate, zinc sulfite, zinc thiocyanate, and hydrates thereof.

5. The coating composition of claim 4 wherein the inorganic zinc compound is zinc dihydrogen phosphate.

6. The coating composition of claim 4 wherein the inorganic zinc compound is zinc cyanamide.

7. The coating composition of claim 1 wherein the organozinc compound is a zinc unsubstituted monocarboxylate or zinc unsubstituted dicarboxylate.



8. The coating composition of claim 1 wherein m is an even integer.

9. The coating composition of claim 1 wherein the zinc containing polymer is selected from a group consisting of polyacrylates, polyacetates, polymethacrylates, polyglycols and polyalcohols.

10. A non-chromate, corrosion-inhibiting coating composition for a metal surface, said composition containing inhibitors, said inhibitors comprising:

a synergistic amount of at least one compound selected from the group consisting of boric acid and salts thereof, calcium dihydrogen phosphate, and mixtures thereof; or

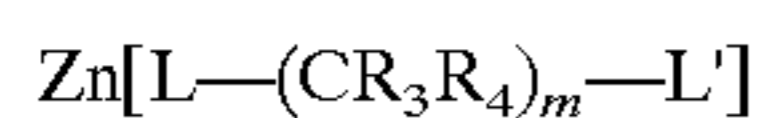
a synergistic amount of at least one compound selected from the group consisting of calcium, strontium phosphosilicate; calcium, strontium, zinc phosphosilicate; and mixtures thereof;

a synergistic amount of at least one compound selected from the group consisting of (2-benzothiazolylthio) succinic acid, the fatty amine salt of (2-benzothiazolylthio) succinic acid, and mixtures thereof;

a synergistic amount of at least one compound selected from the group consisting of: sodium titanium oxide, an inorganic zinc compound and an organo-zinc compound wherein the organo-zinc compound is selected from the group consisting of:



wherein L is selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; R<sub>1</sub> and R<sub>2</sub> are each independently hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, hydroxyl, carboxyl, and amino; n is an integer between 0 and 18 inclusive, mono-unsaturated analogs thereof, di-unsaturated analogs thereof



wherein L and L' are each independently selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; m is an integer between 1 and 8 inclusive; R<sub>3</sub> and R<sub>4</sub> are each independently selected from a group consisting of: hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl, zinc (EDTA) complexes, and zinc containing polymer; and

an effective film-forming amount of a carrier for said inhibitors, said carrier capable of placing said inhibitors in proximity with said metal surface.

11. The coating composition of claim 10 wherein the inorganic zinc compound is selected from the group consisting of zinc borate, zinc bromate, zinc carbonate, zinc chlorate, zinc chloride, zinc cyanamide, zinc fluoride, zinc fluoroborate, zinc hydroxide, zinc iodate, zinc molybdate, zinc nitrate, zinc nitrite, zinc oxide, zinc permanganate, zinc phosphate, zinc dihydrogen phosphate, zinc polyphosphate, zinc pyrophosphate, zinc selenate, zinc orthosilicate, zinc fluorosilicates, zinc sulfate, zinc sulfite, zinc thiocyanate, hydrates thereof.

12. The coating composition of claim 11 wherein the inorganic zinc compound is zinc phosphate.

13. The coating composition of claim 11 wherein the inorganic zinc compound is zinc cyanamide.

14. The coating composition of claim 10 wherein the organozinc compound is a zinc unsubstituted monocarboxylate or zinc unsubstituted dicarboxylate.

15. The coating composition of claim 10 wherein m is an even integer.

16. The coating composition of claim 10 wherein the zinc containing polymer is selected from a group consisting of polyacrylates, polyacetates, polymethacrylates, polyglycols and polyalcohols.

17. The coating composition of claim 10 wherein the boric acid or salt thereof comprises 0.3-10 percent by volume in a dried film.

18. The coating composition of claim 10 wherein said calcium dihydrogen phosphate comprises 0.3-10 percent by volume in a dried film.

19. The coating composition of claim 10 wherein said phosphosilicate comprises 0.3-25 percent by volume in a dried film.

20. The coating composition of claim 10 wherein said succinate compound comprises 0.3-10 percent by volume in a dried film.

21. The coating composition of claim 10 wherein said sodium titanium oxide comprises 1-10 percent by volume in a dried film.

22. The coating composition of claim 10 wherein the inorganic zinc compound comprises 1-15 percent by volume in a dried film.

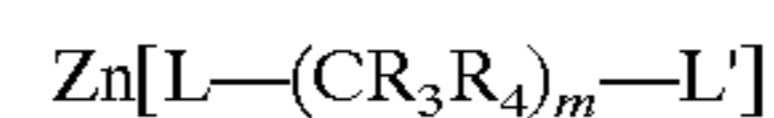
23. A method for inhibiting corrosion on a metal surface, said method comprising contacting said metal surface with a coating composition comprising:

a synergistic amount of at least one compound selected from the group consisting of boric acid and salts thereof; dipotassium monohydrogen phosphate; calcium dihydrogen phosphate; ammonium dihydrogen phosphate; sodium dihydrogen phosphate; potassium dihydrogen phosphate; sodium pyrophosphate; calcium, strontium phosphosilicate; calcium, strontium, zinc phosphosilicate; tetrasodium orthosilicate; and mixtures thereof;

a synergistic amount of at least one compound selected from the group consisting of: sodium titanium oxide, an inorganic zinc compound and an organo-zinc compound wherein the organo-zinc compound is selected from the group consisting of:



wherein L is selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; R<sub>1</sub> and R<sub>2</sub> are each independently hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, hydroxyl, carboxyl, and amino; n is an integer between 0 and 18 inclusive, mono-unsaturated analogs thereof, di-unsaturated analogs thereof



wherein L and L' are each independently selected from carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl; m is an integer between 1 and 8 inclusive; R<sub>3</sub> and R<sub>4</sub> are each independently selected from a group consisting of: hydrogen, C<sub>1</sub>-C<sub>4</sub> alkyl, C<sub>1</sub>-C<sub>4</sub> heteroatom containing alkyl, carboxyl, hydroxyl, sulfonyl, phosphyl, phosphonyl, sulfinyl, amino, cyano, imino, mercapto, nitro, nitroso, siloxy, sulfamino and sulfamyl, zinc (EDTA) complexes, and zinc containing polymer; and

an effective film-forming amount of a carrier for said inhibitors, said carrier capable of placing said inhibitors in proximity with said metal surface.