

# United States Patent [19]

Quinlan

[11] Patent Number: **4,539,140**

[45] Date of Patent: **Sep. 3, 1985**

[54] MIXTURES OF NON-HALOGEN SALTS OF NITROGEN HETEROCYCLICS AND NITROGEN-SULFUR HETEROCYCLICS

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[21] Appl. No.: 588,639

[22] Filed: Mar. 12, 1984

### Related U.S. Application Data

[62] Division of Ser. No. 200,817, Oct. 27, 1980.

[51] Int. Cl.<sup>3</sup> ..... C23F 11/04; C09K 31/00

[52] U.S. Cl. .... 252/391; 252/390; 422/12; 422/16

[58] Field of Search ..... 252/390, 392, 391; 422/12, 16

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### [57] ABSTRACT

This invention relates to non-halogen salts of nitrogen and/or nitrogen-sulfur heterocyclics as corrosion inhibitors for stainless steel-containing systems. These are more effective as corrosion inhibitors when employed in combination with surfactants, preferably of the non-ionic type such as oxyalkylates.

**12 Claims, No Drawings**



**MIXTURES OF NON-HALOGEN SALTS OF  
NITROGEN HETEROCYCLICS AND  
NITROGEN-SULFUR HETEROCYCLICS**

This is a division of application Ser. No. 200,817, filed Oct. 27, 1980.

This invention relates to corrosion inhibitors such as those designed for use in acid cleaning solvents which are used in cleaning industrial high temperature boilers and steam generators. Because of the presence of stainless steel components in these boilers and generators, inhibitors which contain halides have to be avoided since halide contamination will cause stress-corrosion cracking in these components.

In the patent application Ser. No. 101,953 filed Dec. 10, 1979, N. S. Thompson, now U.S. Pat. No. 4,446,056, there is described and claimed nitrogen and/or sulfur-nitrogen heterocycles. While they are useful as corrosion inhibitors in mineral acids, they are not as suitable by themselves in the present system.

I have now discovered that when the compositions described in the above application are reacted with an acid, excluding halogen acids, they become more soluble in the acidic medium in which they are to be used.

I have further discovered that the effectiveness of the salts can further be enhanced by the presence of surfactants, particularly non-ionic surfactants such as oxyalkylated surfactants. The use of surfactants, in the compositions of this invention, not only enhances the overall corrosion inhibition, but also produces a composition of excellent solubility in the various acid cleaning solvents. In effect there is no "gunking out" of the inhibitor or residual film left on the cleaning surface after use.

Thus, this invention relates to acid corrosion inhibitor composition comprising:

- (1) An acid salt(s) of nitrogen and/or sulfur-nitrogen heterocycles described in patent application Ser. No. 101,953.
- (2) A surfactant which is preferably an oxyalkylated surfactant.

Non-limiting examples of suitable acids that may be employed to form the salt(s) of this invention include: sulfuric, nitric, phosphoric, acetic, formic, hydroxyacetic, organic acids such as sulfonic acids, etc., and the like.

Any suitable surfactant can be employed. The surfactants which are most usually employed in the practice of this invention are oxyalkylated surfactants or more specifically polyalkylene ether or polyoxyalkylene surfactants. Oxyalkylated surfactants as a class are well known. The possible sub-classes and specific species are legion. The methods employed for the preparation of such oxyalkylated surfactants are also too well known to require much elaboration. Most of these surfactants contain, in at least one place in the molecule and often in several places, an alkanol or a polyglycoether chain. These are most commonly derived by reacting a starting molecule, possessing one or more oxyalkylatable reactive groups, with an alkylene oxide such as ethylene oxide, propylene oxide, butylene oxide, etc. However, they may be obtained by other methods such as shown in U.S. Pat. Nos. 2,588,771 and 2,596,091-3, or by esterification or amidification with an oxyalkylated material, etc. Mixtures of oxides may be used as well as successive additions of the same or different oxides may be employed. Any oxyalkylatable material may be employed. As typical starting materials may be mentioned

alkyl phenols, phenolic resins, alcohols, glycols, amines, organic acids, carbohydrates, mercaptans, and partial esters of polybasic acids. In general, the art teaches that, if the starting material is water-soluble, it may be converted into an oil-soluble surfactant by the addition of polypropoxy or polybutoxy chains. If the starting material is oil-soluble, it may be converted into a water soluble product. Subsequent additions of ethoxy units to the water-soluble surfactant by the addition of polyethoxy chains tend to increase the water solubility, while, subsequent additions of high alkoxy chains tend to increase the oil solubility. In general, the final solubility and surfactant properties are a result of a balance between the oil-soluble and water-soluble portions of the molecule.

In the practice of this invention I have found that suitable surfactants may be prepared from a wide variety of starting materials. For instance, if I begin with an oil-soluble material such as a phenol or a long chain fatty alcohol and prepare a series of products by reaction with successive portions of ethylene oxide, I find that the members of the series are successively more water-soluble. Similarly it is possible to start with water or a water-soluble material such as polyethylene glycol and add, successively, portions of propylene oxide. The members of this series will be progressively less water-soluble and more oil-soluble. There will be a preferred range where the materials are useful for the practice of this invention.

In general, the compounds which would be selected are oxyalkylated surfactants of the general formula



wherein Z is the oxyalkylatable material, A is the radical derived from the alkylene oxide which can be, for example, ethylene, propylene, butylene, and the like, n is a number determined by the moles of alkylene oxide reacted, for example 1 to 2000 or more and m is a whole number determined by the number of reactive oxyalkylatable groups. Where only one group is oxyalkylatable as in the case of a monofunctional phenol or alcohol R'OH, then m=1. Where Z is water, or a glycol, m=2. Where Z is glycerol, m=3, etc.

In certain cases, it is advantageous to react alkylene oxides with the oxyalkylatable material in a random fashion so as to form a random copolymer on the oxyalkylene chain, i.e., the  $[(OR)_nOH]_m$  chain such as —AA-bAAABBABABBABBA—. In addition, the alkylene oxides can be reacted in an alternate fashion to form block copolymers on the chain, for example



or



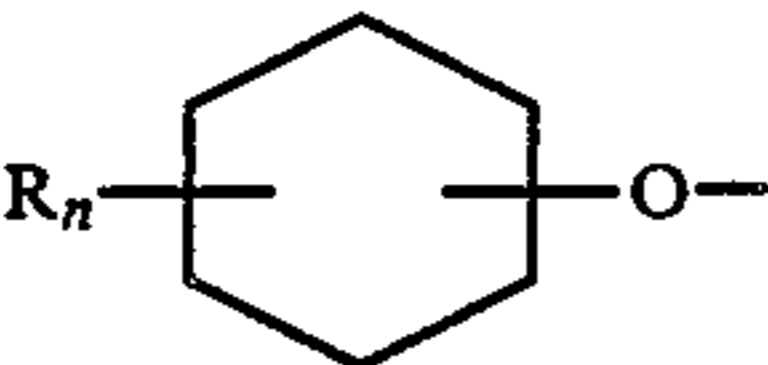
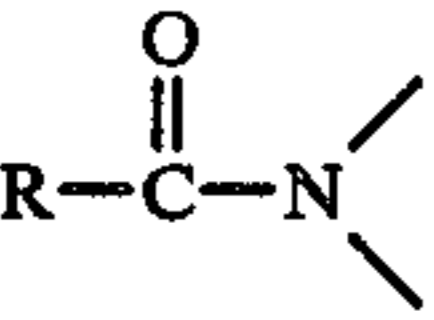

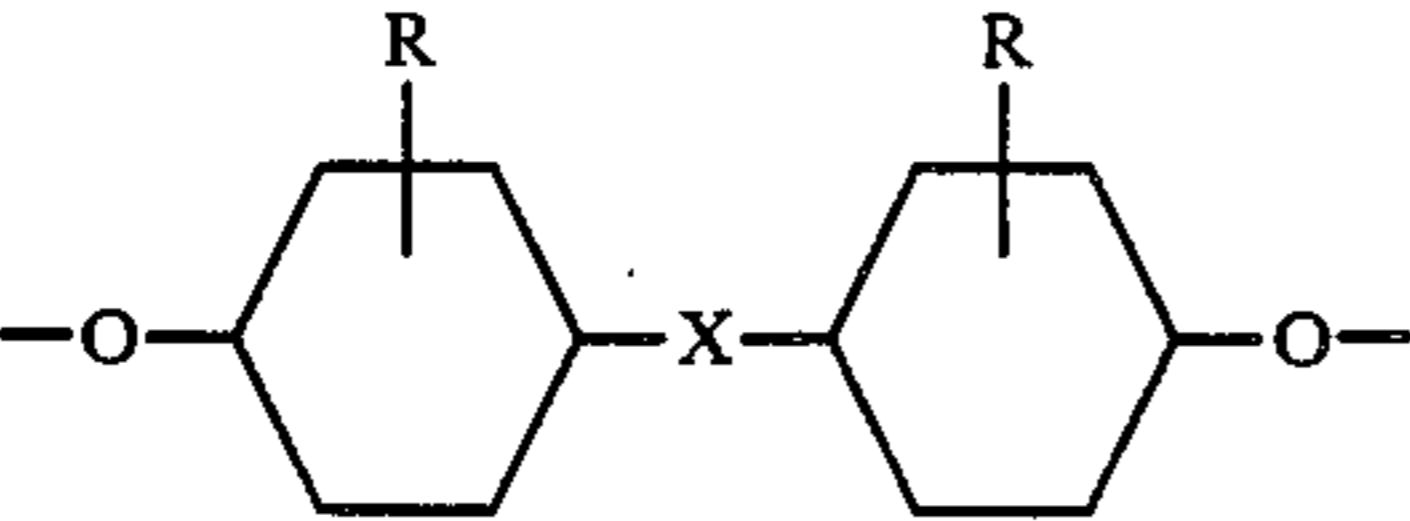
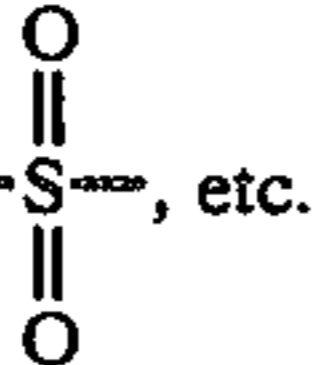
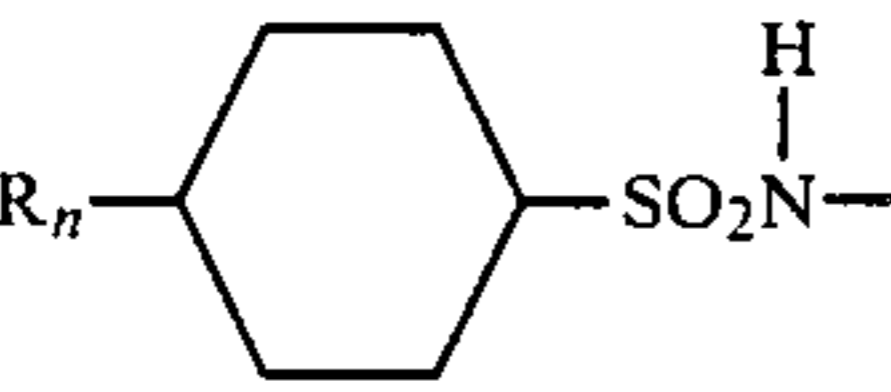
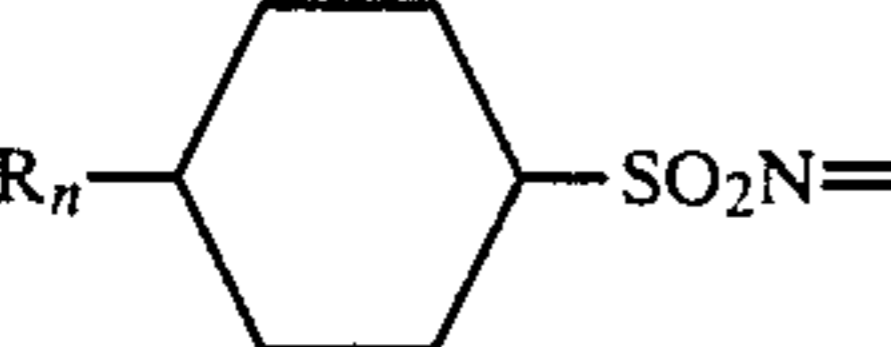
where A is the unit derived from one alkylene oxide, for example ethylene oxide, and B is the unit derived from a second alkylene oxide, for example propylene oxide, and C is the unit derived from a third alkylene oxide, for example, butylene oxide, etc. Thus, these compounds include terpolymers or higher copolymers polymerized randomly or in a blockwise fashion or many variations or sequential additions.

Thus,  $(OR)_n$  in the above formula can be written  $-A_aB_bC_c-$  or any variation thereof, wherein a, b and

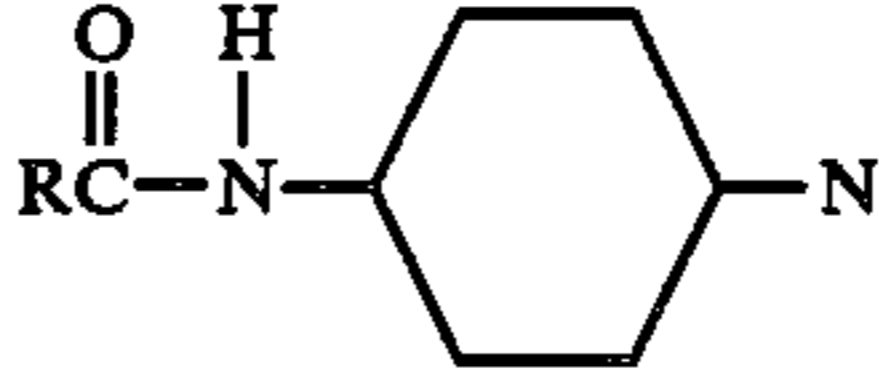
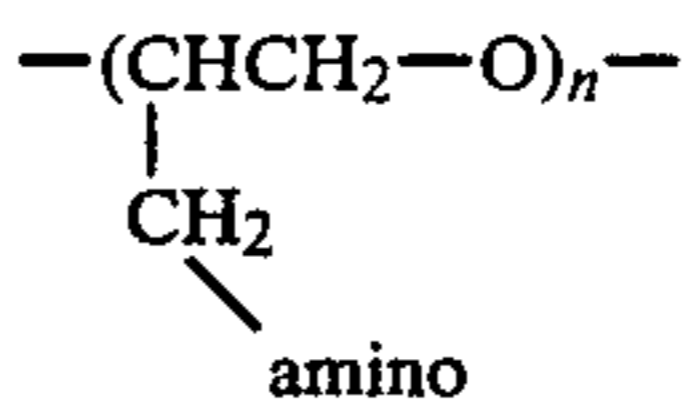


c are 0 or a number provided that at least one of them is greater than 0.

It cannot be overemphasized that the nature of the oxyalkylatable starting material used in the preparation of the surfactant is not critical. Any species of such material can be employed. By proper additions of alkylene oxides, this starting material can be rendered suitable as a surfactant and its suitability can be evaluated by testing in the corrosion system.

REPRESENTATIVE EXAMPLES OF Z	
No.	Z
1	$\begin{array}{c} \text{O} \\ \parallel \\ \text{RC}-\text{O}- \end{array}$
2	
3	R-O-
4	R-S-
5	$\begin{array}{c} \text{O} \quad \text{H} \\ \parallel \quad   \\ \text{R}-\text{C}-\text{N} \end{array}$
6	
7	$\begin{array}{c} \text{H} \\   \\ \text{R}-\text{N}- \end{array}$
8	
9	Phenol-aldehyde resins.
10	-O- (Ex: Alkylene oxide block polymers).
11	 X=O, -S-, -CH <sub>2</sub> -  , etc.
12	$\text{R}-\text{S}-\text{CH}_2-\text{C}(=\text{O})-\text{O}-$
13	RPO <sub>4</sub> H
14	RPO <sub>4</sub> =
15	PO <sub>4</sub> =
16	
17	

-continued

REPRESENTATIVE EXAMPLES OF Z	
No.	Z
5	18 
10	19 Polyol-derived (Ex: glycerol, glucose, pentaerythritol). 20 Anhydrohexitan or anhydrohexide derived (Spans and Tweens). 21 Polycarboxylic derived.
15	22 $-(\text{CHCH}_2-\text{O})_n-$ 

20 Examples of oxyalkylatable materials derived from the above radicals are legion and these, as well as other oxyalkylatable materials, are known to the art. A good source of such oxyalkylatable materials, as well as others, can be found in "Surface Active Agents and Detergents," vols. 1 and 2, by Schwarz et al., Interscience Publishers (vol. 1, 1949, vol. 2, 1958), and the patents and references referred to therein.

25 Non-limiting examples of the preparation of the nitrogen and/or nitrogen-sulfur heterocyclics and the acid  
30 salt(s) thereof are as follows.

## EXAMPLE 1

To a suitable reactor is charged 300 cc of aqueous ammonia (NH<sub>4</sub>OH~58%). Butyraldehyde (200 grams)  
35 is added at a rate such that the reaction mixture is maintained at 35°-45° C. Addition time is about 45 min. Stirring is continued for an additional 90 minutes and then the mixture is allowed to settle and separate into two phases. The upper organic phase was identified as  
40 2,4,6-tripropyl 1,3,5-hexahydrotriazine, containing 1 mole of water. Anal. calculated for C<sub>12</sub>H<sub>27</sub>N<sub>3</sub>. H<sub>2</sub>O: N, 18.2; Found 17.9.

The organic phase (200 grams) was charged to a 500 ml reactor equipped with stirrer, thermometer and off-gas condenser attached to a Dean Stark trap. Ammonium nitrate, 2 grams was added and the mixture heated gradually to 180° C. over a 4 hour period. During this period copious evolution of ammonia gas took place;  
45 low boiling products, mainly water, were removed during the reaction. The product (160 gm) was a clear amber liquid. The product was separated by preparative gas chromatography into two major components. These were identified as 2-propyl 3,5-diethyl pyridine and 1-butyl, 2-propyl, 3,5-diethyl 1,2-dihydropyridine.  
50  
55

## EXAMPLE 2

To 100 grams of 2,4,6-tripropyl 1,3,5-hexahydrotriazine prepared as described in Example 1, was added 1.5  
60 grams ammonium nitrate and 1.0 gram sulfur. The mixture was heated as described in Example 1 and the resulting product was a clear amber stench liquid. Analysis showed incorporation of 90% of the sulfur added.

Using the reactants and procedure of Example 2 except for the following % of sulfur and/or sulfur-containing reactant, the following sulfur-containing compositions were prepared. The results are summarized in Table 1.  
65



TABLE 1

Example No.	Sulfur Compound	Wgt. of Sulfur-containing Reactant	% Sulfur of wgt. of Reactants
3	sulfur	5	5
4	sulfur	10	10
5	sulfur	20	20

## EXAMPLE 6

To a cooled aqueous 28% ammonia solution (300 cc) and sulfur (10 grams) was added 200 grams butyraldehyde at such a rate that a temperature of  $\leq 35^\circ$  C. was maintained. The mixture was stirred for 24 hours and the resulting layers were separated, in an upper organic phase and lower aqueous phase. The organic phase was charged to a reactor equipped with stirrer, thermometer, and an off-gas condenser attached to a Dean Stark trap. Ammonium nitrate, 2 grams, was added and the mixture heated gradually to  $180^\circ$  C. over a period of 4 hours. During this period copious evolution of ammonia gas took place; low boiling products, mainly water, were removed during the reaction. The resulting product was a clear amber stench liquid. Analysis indicated  $\geq 90\%$  of sulfur charged present in the product.

As described in example 6, the following sulfur-containing compositions were prepared. The details are summarized in Table 2.

TABLE 2

Example No.	Sulfur Compound	Wt. Sulfur Compound	% Sulfur of Wgt. of Reactants
7	sulfur	1.0 g	0.5
8	sulfur	5.0 g	2.5
9	sulfur	20.0 g	10.0

## EXAMPLE 10

Place in a pressure reactor, a mixture of 200 grams cyclohexanone and 10 grams ammonium nitrate. Add with stirring ammonia gas and maintain a pressure of 30–40 psi while keeping the reaction temperature between  $50^\circ$ – $60^\circ$  C. Continue stirring for 10–24 hours or until the reaction shows no sign of taking up more ammonia. When the reaction is over the stirrer is turned off and the mass upon settling separates into two layers. The lower viscous organic which is identified as 2,2,4,4-dipentamethylene 5,6-tetramethylene 2,2,4,5-tetrahydropyrimidine is charged to a 500 ml reactor equipped with stirrer, thermometer and an off-gas condenser attached to a Dean Stark trap. Carefully add 4 grams nitric acid or 10 grams ammonium nitrate and heat the mixture gradually to  $190^\circ$ – $200^\circ$  C. for 5–6 hours while removing any water or low boiling distillate. Evolution of ammonia gas commences at about  $110^\circ$ – $120^\circ$  C. and is very strong at  $150^\circ$  C. The progress and completion of the reaction can be followed by the decrease in the off-gas evolution and by the change in the infrared spectral characteristics, the peak at  $6.02\mu$ ,  $-\text{C}=\text{N}$ , disappears and a strong new peak at  $6.4\mu$  (pyridine) appears. The final product (175 grams) is a dark, very viscous liquid and is identified as 9-pentyl 1,2,3,4,5,6,7,8-octahydrophenanthridine).

## EXAMPLE 11

To 100 grams of the intermediate pyrimidine prepared as described in Example 10 was added 10 grams sulfur and 5 grams ammonium nitrate and the mixture

heated gradually to  $190^\circ$ – $200^\circ$  C. for 6 hours. The resulting product was a dark, very viscous stench liquid. Analysis showed  $\geq 85\%$  of sulfur charged present in the final product.

## EXAMPLE 12

Same as Example 11 except 5.0 grams sulfur was added.

## EXAMPLE 13

Same as Example 11 except 1.0 gram sulfur was added.

## EXAMPLE 14

Place in a pressure reactor a mixture of 200 grams cyclohexanone, 10 grams sulfur and 10 grams ammonium nitrate. Add with stirring ammonia gas and maintain a pressure of 30–40 psi while keeping the reaction temperature between  $50^\circ$ – $60^\circ$  C. This reaction is exothermic and some cooling might be required. Continue stirring for 10–20 hours. When the reaction is over, the stirrer is turned off and the mass upon settling separates into two layers. The lower organic layer (205 grams) is drawn off and charged to a 500 ml reactor equipped with stirrer, thermometer and a Dean Stark trap attached to a reflux condenser. Carefully add 4 grams nitric acid or 10 grams ammonium nitrate and heat the mixture gradually to  $190^\circ$ – $200^\circ$  C. for 6 hours. The resulting product is a dark, viscous stench liquid. Gas chromatography shows the product to consist mainly of two components, phenanthridine and 2,2-pentamethylene 4,5-tetramethylene 3-thiazoline. Analysis showed  $\geq 85\%$  of sulfur charged present in the final product.

## EXAMPLE 15

Same as Example 14, except 20 grams sulfur was added. This product was a semi-solid.

## EXAMPLE 16

Same as Example 14, except 5 grams sulfur was added.

## EXAMPLE 17

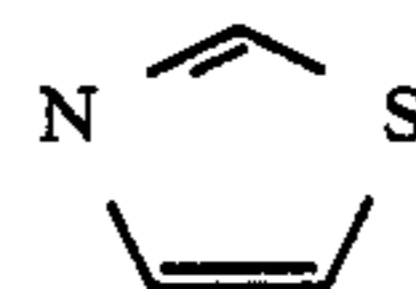
Same as Example 14, except 2 grams sulfur was added.

TABLE 3

Example No.	Sulfur Compound	Wt. Sulfur Compound	% Sulfur of Wgt. of Reactants
18	—	—	—
19	sulfur	10	10
20	"	5	5
21	"	1	1
22	"	10	5
23	"	20	10
24	"	5	2.5
25	"	2	1.0

The sulfur-containing heterocyclic products are primarily thiazoles and thiazolines characterized by the presence of the following moieties:

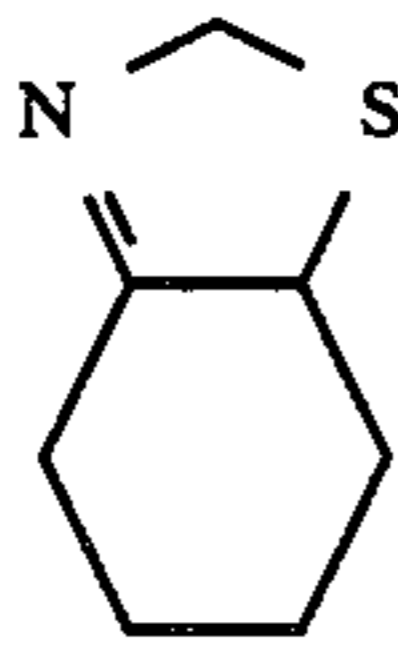
Thiazole



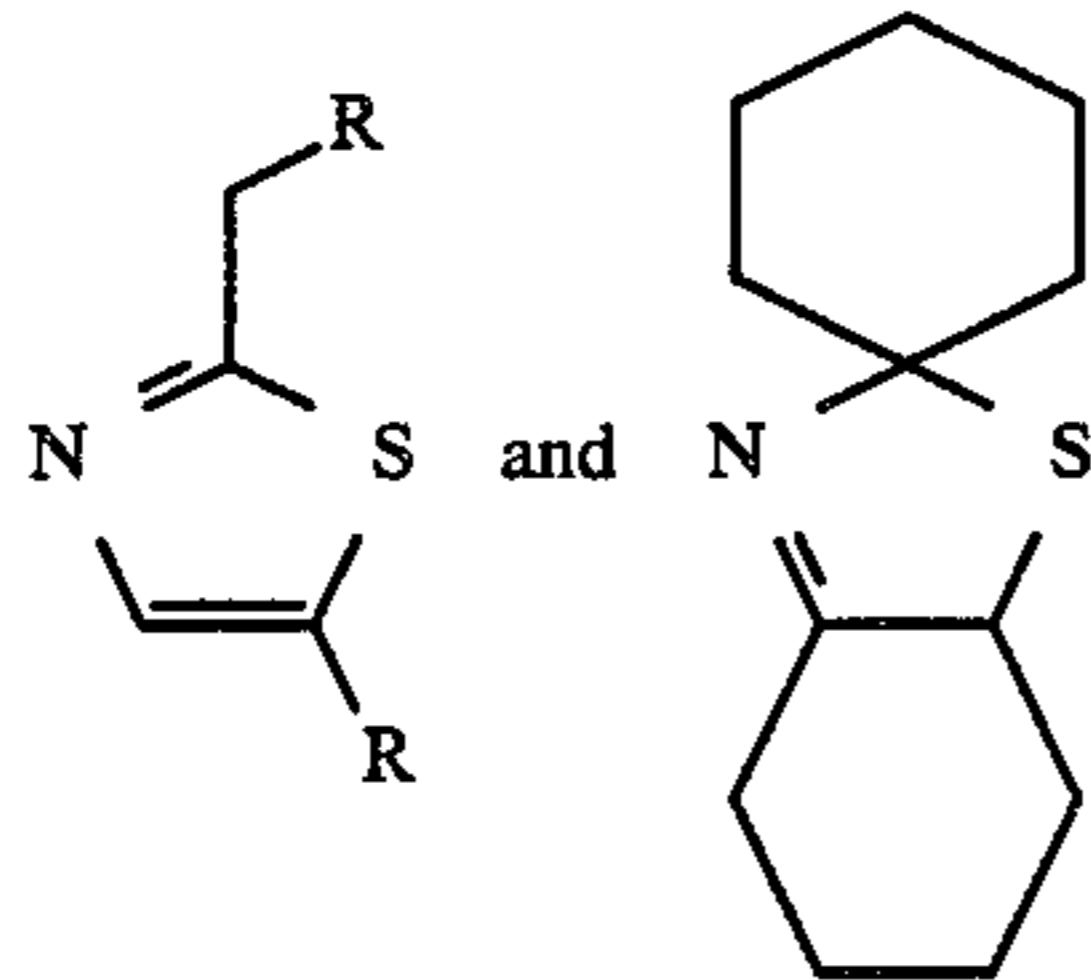
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Thiazoline



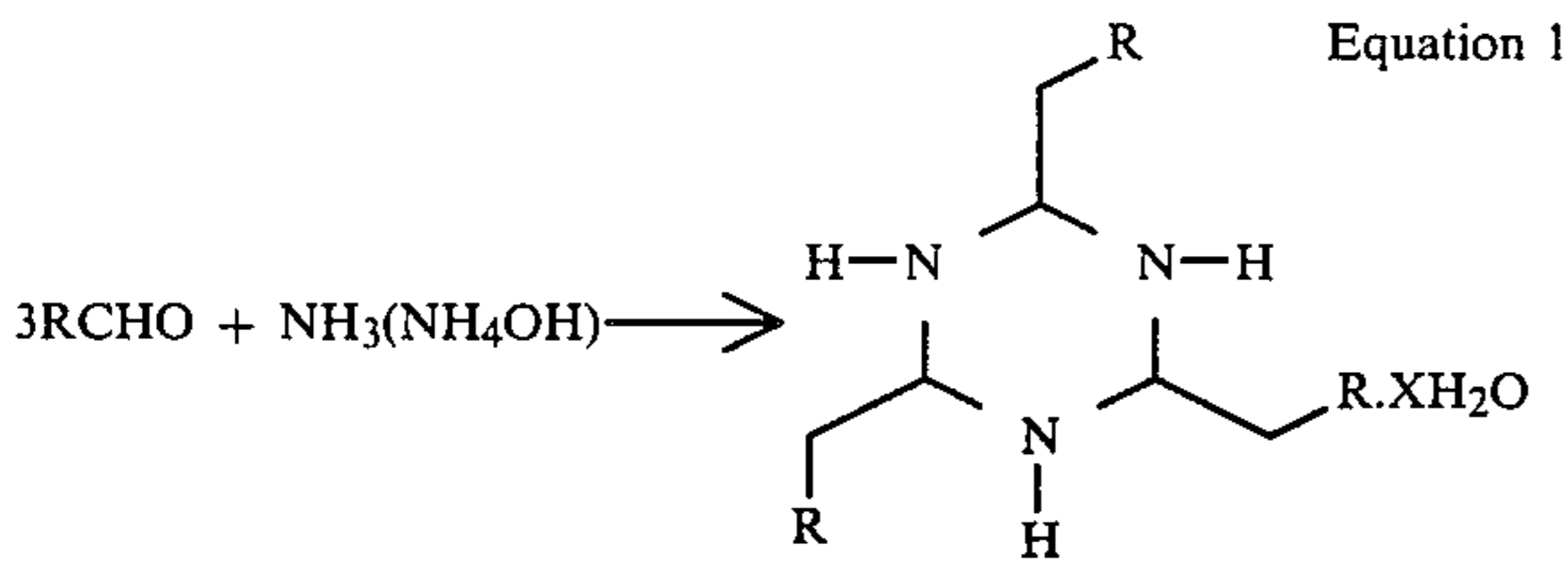
and substituted derivatives thereof, for example



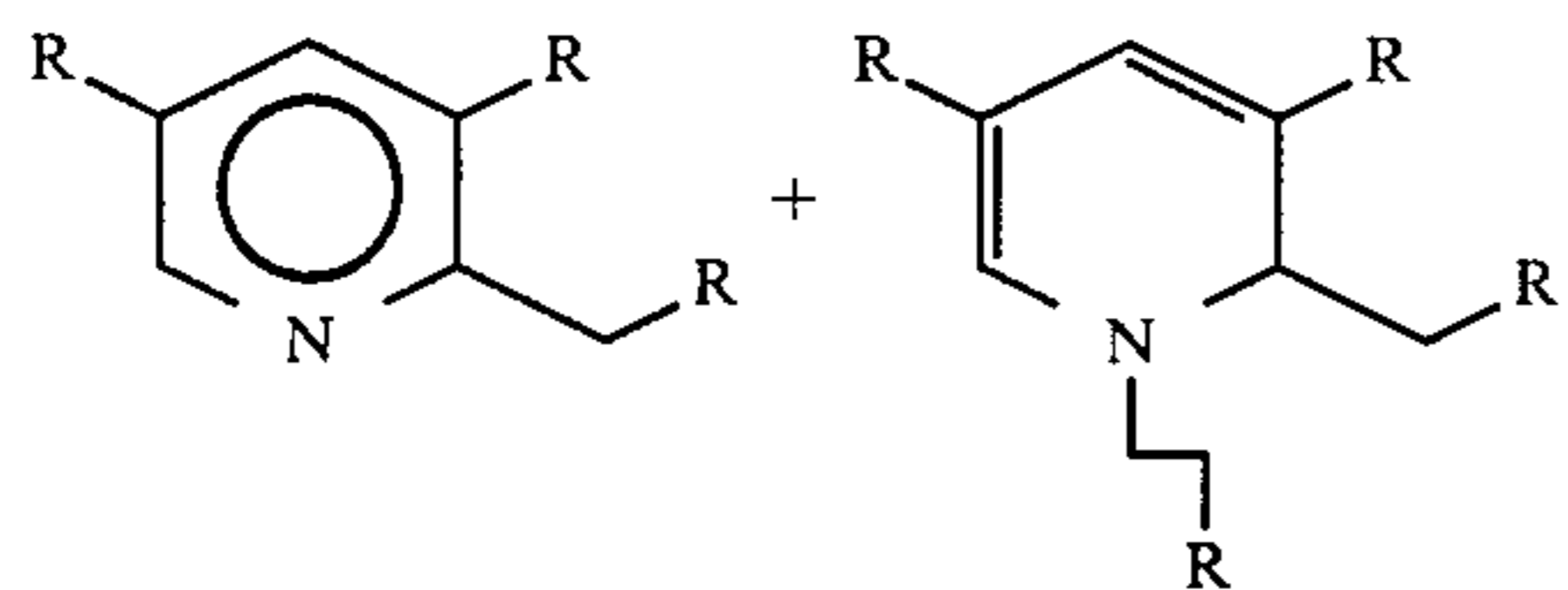
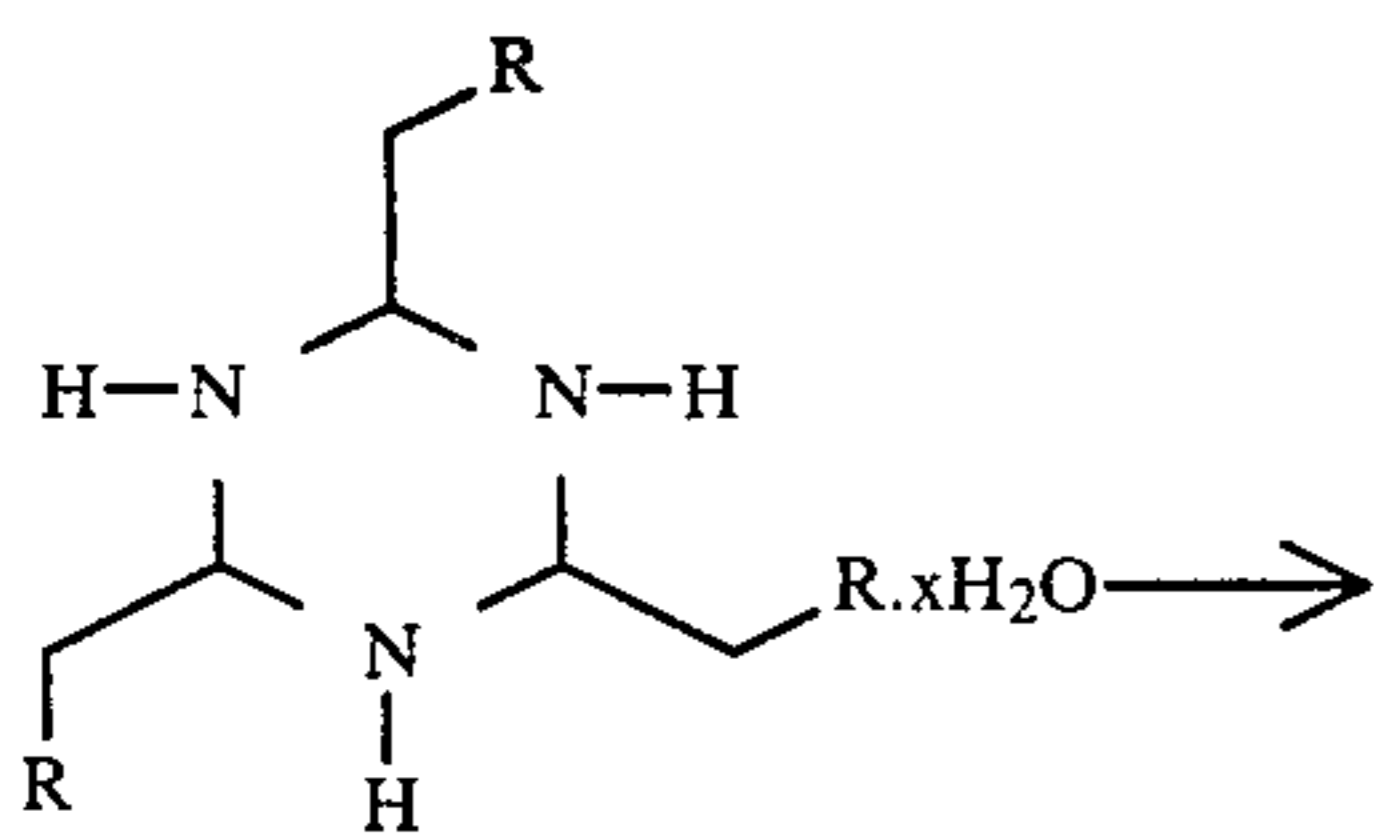
where R is the R moiety of the aldehyde or ketone reacted.

The following is an outline of the reactions occurring in the above examples.

EXAMPLE 1



2,4,6-trialkyl 1,3,5-hexahydrotriazine

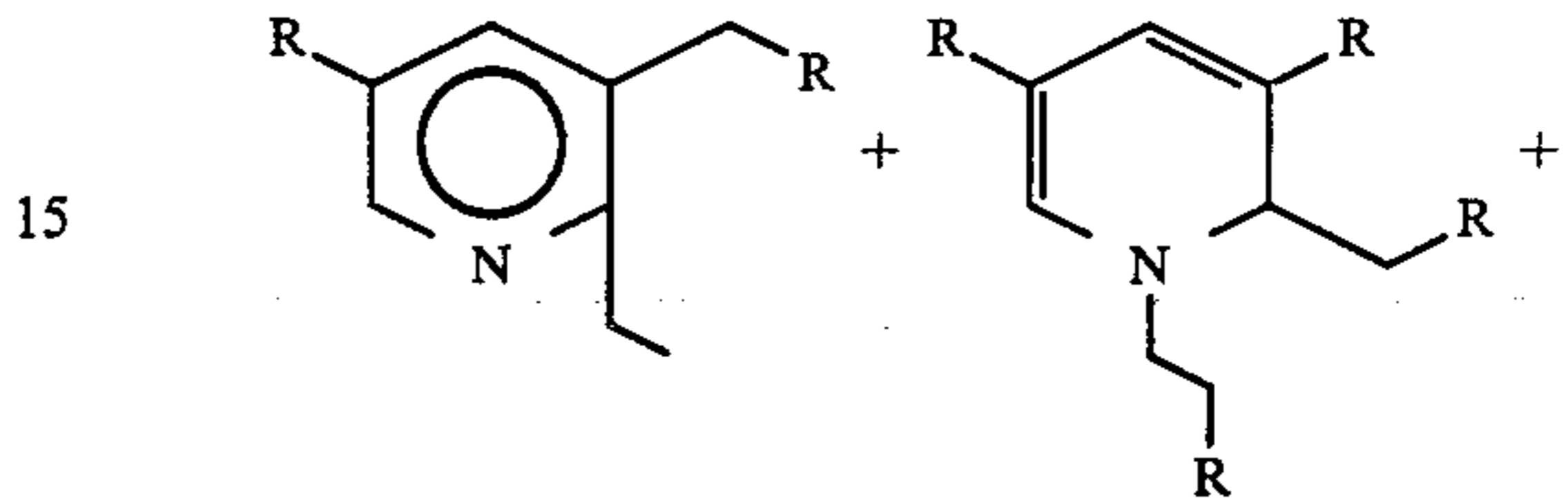
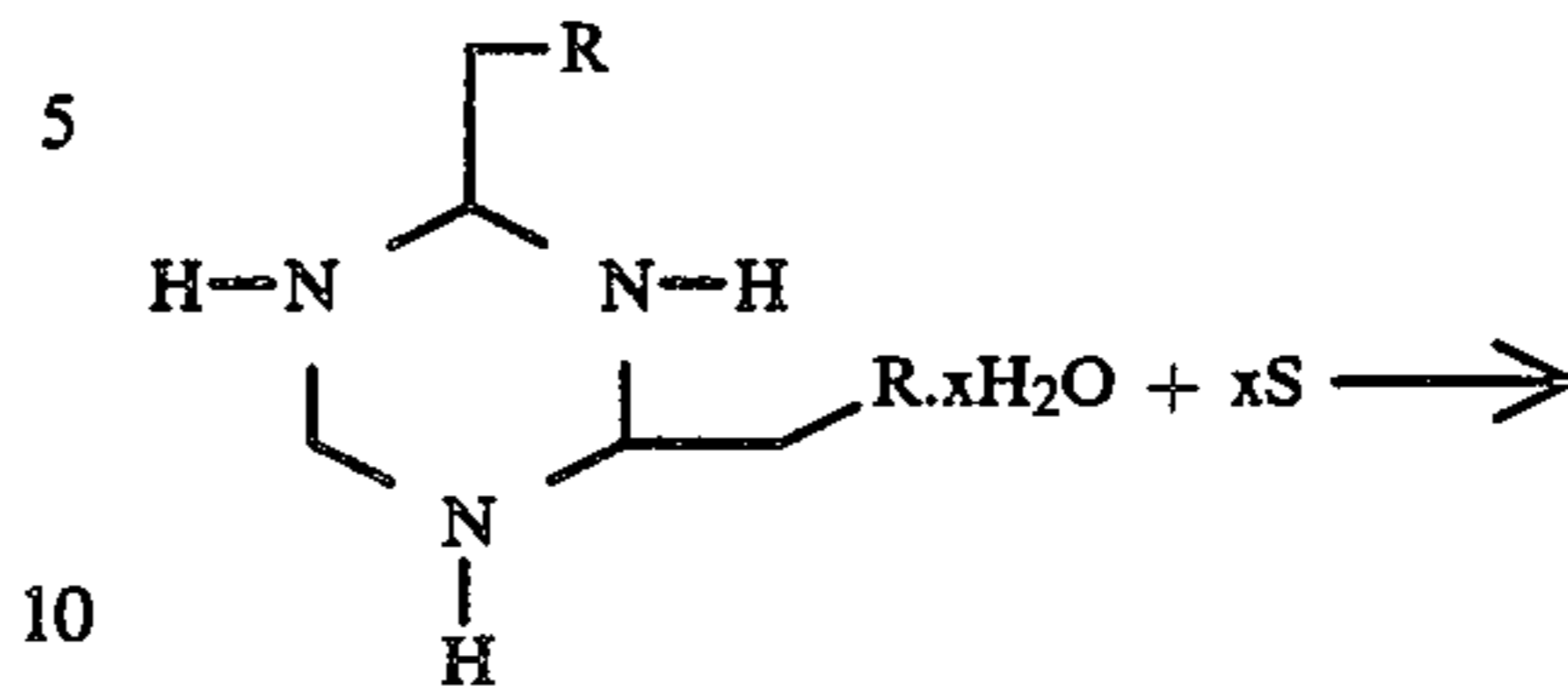


alkylpyridine

N-alkyldihydropyridine

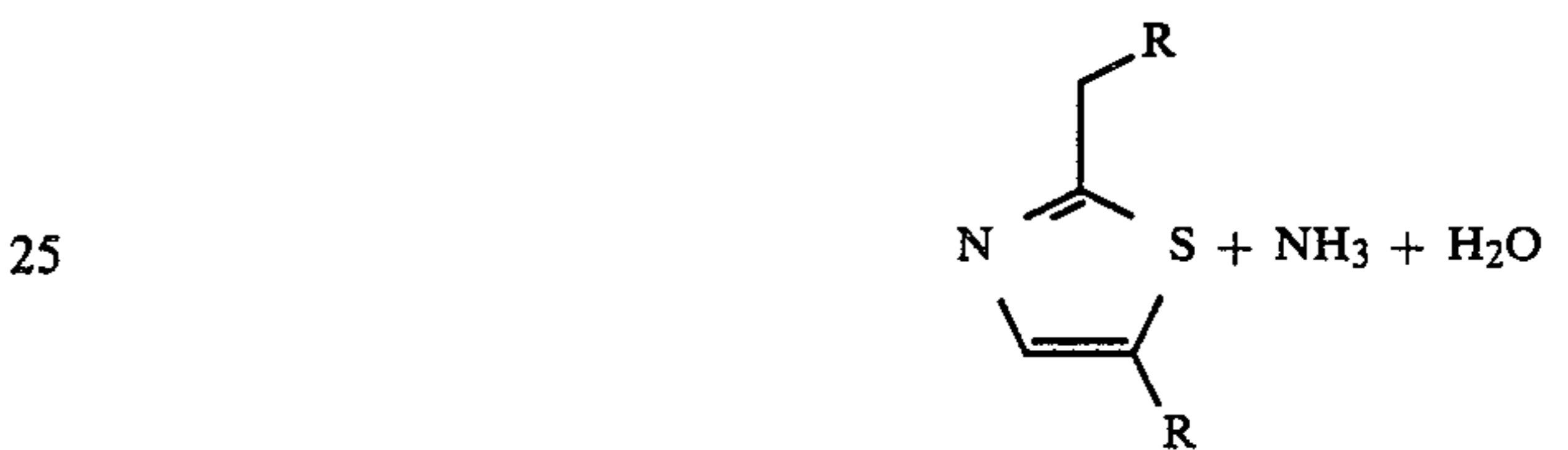
EXAMPLES 2-5

(2)



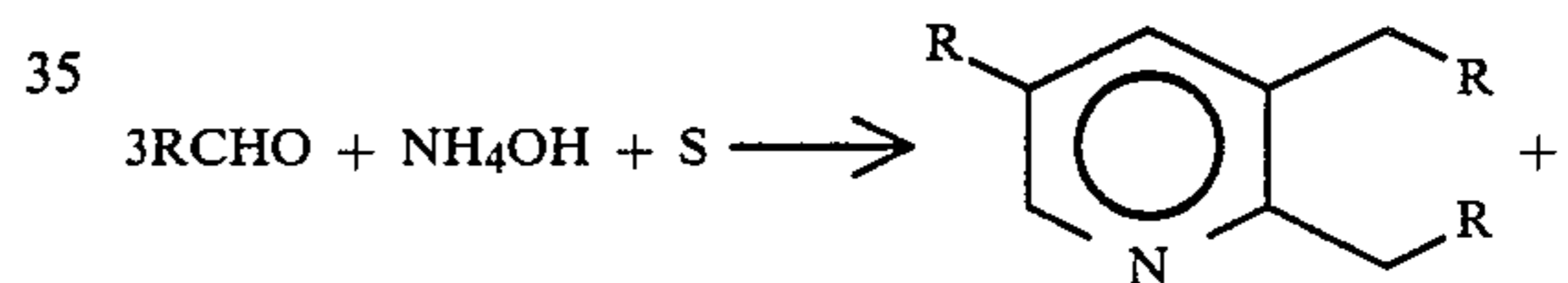
alkylpyridine

N-alkyldihydropyridine

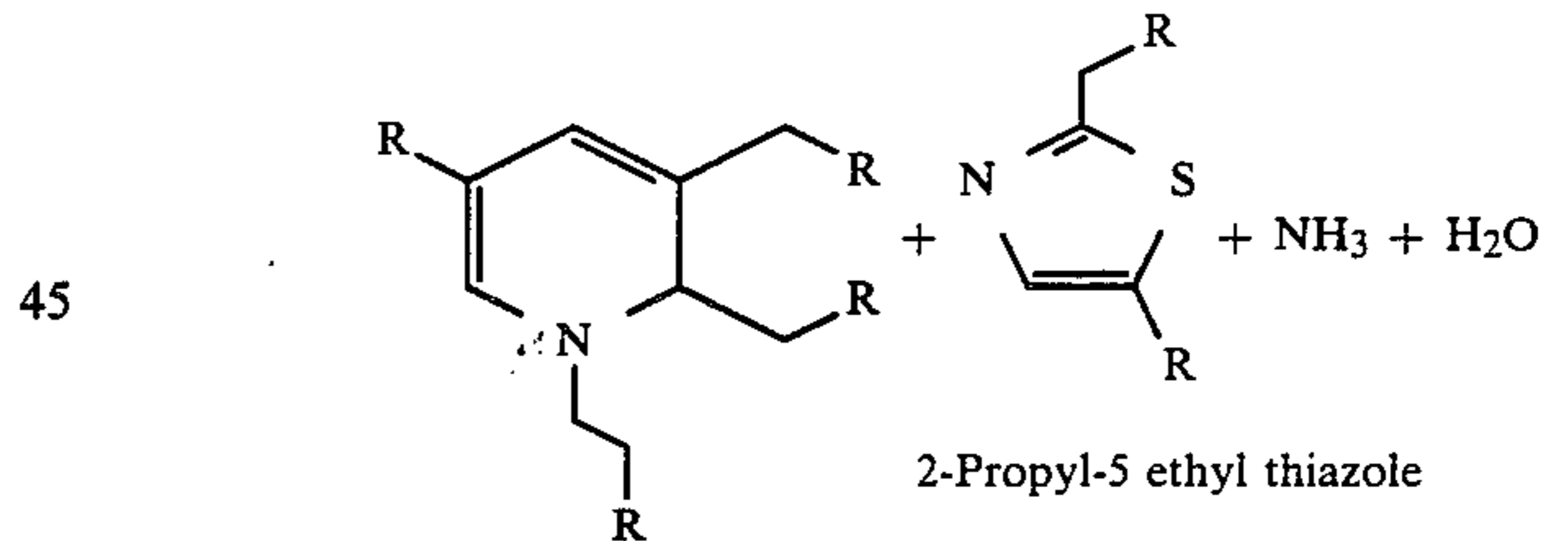


2-Propyl-5 ethyl thiazole

EXAMPLES 6-9

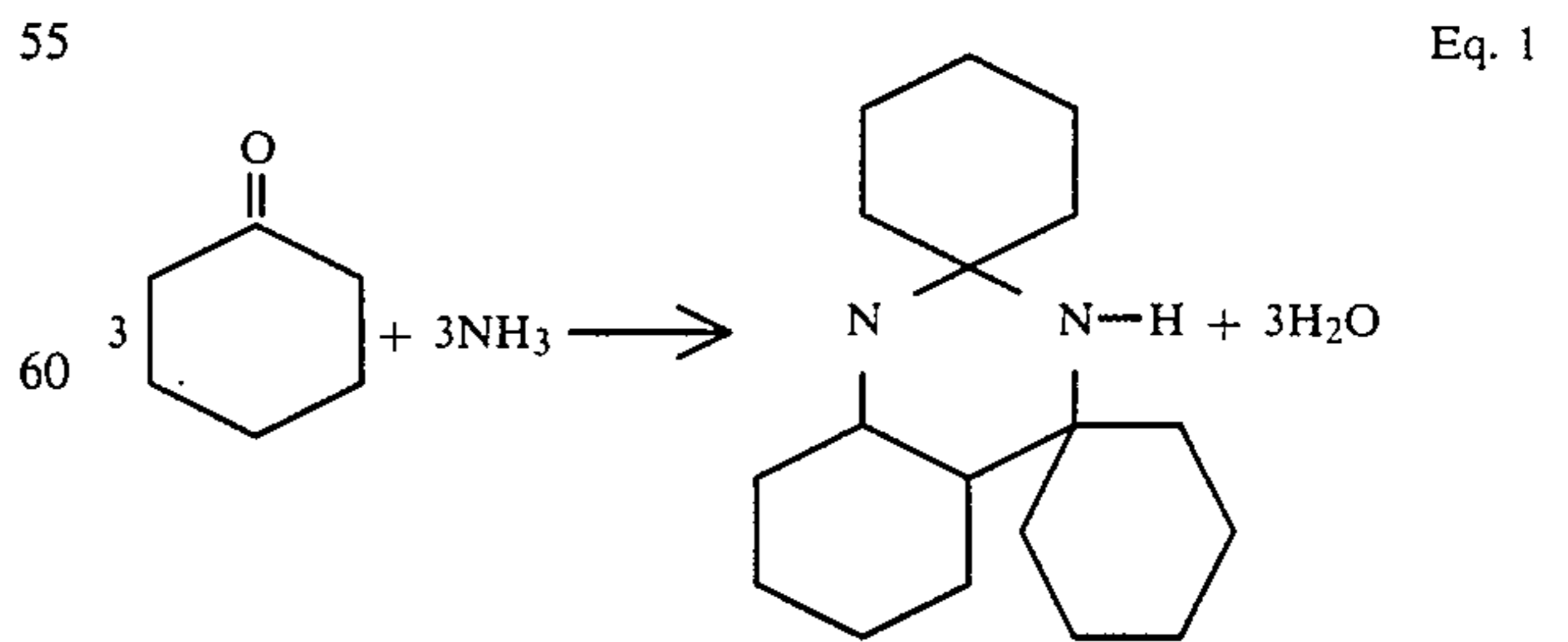


alkylpyridine



N-alkyldihydropyridine

EXAMPLE 10

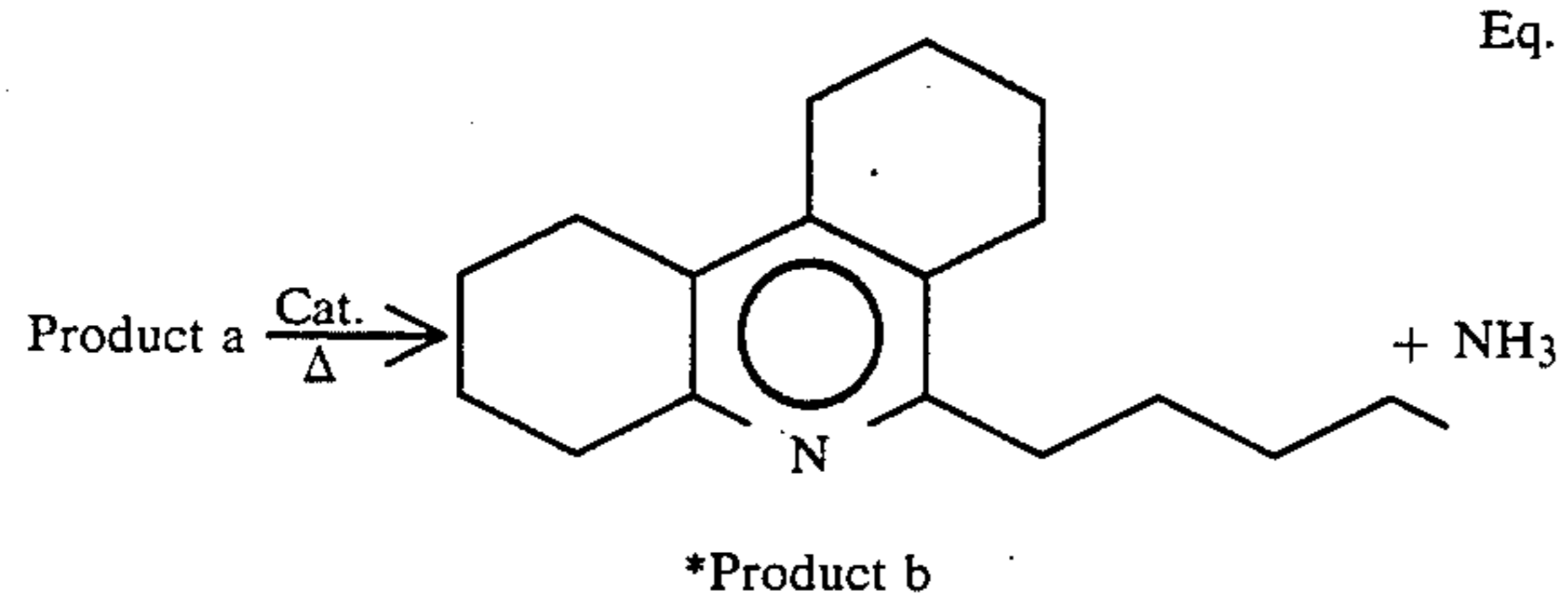


\*Product a  
\*2,2,4,4-dipentamethylene 5,6-tetramethylene  
2,2,4,4-tetrahydropyrimidine



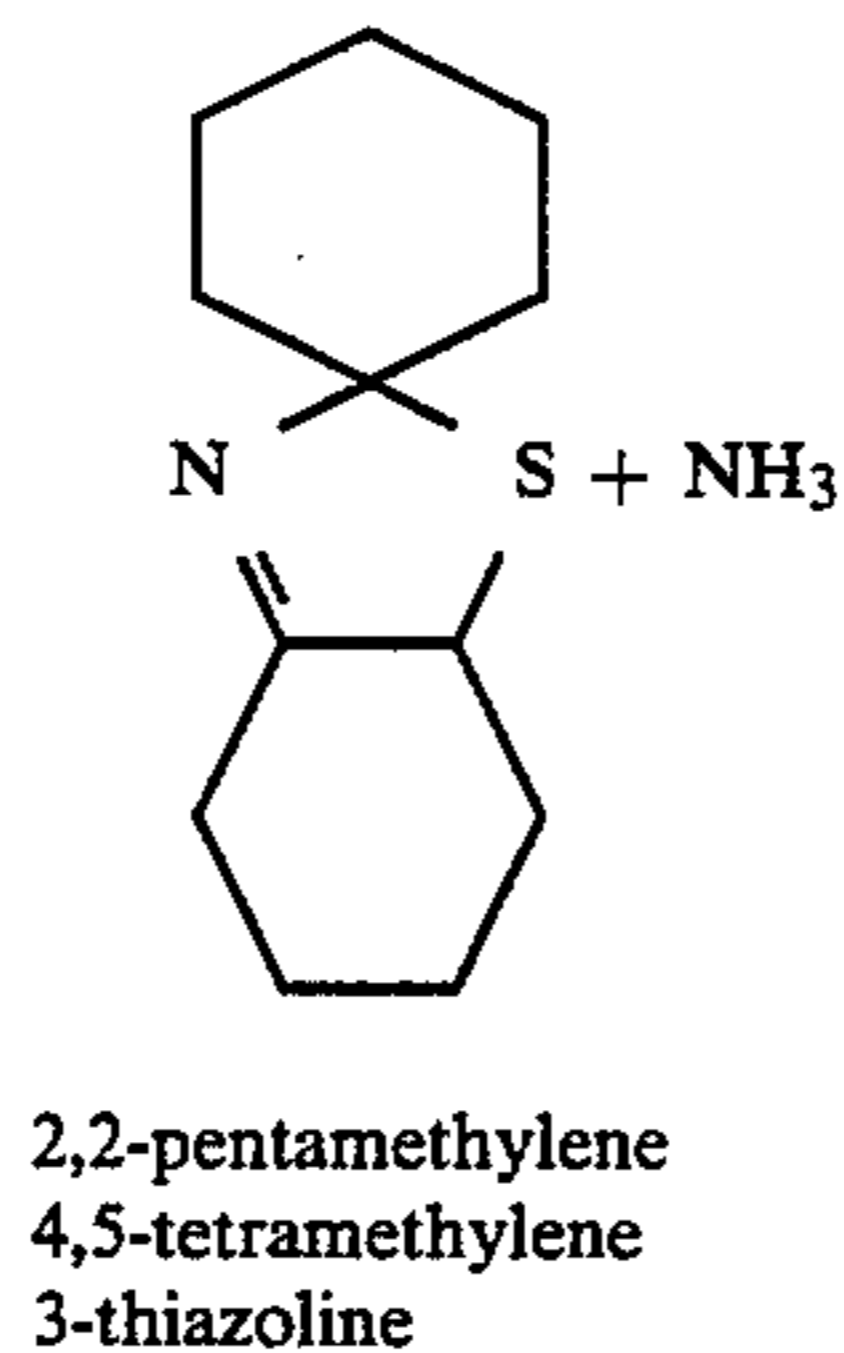
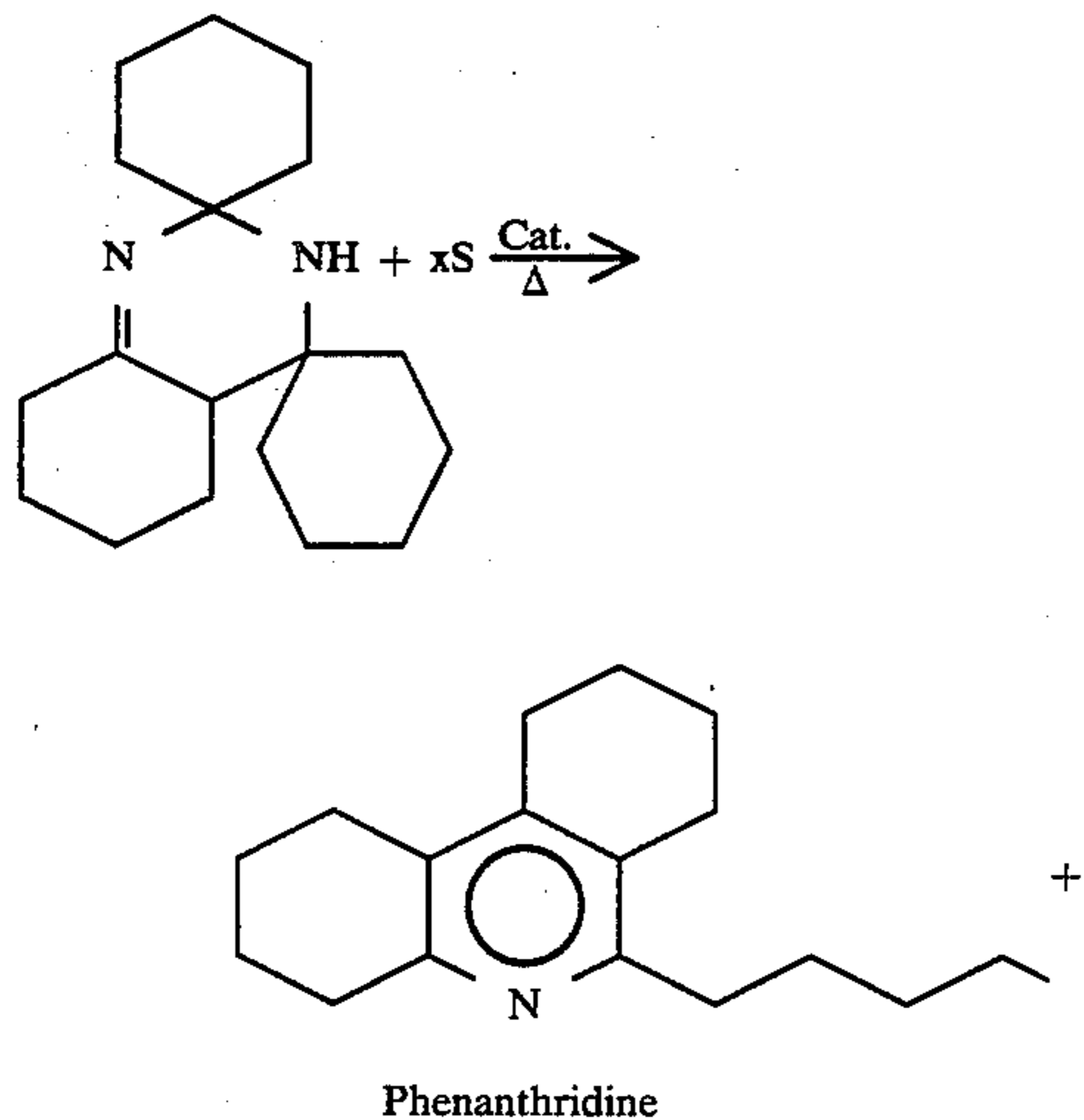
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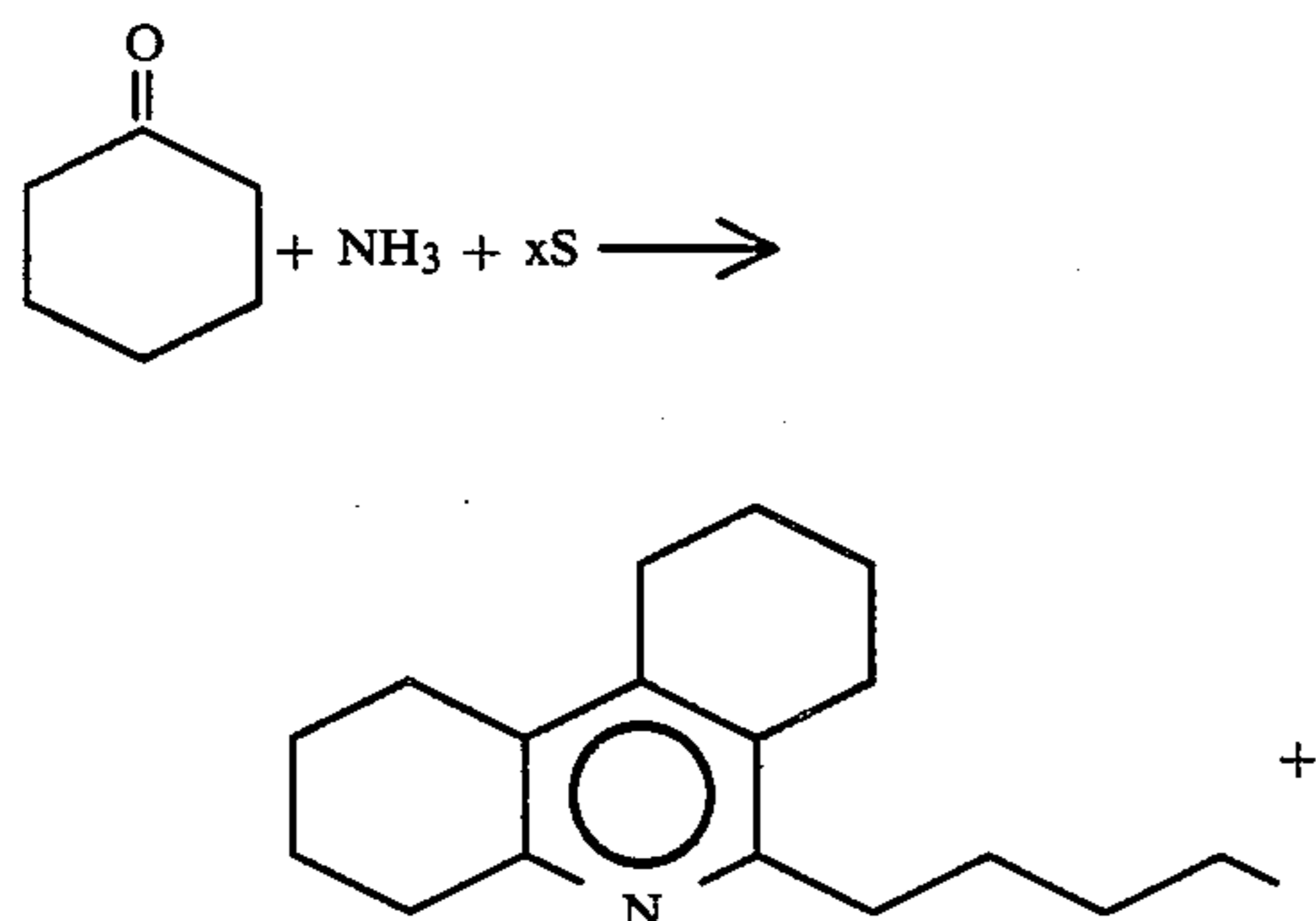


\*9-pentyl 1,2,3,4,5,6,7,8-octahydrophenanthridine

## EXAMPLES 11-13

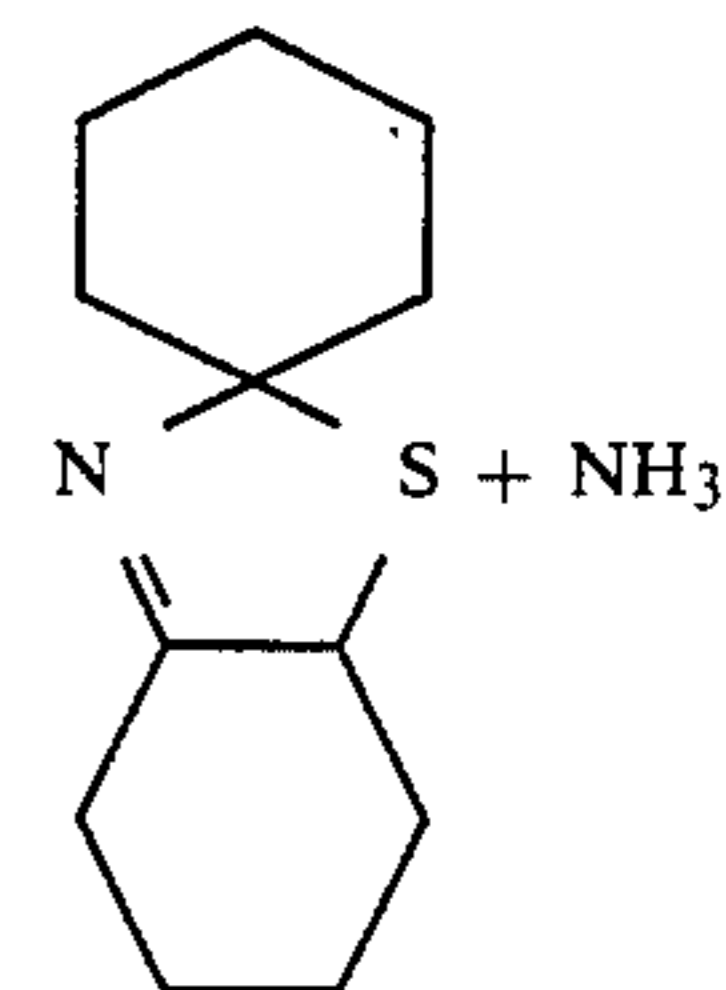


## EXAMPLES 14-17



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The following examples illustrate the preparation of the corrosion inhibitor compositions of this invention.

## EXAMPLE 26

To 50 g. of the reaction products of Example 3 dissolved in 35 g. of 2-propanol and 35 g. of water was slowly added, with cooling and stirring, 20 g. of concentrated sulfuric acid. The product was a clear, bright liquid. To this liquid was added 20 g. of nonyl phenol that had been condensed with 15 mols of ethylene oxide.

## EXAMPLE 27

To 50 g. of the reaction products of Example 4 dissolved in 33.5 g. of 2-propanol and 33.5 g. of water was slowly added, with cooling and stirring, 17 g. of 70% nitric acid. The product was a clear, bright liquid. To this liquid was added 17 g. of nonyl phenol that has been condensed with 15 mols of ethylene oxide.

## EXAMPLE 28

To 50 g. of the reaction products of Example 2 dissolved in 40 g. of 2-propanol and 22 g. of water was slowly added, with cooling and stirring, 12 g. of 85% phosphoric acid. To the resulting liquid was added 13 g. of hexadecyl amine that had been condensed with 23 mols of ethylene oxide.

## EXAMPLE 29

To 54 g. of the reaction products of Example 12 dissolved in 37 g. of 2-propanol and 37 g. of water was slowly added, with cooling and stirring, 20 g. of concentrated sulfuric acid. To the resulting liquid was added 13 g. of nonyl phenol that had been condensed with 15 mols. of ethylene oxide.

In a similar manner the following corrosion inhibitor compositions were prepared.

## EXAMPLE 30

Products of Ex. 12—54 g.  
Nitric acid (70%)—17 g.  
Dinonyl phenol condensed with 17 mols ethylene oxide—13 g.  
2-propanol—35.5 g.  
Water—35.5 g.

## EXAMPLE 31

Products of Ex. 14—54 g.  
Sulfuric acid (conc.)—20 g.  
Nonyl phenol condensed with 15 mols ethylene oxide—14 g.  
2-propanol—37 g.  
Water—37 g.

## EXAMPLE 32

Products of Ex. 10—54 g.

Nitric acid (70%)—17 g.  
 Nonyl phenol condensed with 10 mols ethylene oxide—14 g.  
 2-propanol—35.5 g.  
 Water—35.5 g.

EXAMPLE 33

Products of Ex. 12—54 g.  
 Phosphoric acid (85%)—12 g.  
 nonyl phenol condensed with 15 mols ethylene oxide—13 g.  
 2-propanol—66 g.

EXAMPLE 34

Products of Ex. 11— $\sqrt[3]{g}$ .  
 Glacial acetic acid—17 g.  
 oleyl amine condensed with 22 mols ethylene oxide—14 g.  
 2-propanol—35.5 g.  
 Water—35.5 g.

EXAMPLE 35

Products of Ex. 19—52 g.  
 Sulfuric acid (conc.)—19 g.  
 nonyl phenol condensed with 15 mols ethylene oxide—12 g.  
 2-propanol—35.5 g.  
 Water—35.5 g.

The following tests were conducted to demonstrate the increased effectiveness of the acid cleaning inhibitor compositions of this invention over the nitrogen and sulfur-nitrogen and heterocycles described in patent application Ser. No. 101,953.

TABLE I

Temperature	175° F.
Duration of test	6 hours
Corrodent	8% Formic Acid
Inhibitor Concentration	0.2%
Coupons	AISI 1020 mild steel
Volume/Surface ratio	25 ml./sq. in.
Product of Example	Corrosion rate (lbs./ft <sup>2</sup> /24 hrs.)
3 (Composition of S.N. 101,953)	0.009
20 (Composition of S.N. 101,953)	0.010
26 (Composition of this invention)	0.0035
31 (Composition of this invention)	0.0020
Blank	0.070

The following tests results illustrate the effectiveness of the compositions of this invention as corrosion inhibitors in acid cleaning solvents. The test procedure utilized was the same as before.

TABLE II

Temperature	200° F.
Duration of test	16 hours
Corrodent	4% Hydroxyacetic acid and 2% Formic acid mixture
Inhibitor concentration	0.2%
Coupons	AISI 1020 mild steel
Volume/Surface ratio	25 ml./sq. in.
Product of Example	Corrosion rate (lbs./ft <sup>2</sup> /24 hrs.)
26	0.0075
27	0.0070
29	0.0027
30	0.0027

TABLE II-continued

33	0.0031
Blank	0.065

TABLE III

Temperature	175° F.
Duration of test	6 hours
Corrodent	10% formic acid
Inhibitor concentration	0.2%
Coupons	AISI 1020 mild steel
Volume/Surface ratio	25 ml./sq. in.
Product of Example	Corrosion rate (lbs./ft <sup>2</sup> /24 hrs.)
27	0.0085
28	0.0076
30	0.0020
31	0.0022
32	0.0023
Blank	0.060

I claim:

1. A corrosion inhibiting composition of matter comprising a mixture of non-halogen salts of nitrogen heterocyclics and nitrogen-sulfur heterocyclics where the nitrogen heterocyclics are selected from phenanthridines, tetrahydropyrimidines, pyridines, dihydropyridines and mixtures thereof and the sulfur-nitrogen heterocyclics are selected from thiazoles and thiazolones and mixtures thereof.

2. The composition of claim 1 where the phenanthridine is 9-pentyl 1,2,3,4,5,6,7,8-octahydrophenanthridines; where the alkyl pyridine is 2-propyl 3,5-diethyl pyridine and the dihydropyridine is 1-butyl, 2-propyl 3,5-diethyl 1,2-dihydropyridine; where the tetrahydropyrimidine is 2,2,4,4-dipentamethylene-5,6-tetramethylene-2,2,4,5-tetrahydropyrimidine; where the thiazolene is 2,2-pentamethylene-4,5-tetramethylene-3 thiazolene; where the thiazole is 2-propyl-5 ethyl thiazole.

3. The composition of claim 1 which also contains a surfactant.

4. The composition of claim 2 which also contains a surfactant.

5. The composition of claim 3 where the surfactant is also an oxyalkylate.

6. The composition of claim 4 which is also an oxyalkylate.

7. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 1.

8. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 2.

9. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 3.

10. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 4.

11. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 5.

12. A process of inhibiting corrosion in a system containing stainless steel which comprises treating such system with the composition of claim 6.

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