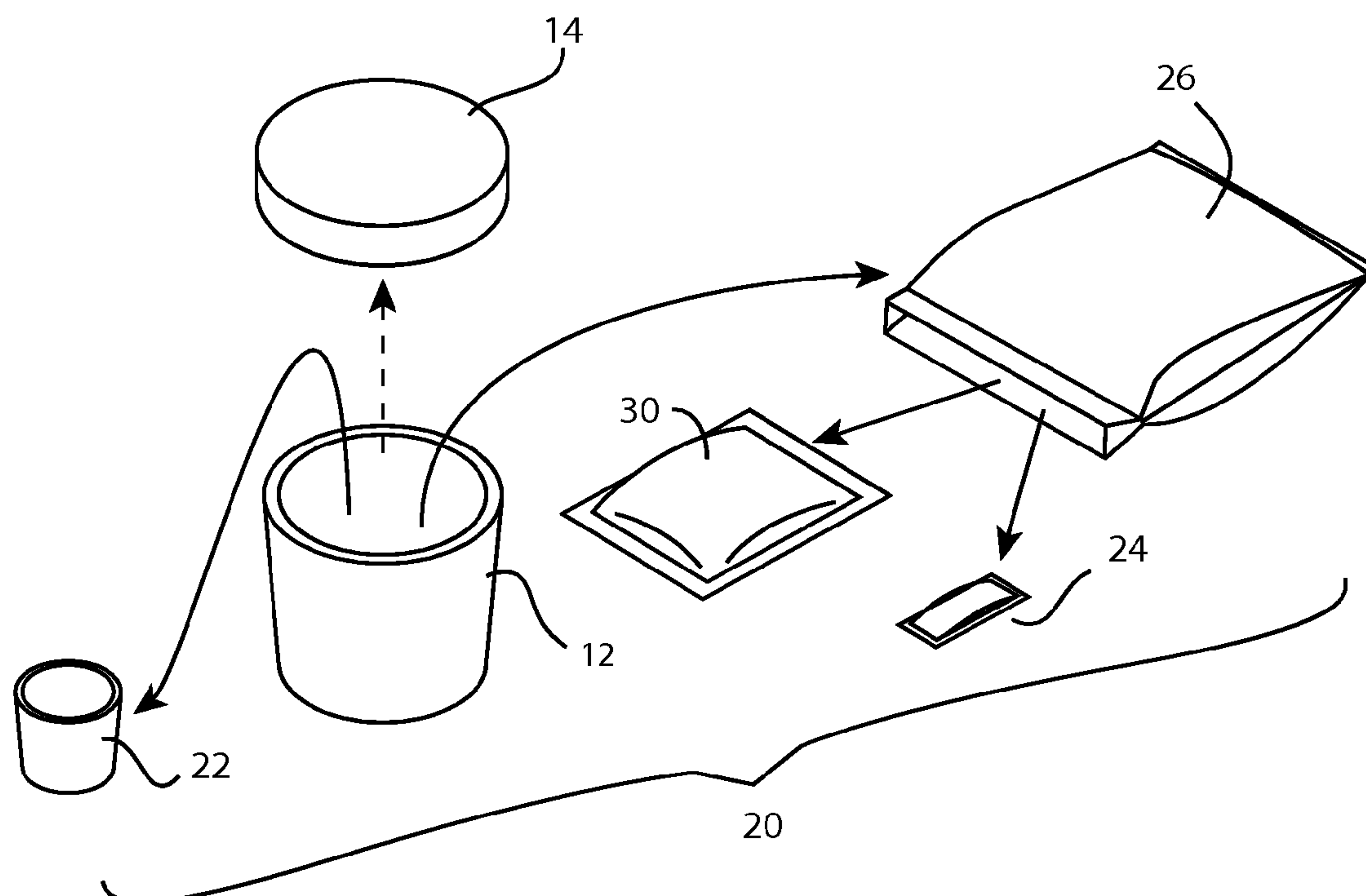


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(19) **United States**(12) **Patent Application Publication**
Baselli et al.(10) **Pub. No.: US 2013/0136685 A1**(43) **Pub. Date: May 30, 2013**(54) **PORTABLE CHLORINE DIOXIDE
GENERATOR**(71) Applicants: **Juan Carlos Baselli**, West Hills, CA
(US); **Spencer Blua**, Torrance, CA (US)(72) Inventors: **Juan Carlos Baselli**, West Hills, CA
(US); **Spencer Blua**, Torrance, CA (US)(21) Appl. No.: **13/684,638**(22) Filed: **Nov. 26, 2012****Related U.S. Application Data**(60) Provisional application No. 61/563,723, filed on Nov.
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(2013.01)USPC **423/477**; **422/164**(57) **ABSTRACT**

The present invention provides a safe, disposable and biodegradable chlorine dioxide micro generator that uses water soluble paper and hydrogel or compressed cellulose encased in filter paper pouch. The chemicals are kept in a stabilize form until activated by the addition of water. Multiple levels of protection against early exposure to water such as a foil pouch and an impermeable outer container allow for the safe transportation and storage in small, ready for deployment amounts of the chemicals. Wicking materials packaged around the chemicals provide for the ready introduction of water to the chemicals at the proper time. Water dissolves the paper walls of the chemical pack housings and then the water facilitates the reaction between the acid and the sodium chlorite to form chlorine dioxide gas as will be described further hereunder.



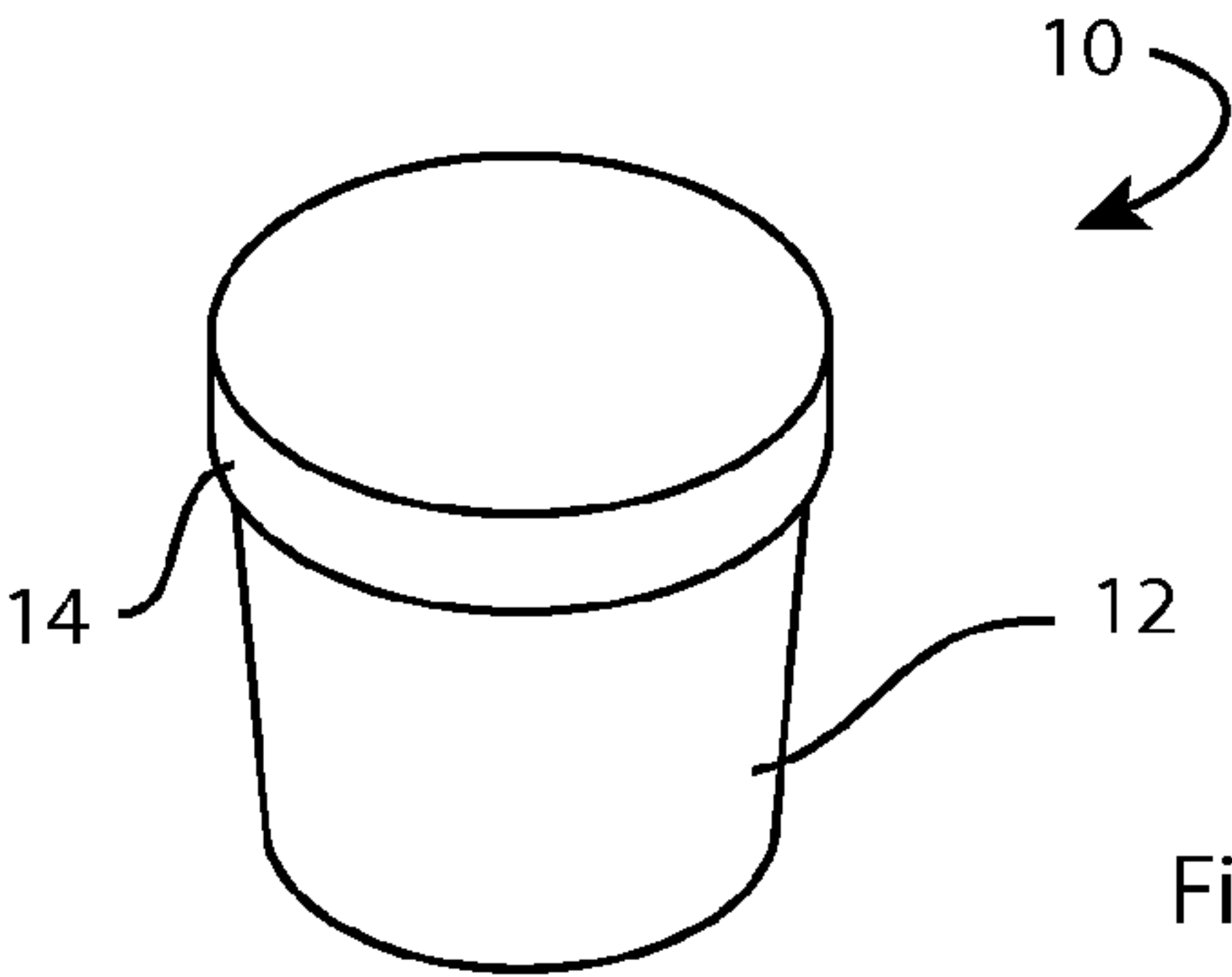


Figure 1

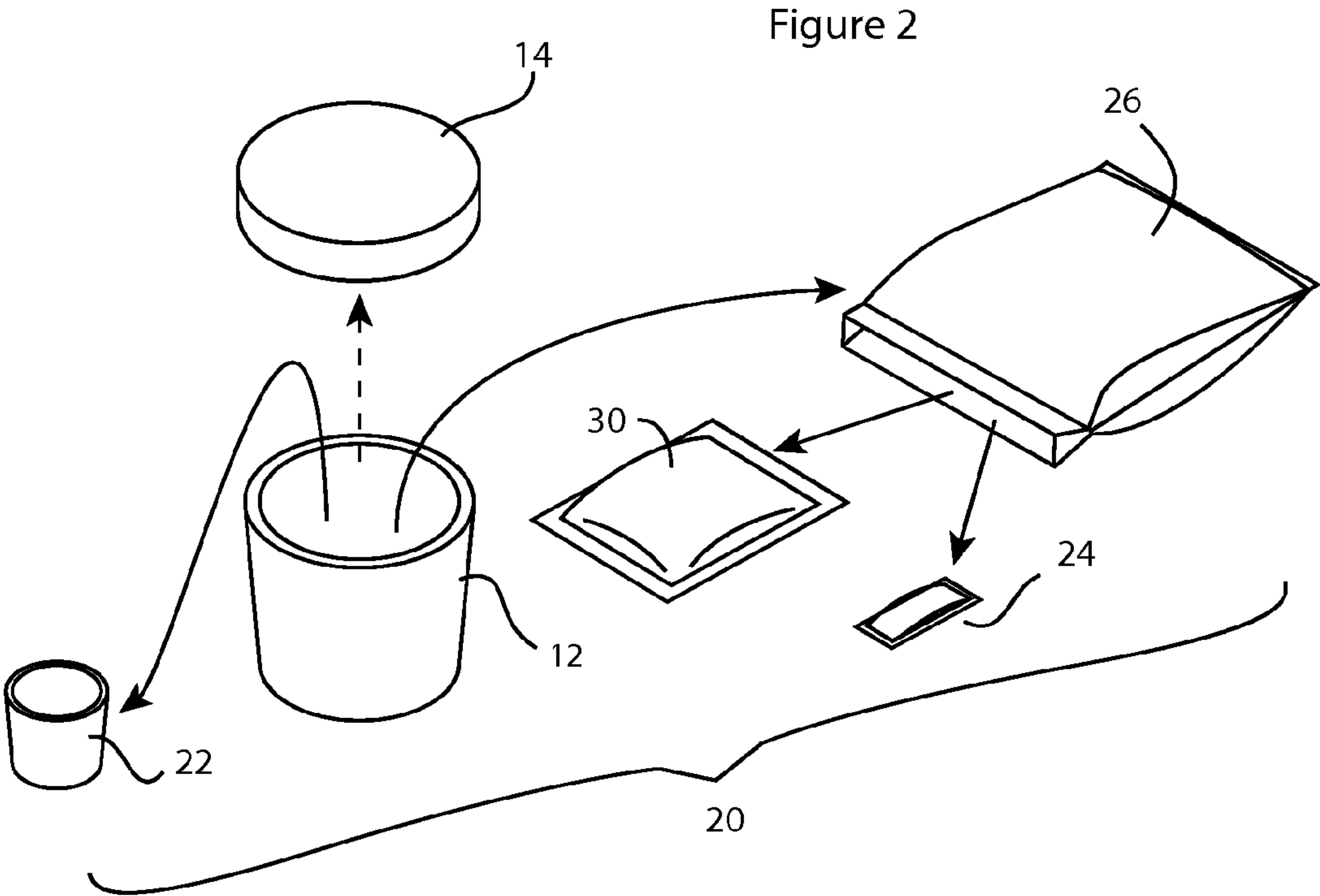


Figure 2

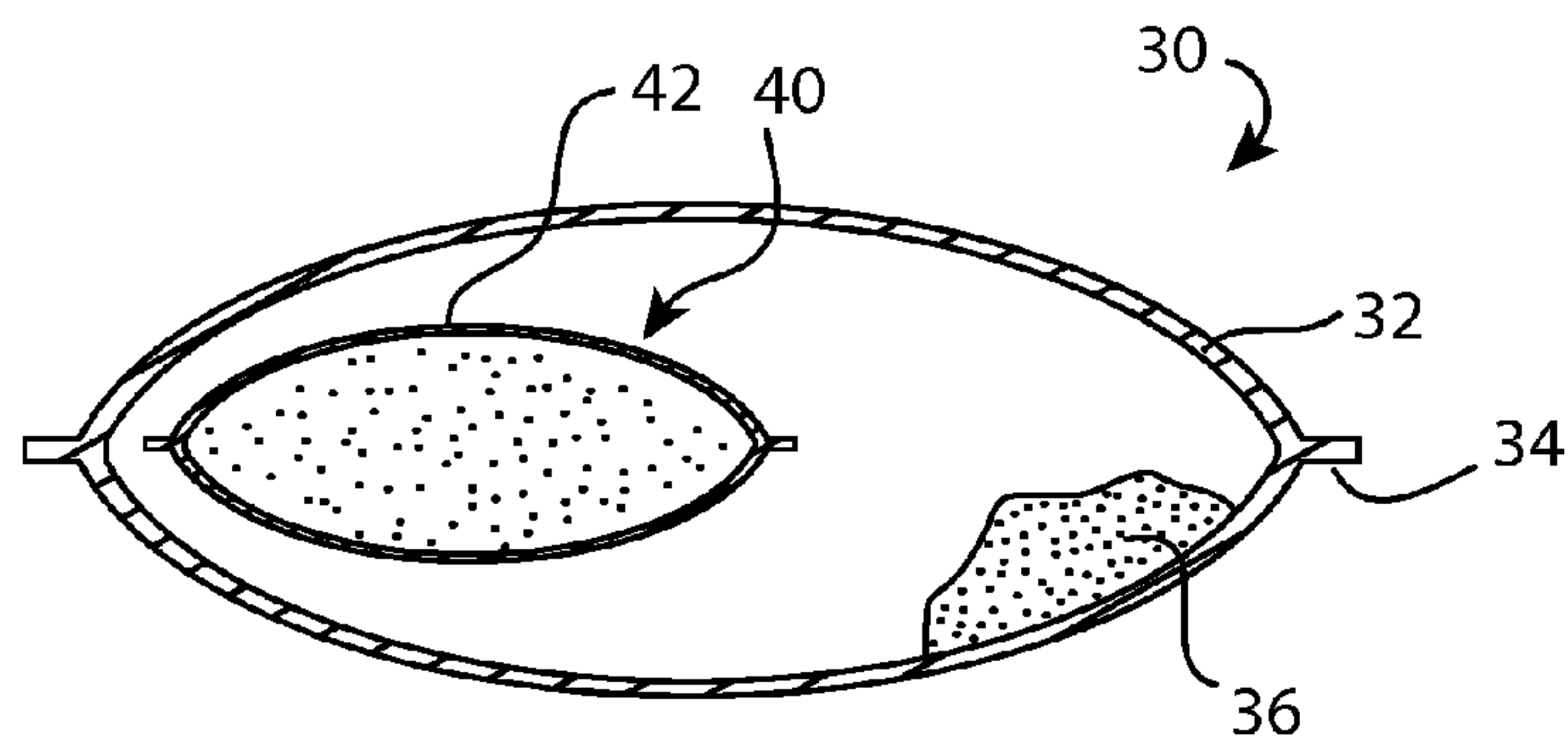


Figure 3

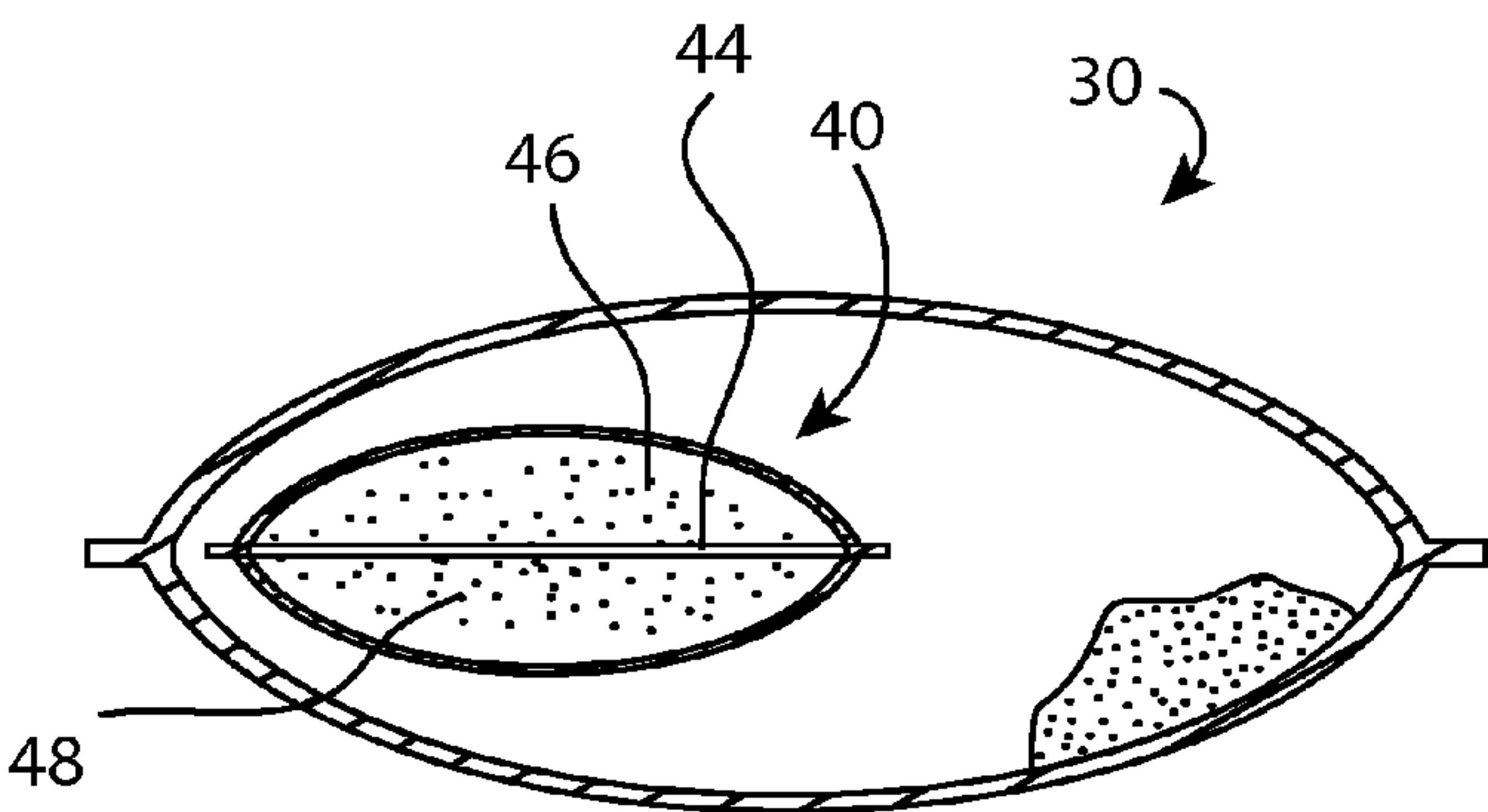


Figure 4

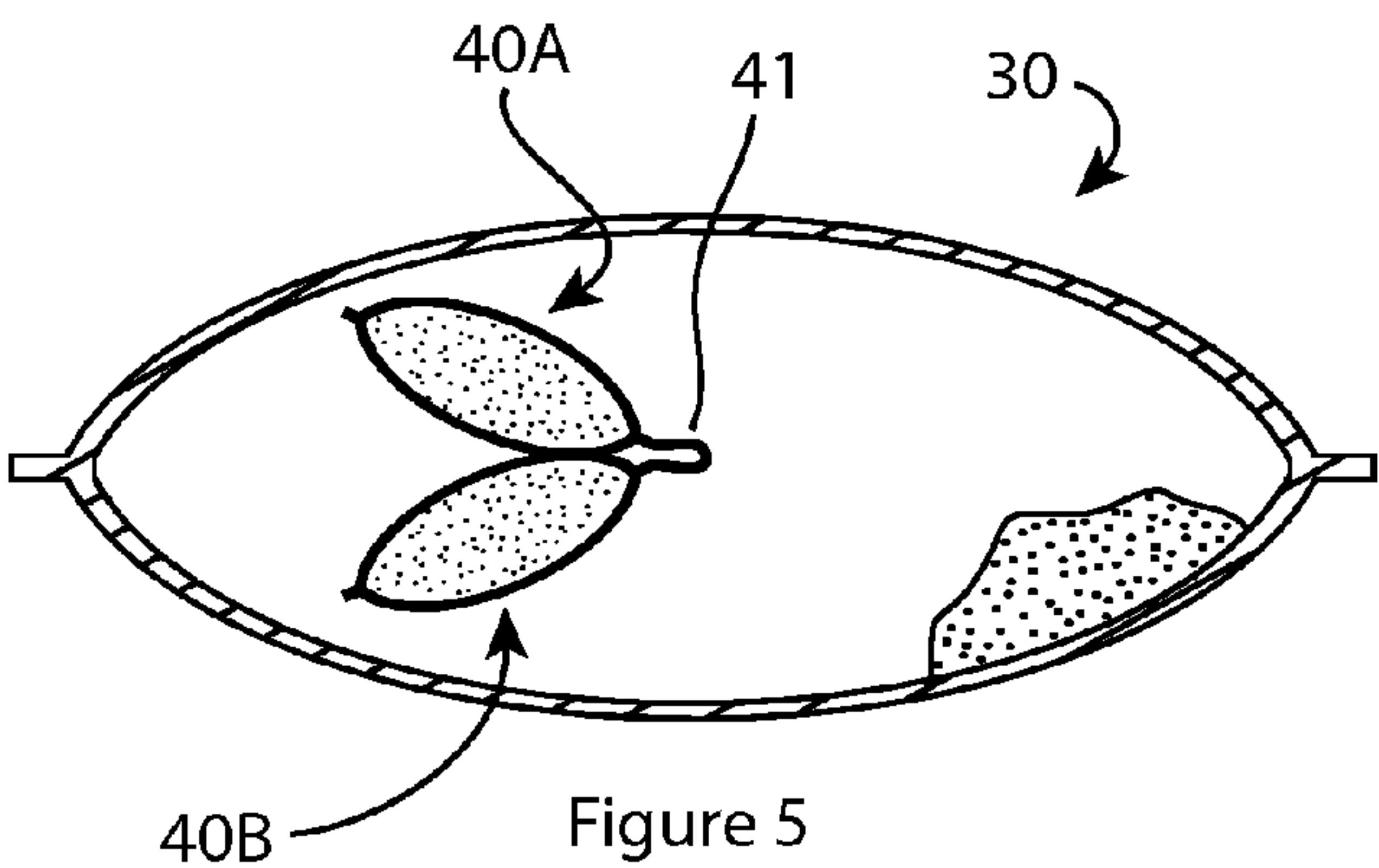


Figure 5

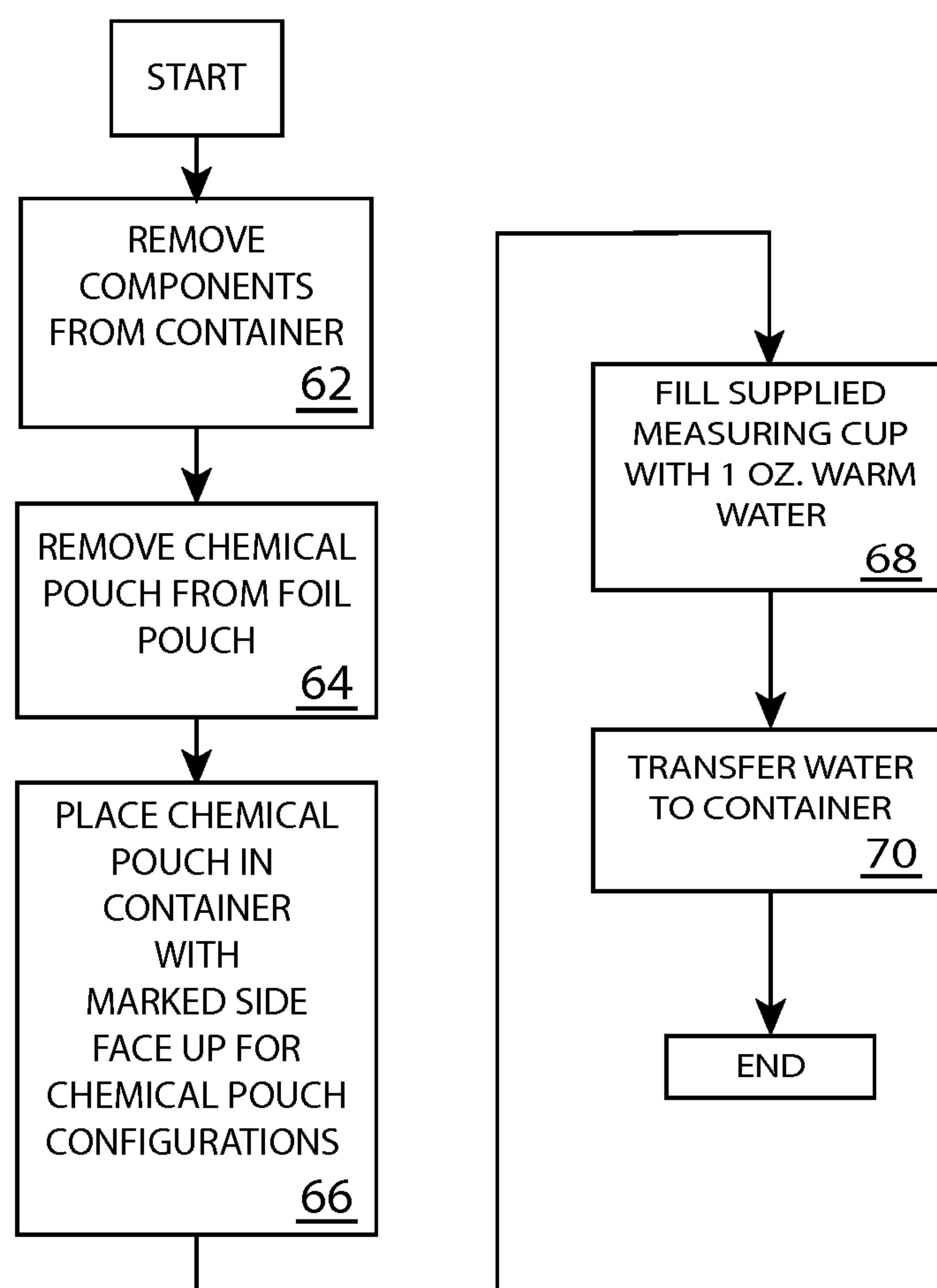


Figure 6

PORTABLE CHLORINE DIOXIDE GENERATOR

CROSS REFERENCE TO RELATED APPLICATIONS

[0001] This application claims the benefit of U.S. Provisional Application 61/563,723, filed Nov. 25, 2011, entitled Portable Chlorine Dioxide Generator, which is incorporated herein by reference.

BACKGROUND OF THE INVENTION

[0002] 1. Field of the Invention

[0003] The present application relates to a disposable and biodegradable chlorine dioxide micro generator from portable, stable chemicals, using for example water soluble paper and hydrogel or compressed cellulose encased in a filter paper pouch.

[0004] 2. Description of the Prior Art

[0005] Chlorine Dioxide (herein also referred to by "ClO₂" or "ClO₂") is a known biocide and disinfectant. It works by oxidizing single cell organisms in a known manner to kill the organism. Chlorine dioxide is currently used in commercial buildings to disinfect and deodorize various rooms and other enclosed areas. It has been known to be used in gymnasiums and other sports facilities to prevent staph infections while simultaneously deodorizing the facility. However, ClO₂ is an unstable chemical that breaks down especially in ultraviolet light and must be generated on site by large, bulky industrial equipment making it inaccessible to smaller sites at a reasonable cost.

[0006] Because of the inherent instability of chlorine dioxide, it is currently generated as needed. This is typically done by mixing a small amount of sodium chlorite and acids from large canister reservoirs. The sodium chloride is mixed with the acid, such as for example, Citric acid, sodium bisulfate, hydrochloric acid, etc. in large, industrial machinery. The separate canisters prevent unintended or premature mixing of the chemicals, but require porting around excess equipment to the desired site. It is therefore desired to provide a portable chlorine dioxide generator that can deploy small amounts of chlorine dioxide gas, while ensuring that the gas generation does not occur before the point of deployment. While the chlorine dioxide is not poisonous, it is at a minimum unpleasant or unhealthy to breathe, analogous to many household cleaners and thus premature mixing or "leakage" could have unwanted or deleterious consequences.

SUMMARY OF THE INVENTION

[0007] The present invention provides a safe, disposable and biodegradable chlorine dioxide micro generator that uses exposure to water to trigger a reaction between small quantities of provided chemicals, such as sodium chlorite and an acid to produce the chlorine dioxide. In one embodiment, water soluble paper and hydrogel or compressed cellulose encased in a filter paper pouch surround the chemicals and act to wick the water to the chemicals at the time of generation. The chemicals are kept in a dry, stabilized form until activated by the addition of water by multiple levels of protection. These levels include desiccants, physical separation, stabilizers, and impermeable barriers. These levels protect against early exposure to water to allow for the safe storage and

transportation of the chemicals in small, pre-measured amounts of the chemicals suitable for the intended application site.

[0008] Accordingly, it is a principal object of a preferred embodiment of the invention to provide a one time, single use chlorine dioxide generator that is safe to use, stable during storage and shipment, and is readily deployable.

[0009] It is another object of the invention to provide a stable environment for the sodium chlorite and acids to exist in a single package without prematurely forming chlorine dioxide.

[0010] It is a further object of the invention to provide packaging for the chemicals that in the absence of water acts to separate the chemicals, and during introduction of water to the packaging facilitates a reaction between the enclosed chemicals to form chlorine dioxide.

[0011] Still another object of the invention is to provide in at least one embodiment separate compartments for the chemicals to further forestall a premature or unintended reaction between the chemicals.

[0012] It is yet another object of the invention according to at least one embodiment to provide cellulose material to act as wick to rapidly introduce water to the chemicals at the time of reaction without interfering with the release of gas from the system.

[0013] It is an object of the invention to provide improved elements and arrangements thereof in an apparatus for the purposes described which is inexpensive, dependable and fully effective in accomplishing its intended purposes.

[0014] These and other objects of the present invention will be readily apparent upon review of the following detailed description of the invention and the accompanying drawings. These objects of the present invention are not exhaustive and are not to be construed as limiting the scope of the claimed invention. Further, it must be understood that no one embodiment of the present invention need include all of the aforementioned objects of the present invention. Rather, a given embodiment may include one or none of the aforementioned objects. Accordingly, these objects are not to be used to limit the scope of the claims of the present invention.

BRIEF DESCRIPTION OF THE DRAWINGS

[0015] FIG. 1 is an environmental perspective view of a container housing the components of the system.

[0016] FIG. 2 is an exploded view of the components of the system.

[0017] FIG. 3 is a break-away view of the chemical pouch and outer housing according to at least a first embodiment of the invention.

[0018] FIG. 4 is a break-away view of the chemical pouch and outer housing according to at least a second embodiment of the invention.

[0019] FIG. 5 is a break-away view of the chemical pouch and outer housing according to at least a third embodiment of the invention.

[0020] FIG. 6 is a flow diagram showing the steps for implementing the system.

[0021] Similar reference characters denote corresponding features consistently throughout the attached drawings.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENT(S)

[0022] The present invention is to a single use, compact chlorine dioxide generator. As best shown with reference to the drawings, the system and method for generating chlorine dioxide is shown. The system includes a container **10** (FIG. 1) having a lower cup portion **12** and lid **14** for containing all of the parts of the system and for preventing the introduction of any moisture to the system. As shown in FIG. 2, removal of the lid **14** from the cup **12** allows access to and removal of the optional components of the system **20**. Within the cup **12** are initially contained a water measuring cup **22**, and a foil pouch **26** containing a chemical pack housing **30** and a desiccant pack **24**.

[0023] In general, the materials are removed from the cup and the chemical pack housing **30** is removed from the foil pouch. A measured amount of water and the chemical pack housing are placed in the cup in the proper orientation, preferably with the chemical pack placed in the cup prior to the water. The water eventually is brought into contact with the chemicals within the chemical pouch **40** and facilitates the reaction between the acid and the sodium chlorite to form chlorine dioxide as will be described further hereunder.

The Components

[0024] The cup **12** and lid **14** that comprise the container are preferably made of an impermeable material, such as paper, plastic, etc. In the most preferred form, the cup is made from an impermeable paper so that the elements of the system may be readily be recycled or biodegraded. The container may be an important layer in preventing premature introduction of any water to the chemicals, but in some circumstances may be optional.

[0025] A foil pouch **26** is provided to house the chemical pack housing **30** during storage and shipment and is used to redundantly protect the chemicals from the premature introduction of moisture or liquids to the chemicals to prevent an unintended reaction of the chemicals with each other. Foil or other impermeable materials can be used on the walls of the pouch **26** to prevent moisture in or out of the pouch.

[0026] A desiccant **24** is provided within the foil pouch **26** to prevent moisture from accumulating within the foil pouch **26** during shipping and storage. One skilled in the art would understand that the use of a desiccant is well known for lowering the humidity of a closed compartment and that its use or an alternative humidity lowering device is not necessary for the operation of the invention, but is merely provided to maintain a stable environment. The heart of the system, however, is the chemical pack **30** housing the chemical pouch **40** with the active ingredients necessary for generating chlorine dioxide.

Chemical Pack Construction

[0027] The chemical pack housing **30** may have many configurations, several of which are described herein. The configuration used in a particular system **20** may vary based on the measure of chemicals used in order to achieve different levels of ClO₂ concentrations and with the size of confined environment to be treated.

[0028] With reference to FIG. 3, the chemical pack housing **30** has an outer filter paper (or other porous or mesh material) wall **32**, preferably joined together along a seam **34** joining the halves of the filter paper wall **32** together. Alternatively,

any paper wall of the system could also use other materials other than water soluble paper. For example water soluble film could be used, or even non-dissolvable filtration material could be used as long as water penetrates to allow the desired reaction to occur.

[0029] Within the filter paper outer wall **34**, the chemical pack housing is stuffed with a wicking material such as hydro gel or compressed cellulose **36** or other open cell foam/material/synthetic material, preferably having a high wicking capability. More preferably the cellulose takes the form of a sponge **36** as will be described further hereunder. The wicking material serves several functions. First, the material **36** serves as a physical buffer between the outer wall and the chemical pouch **40** within the outer wall **32**. The material isolates and protects the chemical pouch **40** from jostling and from damage during shipment. And in operation, the wicking material **40** also acts to rapidly wick the water from outside the chemical pack housing **30** into contact with the chemical pouch **40** to introduce the water to the chemicals within the pouch **40** to cause the chemical reaction.

[0030] In a preferred embodiment, the cellulose material of the chemical pack housing **30** is comprised of a number of smaller cellulose blocks **36**. Each block may be made of a dehydrated, compressed natural sponge or other cellulose source. By cutting a compressed sponge into approximately $\frac{2}{16}$ " by $\frac{2}{16}$ " cubes, it has been found that the wicking properties of the sponge increases to optimum levels to wick water rapidly to the chemical pouch **40**, while allowing for sufficient pathways for the gas created in the chemical reaction to escape between the cellulose cubes better than a single layer or layers of cellulose material.

[0031] The chemical pouch **40** according to at least one preferred embodiment contains both the sodium chlorite and the activating acid in a single structure. To prevent contact and/or an unintended reaction between the chemicals, a stabilizer is provided with and between the chemicals to forestall reaction. A preferred stabilizer is talcum powder ("talc"), but other stabilizers could also be used such as calcium chloride. One reason that the talcum powder may work well as a stabilizer is that it coats the granular surfaces of the chemicals to create a physical boundary between the chemicals. The talcum powder may also lower humidity reaching the chemicals to further prevent an unintended, premature reaction between the chemicals.

[0032] The stabilizer allows both of the chemicals to be safely stored within the same compartment. Housing both of the chemicals within the same housing has the benefit that once water is introduced to initiate chemical reaction, the acid and sodium chlorite are in close contact encouraging a full and complete reaction with low barriers to the escaping chlorine dioxide gas. The stabilizer also has the added advantage of providing a long shelf life for the product in the range of two years, well beyond what would be expected for this type of chlorine dioxide generator.

[0033] Depending on the size/amount of the chemicals and the intended environment that the chemicals will be used in, it may be desirable to store the chemicals in separate compartments to further ensure the chemicals cannot come in contact with each other prior to deployment at the intended site. As shown in FIG. 4, an additional paper barrier **44** can be introduced between the chemicals to keep the chemicals further separated. In this embodiment, it is preferred that the acid, such as citric acid be stored in the top compartment **46** and the sodium chlorite in the bottom compartment **48**.

Instructions or indicia on the chemical pouch **40** and/or the chemical pack housing **30** may be provided to ensure that the chemicals are deployed in the proper orientation. Having the acids on top of the sodium chlorite as the paper barrier dissolves during introduction of water to the system will ensure the most robust reaction between the acid and the sodium chlorite.

[0034] Another embodiment is shown in FIG. **5** having two separate pouches, one pouch **40A** for the acid and a pouch **40B** for the sodium chlorite. The chemicals pouch may be secured in close proximity by a tether **41** linking the pouches **40A** & **40B**. While shown in loose configuration in the figure, it may more practical to secure the pouches to each other at one or more point such that they maintain proper contiguous or close orientation to each other.

[0035] The chemical pack housing **30** itself may be fabricating with an outer pack wall **32** formed from filter paper filled with a measured amount of hydro gel (water absorbing polymer) or compressed cellulose **26**. The process of fabrication is preferably performed in a room with less than 35% humidity and chemicals holding less than 1% humidity, filling the chemical pouch **40** with compound made with 80% technical grade sodium chlorite, organic acid (i.e., Citric acid, sodium bisulfate, et.) and a stabilizer, then heat sealing the chemical pouch walls **40** together. The chemical pouch **40** is then placed within outer pack ("chemical pack housing") **30** with polymer and heat sealing outer pouch wall **24**. This is then placed in the foil pouch **26** along with desiccant **24** and the foil pouch is sealed to create a water tight housing. This foil pouch is then placed with the other components in the water impermeable container **10**, with lid sealed over the cup portion **12** to provide a long shelf life container for generating chlorine dioxide at the desired time in a small quantity.

Operation of the System

[0036] In operation, as shown in the block diagram of FIG. **6**, the self-contained chlorine dioxide generator is capable of creating the gas by the mere arrangement of the provided components plus the addition of a small amount of water. Referring to FIGS. **1**, **2** and **6**, operation of the system will be described.

[0037] To generate chlorine dioxide from the pre-packaged system, the lid **14** of the container **10** is separated from the cup portion **12**. The components, namely, the measuring cup **22** and foil pouch **26** are removed **62** from the container cup portion **12** and set aside

[0038] The chemical pack housing **30** is prepared by first removing **64** the pack **30** from the protective pouch **26**. The desiccant **24** (also in foil pouch **26**) is used merely to remove moisture from the interior of the container **10** and is no longer required, so may be discarded or otherwise disposed of or recycled.

[0039] The proper orientation of the chemical pack housing **30** is then determined to maintain the chemical pouch within the pack housing in the proper position. The chemical pack housing **30** is then placed **66** in the cup in the desired orientation.

[0040] With the chemical pouch properly placed in the cup **12**, the measuring cup **22** is then filled **68** with the required amount of water (not shown) using indications on the cup or in accordance with provided instructions that may be provided. Once the proper amount of water is measured **66**, it is poured into the cup **12** on top of the chemical pouch to initiate generation of the chlorine dioxide.

[0041] The water is first absorbed by the outer paper wall **32** of the chemical pack housing causing the paper walls to dissolve. The cellulose or hydro gel **36** then acts as a wick to direct the water at the desired rate to the active chemical pouch **40** (or **40A** & **B**). As the water is absorbed by the hydro gel or compressed cellulose **36** in the outer pouch **30**, it pushes the inner chemical pouches **40** upward against the outer pouch **30**, creating a pillow-like gel bed, where the chemical pouch walls **42** may dissolve and initiate the chemical reaction by mixing all the chemicals together.

[0042] The outer paper walls **42** continue to dissolve in the water allowing for more water to reach the chemicals and stabilizer within the chemical pouch. Where the chemicals are stored in separate compartments (**46,48** FIG. **4**) or separate pouches (**40A, 40B** FIG. **5**), interposed walls break down in the water and allow the chemicals to react and proceed analogously to the single pouch embodiment.

[0043] When sufficient water has permeated into the chemicals to overcome the stabilizer, the acid and sodium chlorite can react with each other to form a chlorine dioxide gas. The gas then percolates out through the cellulose and through gaps in the outer chemical pack housing walls **32** that have dissolved. The gas continues to expand and flow out of the cup **12** into the enclosed environment around the cup. The chlorine dioxide gas oxidizes or otherwise eliminates single celled organisms within the enclosed space around the cup, preferably killing any odor causing organisms. After a short period of time, the chlorine dioxide then begins to break up into salts and water. Since only a small amount of gas per volume of room is required to neutralize the organisms, the amount of salt and water should be negligible and should not create a need for separate cleanup of the resulting byproducts. In this way, the self-contained chlorine dioxide generator generates sufficient gas to deodorize a confined space and breaks up easily to into simple, environmentally friendly compounds. The articles left after the process, namely the chemical pack **30**, and active ingredients as well as the paper cup can be recycled or will biodegrade. And unlike the commercial gas generators, there is no industrial equipment or canisters left after the application of ClO_2 at the site to haul away.

Slow Release System

[0044] An alternative to adding water to the cup is to merely deploy the system as the chemical pack housing **30** within the cup and allow the ambient humidity of the surrounding environment to slowly allow the gradual introduction of water to the chemicals. Eventually as the humidity level of the interior of the chemical pack **30** and later chemical pouch **40** rise to the level of the room, the chemicals will slowly react to the water carried by the air to the chemical pouch **40**. The time that the humidity takes to reach the chemicals could be enhanced by scoring the walls of the chemical pack **30** or simply by removing the chemical pouch **40** from the chemical pack **30**, however this is not desirable because the chemical pack **30** acts as a fuse to delay the introduction of water/humidity to the chemicals and this may be defeated by removing the chemical pouch **40** from the pack. Additionally, the chemical pack housing **30** acts a physical boundary between the chemicals and the user to further safeguard the user from the chemicals.

Water Suspended Chlorine Dioxide

[0045] A similar process could also be used to create chlorine dioxide suspended in water by introducing the chemical

pack housing **30** to a larger volume of water. In such case, certain components of the system may be unnecessary, such as the cup **12** and measuring cup **22**. The chemical pack could be provided within the foil pack, and deployment could be as simple as dropping the chemical pack **30** (when removed from the foil pack) into a large volume of water to initiate generation of the gas. If the volume of water is contained in a closed tank of water, the ClO₂ will stay within the water for a longer period of time instead of being released from the water as a gas. A pump/sprayer on the tank could be used to spray ClO₂ containing water on a desired spot to remove odor or to kill single cell organisms or other affected materials. Alternatively, the pouch could be dropped into an open volume of water, and a mop or similar device could be used to apply the ClO₂ containing water.

[0046] The housing **30** could be provided within a webbing, netting or other housing configuration to prevent the housing from be introduced to downstream pumps, for example in a system that pumps the combined water and chlorine dioxide directly onto the desired areas. Additionally, individual components of the system can be made to be dissolvable or non-dissolvable as needed for various applications.

[0047] While this invention has been described as having a preferred design, it is understood that it is capable of further modifications, uses and/or adaptations of the invention following in general the principle of the invention and including such departures from the present disclosure as come within the known or customary practice in the art to which the invention pertains and as maybe applied to the central features hereinbefore set forth, and fall within the scope of the invention and the limits of the appended claims. It is therefore to be understood that the present invention is not limited to the sole embodiment described above, but encompasses any and all embodiments within the scope of the following claims.

We claim:

1. A system for generating chlorine dioxide, comprising:
 - a chemical pouch housing active chemical ingredients including sodium chlorite and an acid;
 - a stabilizer coating said sodium chlorite and said acid to allow for water-free storage of the sodium chlorite and acid in close contact without allowing a reaction between the sodium chlorite and the acid;
 - said chemical pouch having a non-water resistant outer wall for allowing water to pass through said wall into said active chemical ingredients;
 - wherein introduction of water to said active chemical ingredients activates said active chemical ingredients to form chlorine dioxide.
2. The system for generating chlorine dioxide of claim 1, wherein said active chemical ingredients in said system only include sodium chlorite and an acid.
3. The system for generating chlorine dioxide of claim 1, wherein said chemical pouch outer wall is made of dissolvable paper.
4. The system for generating chlorine dioxide of claim 1, wherein said chemical pouch is contained within a chemical pack outer housing, said chemical pack having a wicking material between at least one outer wall and said chemical pouch to wick water from an outer wall of said chemical pack to said chemical pouch outer wall.
5. The system for generating chlorine dioxide of claim 4, including a foil pouch for storing said chemical pack contain-

ing said chemical pouch in a water-free state to prevent water external to said foil pouch from prematurely reaching said active chemicals.

6. The system for generating chlorine dioxide of claim 5, including:

- a nonpermeable cup for storing said foil pouch, chemical pack and chemical pouch in a water-free state during storage; and
- a measuring cup for measuring an amount of water to be added to said active chemicals to form the chlorine dioxide at a preselected time of activation.

7. A method for generating chlorine dioxide, comprising: providing a chemical pouch housing a active chemical ingredients including sodium chlorite and an acid, said chemical pouch having a non-water resistant outer wall for allowing water to pass through said wall into said active chemical ingredients during activation of the active chemical ingredients;

providing a stabilizer within said chemical pouch for coating said sodium chlorite and said acid to allow for stable storage of the sodium chlorite and acid in close contact without allowing a reaction between the sodium chlorite and the acid prior to addition of a preselected amount of water; and

introducing water to said chemical pouch to cause water to breach through said chemical pouch outer wall and overcome said stabilizer to allow contact between said active ingredients to cause said active chemical ingredients to form chlorine dioxide.

8. The method for generating chlorine dioxide of claim 7, further comprising:

- storing said chemical pouch in a chemical pack;
- said chemical pack having an amount of wicking material within said chemical pack to secure said chemical pouch within said chemical pack and physically separated from at least one portion of said chemical pack wall by said wicking material;
- said chemical pack having a non-water resistant housing wall for allowing water to pass through said housing wall, such that the introduced water contacting the chemical pack wall travels through said chemical pack wall and is wicked in to contacting said chemical pouch wall to cause a reaction between said sodium chlorite and said acid.

9. The method for generating chlorine dioxide of claim 8, further comprising:

- providing a foil pouch housing selectively surrounding said chemical pack and chemical pouch for preventing moisture from contacting said chemical pack outer wall prior to introducing water with said chemical pouch.

10. The method for generating chlorine dioxide of claim 9, further comprising:

- providing a cup formed of non-water permeable material for storing said foil pouch, chemical pack and chemical pouch during shipment;
- and for placing said chemical pack containing said chemical pouch within said cup during the introduction of water to the chemical pouch to form chlorine dioxide gas within said cup.

11. A method for generating chlorine dioxide, comprising: providing a first non-water permeable housing for storing a chemical pack;

said chemical pack having outer walls formed of a non-water resistant material;

said chemical pack having a water wicking material stuffing within said chemical pack outer walls and a chemical pouch within said wicking material stuffing;

said chemical pouch having chemical pouch outer walls for containing active chemical ingredients within said chemical pouch outer walls;

said active ingredients comprising at least sodium chlorite and an acid within said chemical pouch;

providing a stabilizer within said chemical pouch to separate said active chemical ingredients and to prevent activation of the active chemical ingredients prior to introduction of a preselected amount of water to the chemical pouch during an activation step;

introducing water to the chemical pouch by removing said chemical pack from said non-water permeable housing and allowing water to contact said chemical pack outer walls;

said introduced water permeating or dissolving said chemical pack outer walls and contacting said wicking material;

said wicking material wicking said introduced water into contact with said chemical pouch outer wall;

said introduced water permeating or dissolving said chemical pouch outer walls and contacting said active ingredients and said stabilizer;

said introduced water dissolving or washing away said stabilizer to allow contact of the sodium chlorite with said acid to form chlorine dioxide.

12. The method for generating chlorine dioxide of claim **11**, wherein said stabilizer is a talcum powder.

13. The method for generating chlorine dioxide of claim **11**, wherein said chemical pack is replaced within said first

non-water permeable housing and the introduced water is brought into contact with the chemical pack within said first non-water permeable housing.

14. The method for generating chlorine dioxide of claim **13**, wherein said chemical pack is stored within a second non-permeable housing within said first non-water permeable housing, and said chemical pack is removed from said first non-water permeable housing and from said second non-water permeable housing prior to replacing the chemical pack within said non-permeable housing and the introduced water is then added to the first non-water permeable housing to bring the water into contact with the chemical pack.

15. The method for generating chlorine dioxide of claim **14**, wherein said first non-permeable housing is a paper cup and the second said non-permeable housing is a foil pack, and said introduced water is in the form of liquid or vapor.

16. The method for generating chlorine dioxide of claim **11**, further comprising:

providing a measuring cup within said first non-water permeable housing for measuring a preselected amount of water to be introduced to the chemical pouch.

17. The method for generating chlorine dioxide of claim **11**, wherein at least one of said chemical pack outer walls and said chemical pouch outer walls are formed from dissolvable filter paper.

18. The method for generating chlorine dioxide of claim **11**, wherein said chemical pouch includes at least one interior wall for separating said sodium chlorite from said acid prior to introduction of said water, and said at least one interior wall dissolves when said water is introduced to allow contact between said acid and said sodium chlorite.

* * * * *