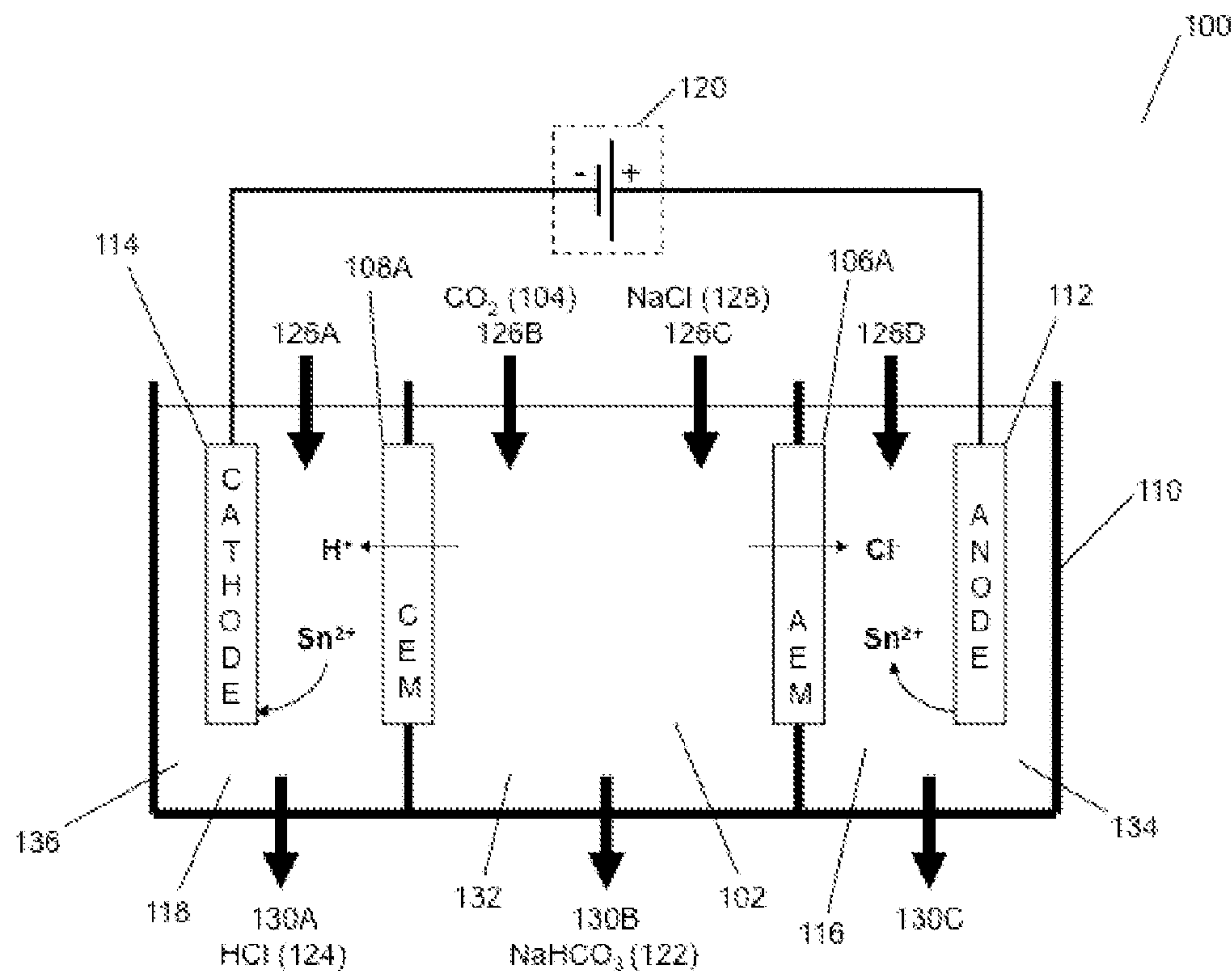


US 20110042230A1

(19) **United States**(12) **Patent Application Publication**
Gilliam et al.(10) **Pub. No.: US 2011/0042230 A1**(43) **Pub. Date: Feb. 24, 2011**(54) **LOW-ENERGY ELECTROCHEMICAL
BICARBONATE ION SOLUTION****Publication Classification**(76) Inventors: **Ryan J. Gilliam**, San Jose, CA
(US); **Bryan Boggs**, Campbell, CA
(US); **Valentin Decker**, San Jose,
CA (US)(51) **Int. Cl.**
C25B 1/14 (2006.01)
C25B 9/00 (2006.01)
C25B 1/00 (2006.01)
C25B 1/26 (2006.01)
C25B 1/18 (2006.01)Correspondence Address:
Calera Corporation
Eric Witt
14600 Winchester Blvd.
Los Gatos, CA 95032 (US)(52) **U.S. Cl. 205/482; 204/242; 204/252; 205/555;
205/480**(21) Appl. No.: **12/989,781**(22) PCT Filed: **Jan. 28, 2009**(86) PCT No.: **PCT/US09/32301**§ 371 (c)(1),
(2), (4) Date: **Oct. 26, 2010**(57) **ABSTRACT**

A low-energy electrochemical method and system of forming bicarbonate ion solutions in an electrochemical cell utilizing carbon dioxide in contact with an electrolyte contained between two ion exchange membranes in an electrochemical cell. On applying a low voltage across an anode and cathode in electrical contact with the ion exchange membranes, bicarbonate ions form in the electrolyte without forming a gas, e.g., chlorine or oxygen at the electrodes.



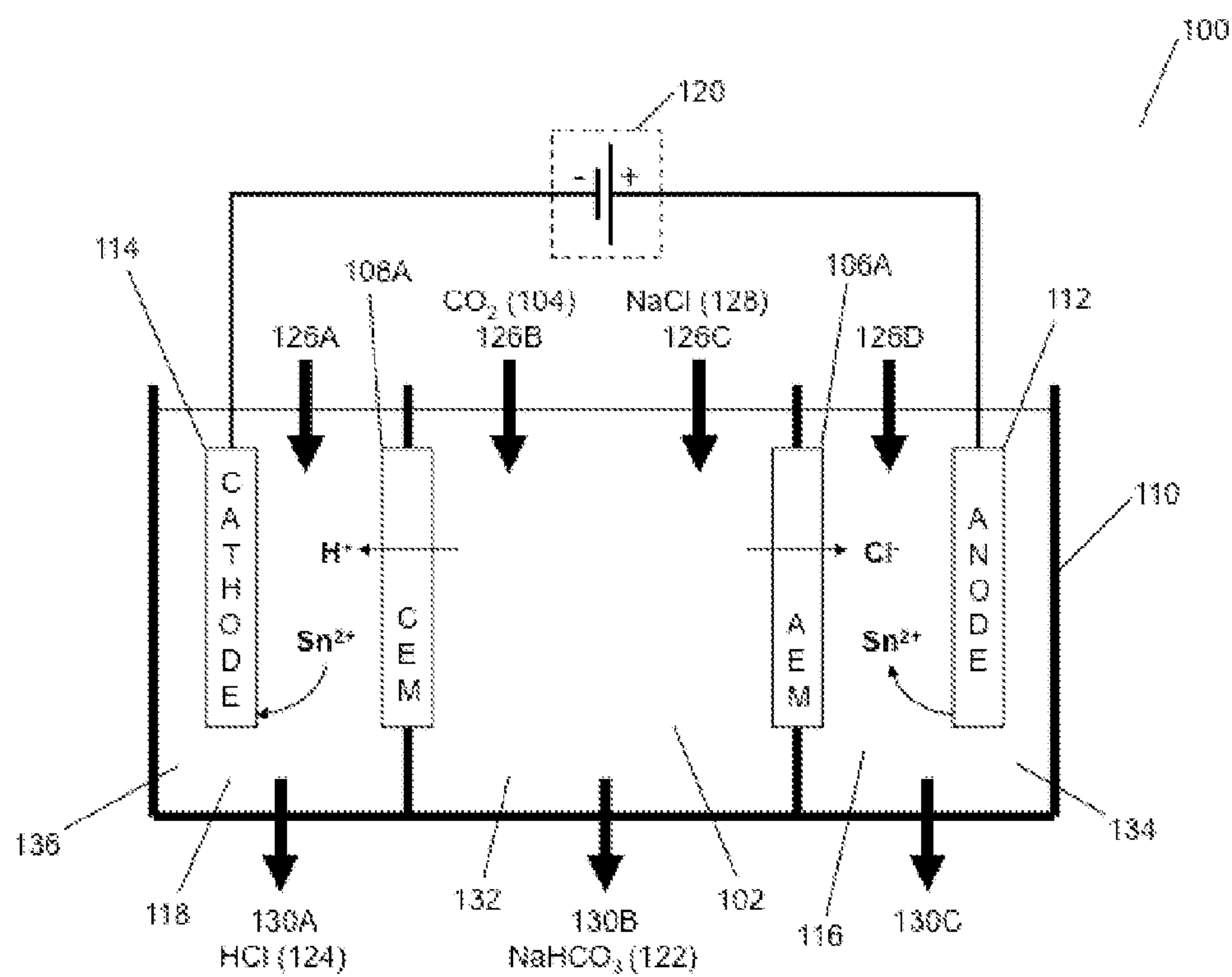


Fig. 1

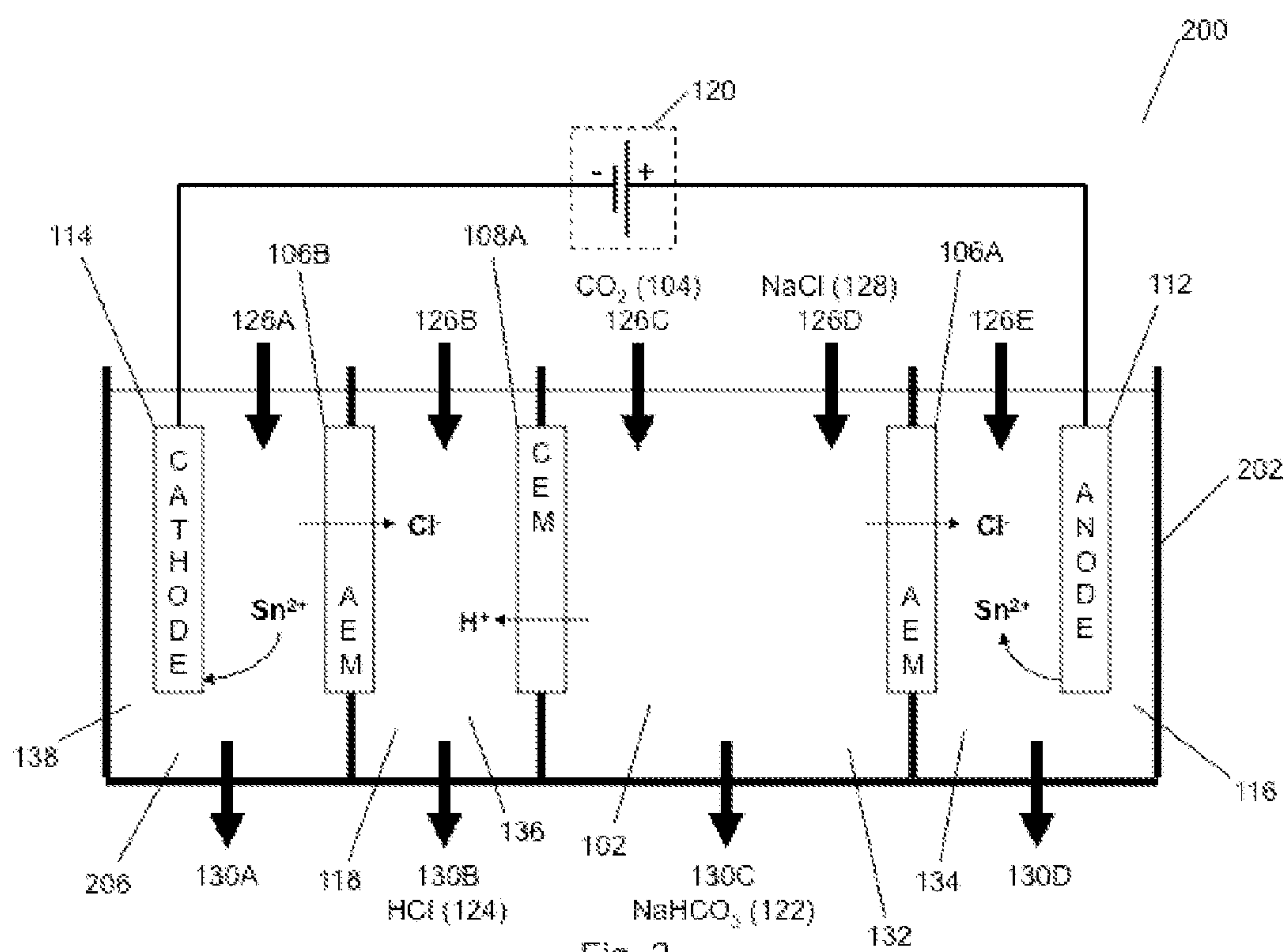
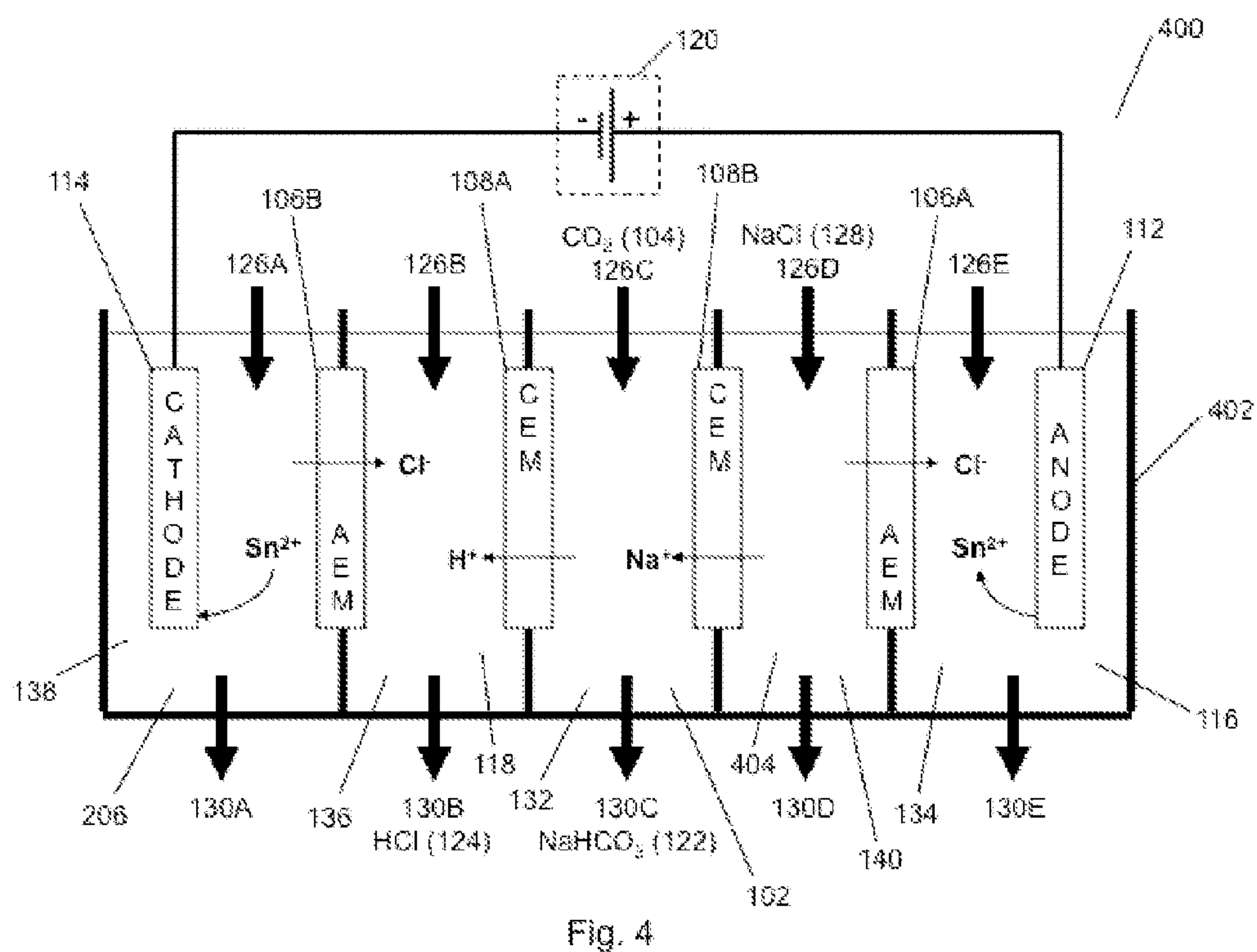
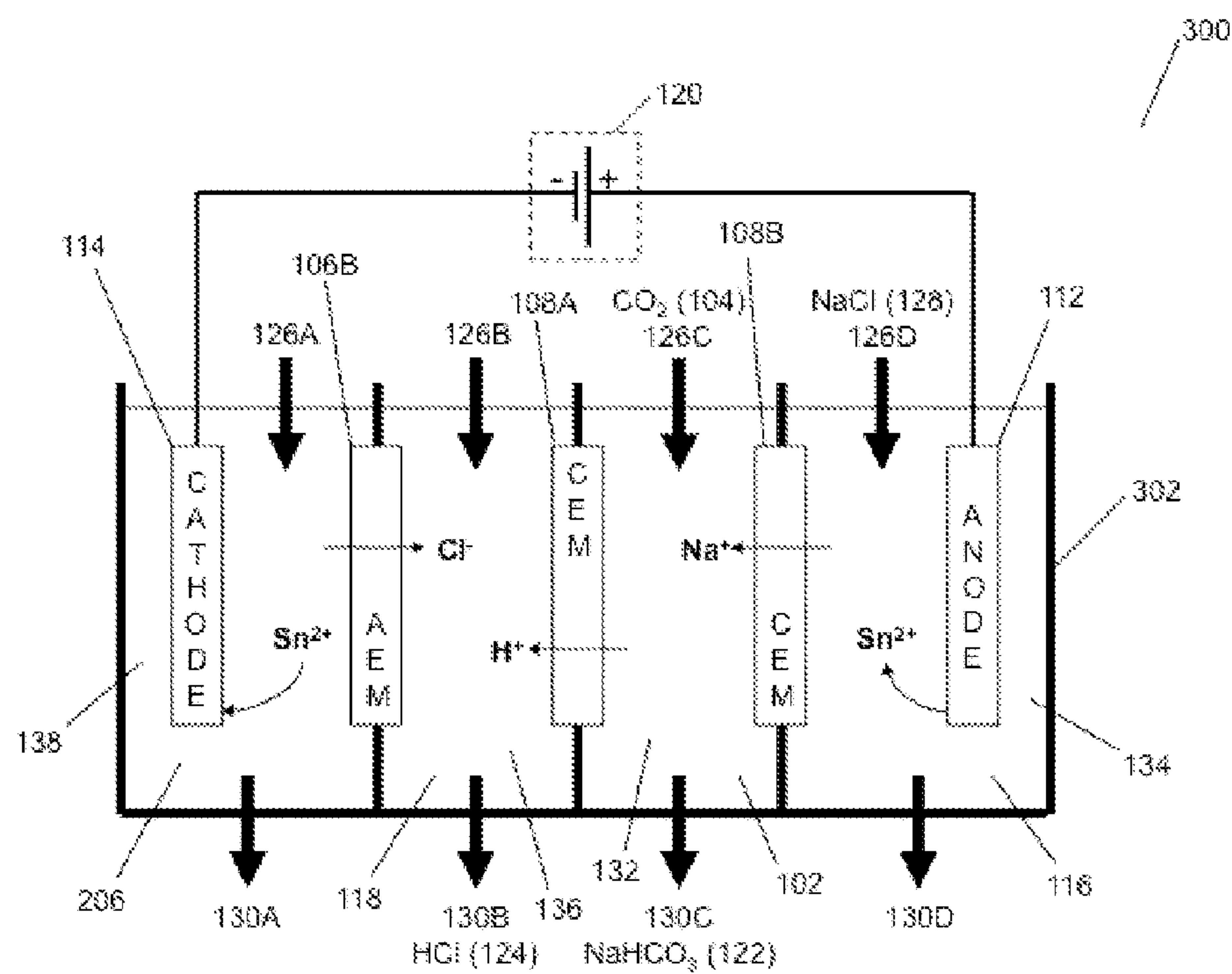


Fig. 2



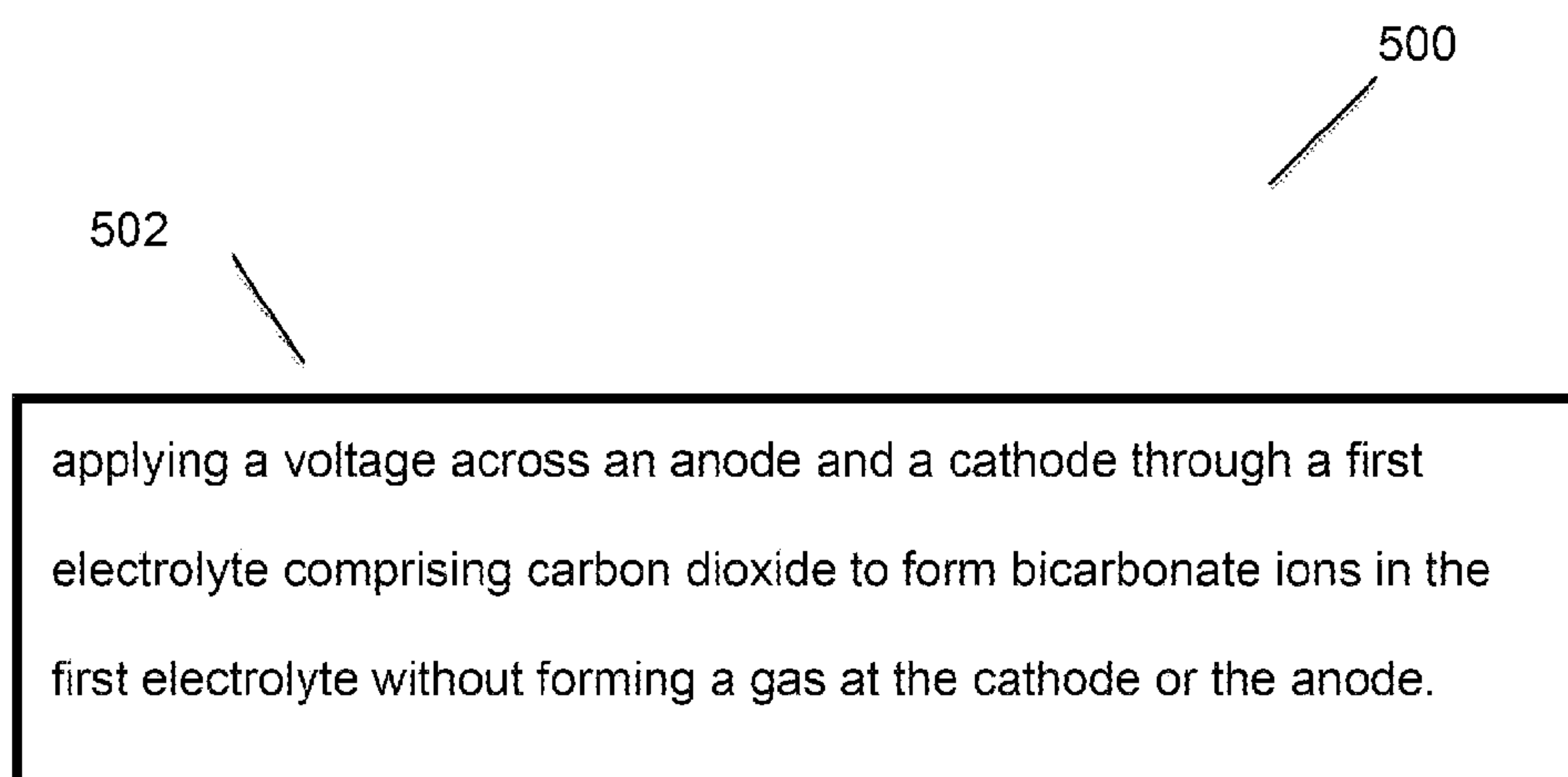


Fig. 5

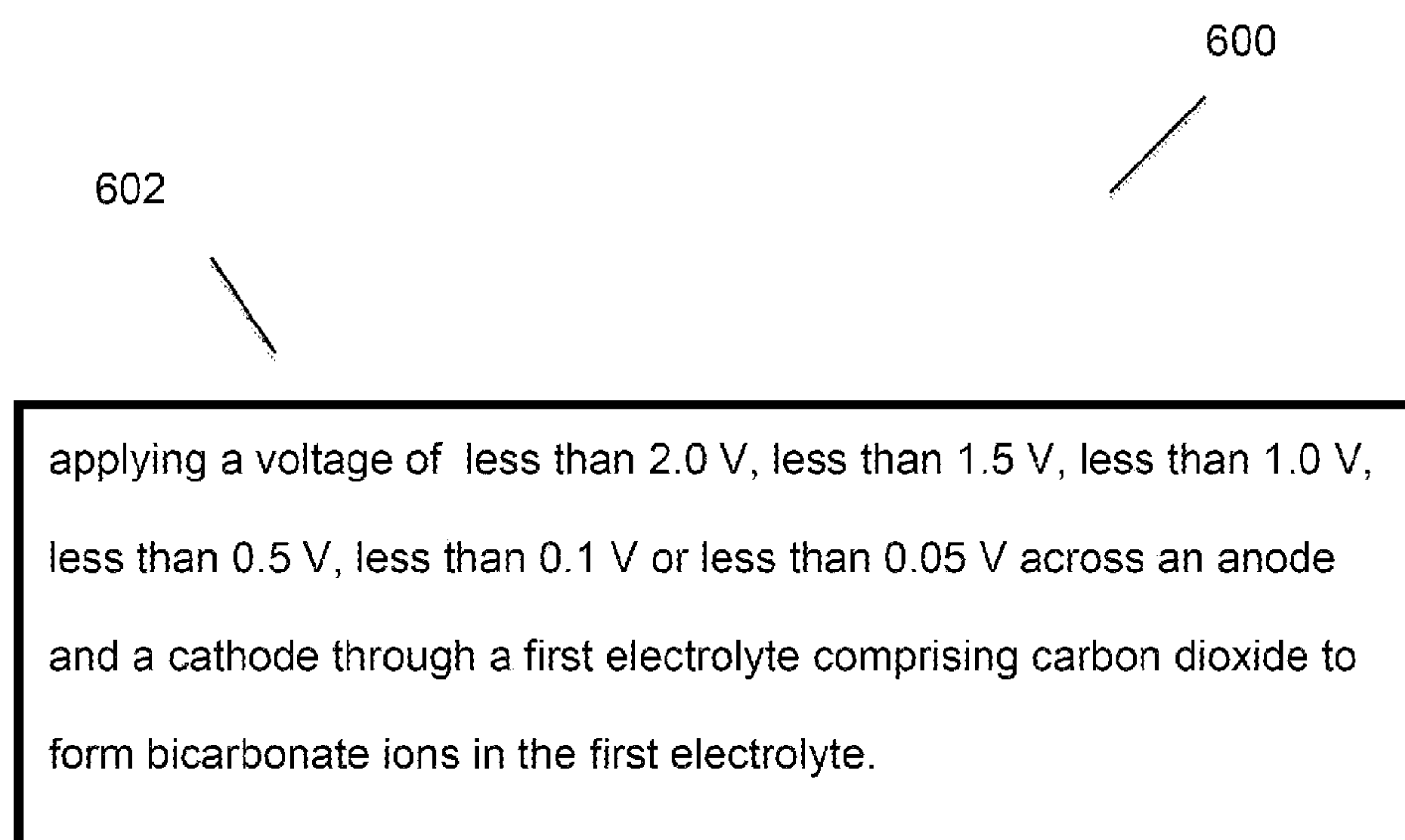


Fig. 6

LOW-ENERGY ELECTROCHEMICAL BICARBONATE ION SOLUTION

BACKGROUND

[0001] Bicarbonate ion solutions are used to regulate or achieve a chemical reaction or buffer the pH of a solution. Conventionally, bicarbonate ion solutions are obtained by dissolving bicarbonate salts, e.g., sodium bicarbonate, in water. However, producing bicarbonate salts conventionally is energy intensive and, consequently, bicarbonate ion solutions are expensive.

SUMMARY

[0002] This invention pertains to a low energy system and method of producing bicarbonate ions utilizing an electrolyte and carbon dioxide in an electrochemical cell. In one embodiment, the system comprises an anode, a cathode and an electrolyte contained between ion exchange membranes in an electrochemical cell. On applying a voltage across the anode and cathode while contacting the electrolyte with carbon dioxide, the system is capable of forming bicarbonate ions in the electrolyte without forming a gas at the electrodes, e.g., without forming hydrogen at the cathode or chlorine at the anode. The system is also capable of forming an acid, e.g., hydrochloric acid in another electrolyte in contact with an ion exchange membrane; and, in various embodiments, ions of the anode can be recovered at the cathode by reusing the anode electrolyte at the cathode.

[0003] In another embodiment, the system comprising an anode, a cathode and an electrolyte contained between ion exchange membranes, is capable of forming bicarbonate ions in the electrolyte on applying a voltage of, e.g., less than 0.05 V across the anode and cathode while contacting the electrolyte with carbon dioxide. The system is also capable of forming an acid, e.g., hydrochloric acid in another electrolyte in contact with an ion exchange membrane; and, in various embodiments, ions of the anode can be recovered at the cathode by reusing the anode electrolyte at the cathode.

[0004] In one embodiment, the method comprises applying a voltage across an anode and a cathode in an electrochemical cell containing an electrolyte comprising carbon dioxide and contained between ion exchange membranes, to form bicarbonate ions in the electrolyte without forming a gas at the electrodes, e.g., without forming chlorine at the anode or hydrogen at the cathode. The method is also capable of forming an acid, e.g., hydrochloric acid in another electrolyte in contact with an ion exchange membrane; and, in various embodiments, ions of the anode can be recovered at the cathode by reusing the anode electrolyte at the cathode.

[0005] In another embodiment, the method comprises forming bicarbonate ions in an electrolyte contained between ion exchange membranes in an electrochemical cell by applying a voltage of less than 2.0 V, less than 1.5 V, less than 1.0 V, less than 0.5 V, less than 0.1 V or less than 0.05 V across the anode and cathode while contacting the electrolyte with carbon dioxide. The system is also capable of forming an acid, e.g., hydrochloric acid in another electrolyte in contact with an ion exchange membrane; and, in various embodiments, ions of the anode can be recovered at the cathode by reusing the anode electrolyte at the cathode.

[0006] With the present system and method, carbon dioxide from any convenient source can be used to contact the electrolyte between the ion exchange membranes. Such sources

include carbon dioxide dissolved in a liquid, carbon dioxide in solid form, e.g., dry ice, or gaseous carbon dioxide. In particular embodiments, carbon dioxide in combustion gases of an industrial plant, e.g., the stack gases of fossil fuel power-generating plants or cement plants can be used.

[0007] In various embodiments, the present system and method are adaptable for batch, semi-batch or continuous flows of electrolytes, bicarbonate ions, carbon dioxide and acid in the electrochemical cell. In various embodiments, the solution comprising bicarbonates ions can be used to sequester carbon dioxide by contacting the bicarbonate ion solution with an alkaline earth metal ion solution in the presence of carbon dioxide to precipitate carbonates, e.g., to precipitate calcium and magnesium carbonates from saltwater as described in U.S. patent application Ser. No. 12/126,776, filed on May 23, 2008, herein incorporated by reference. The precipitated carbonates, in various embodiments, can be used as building products, e.g., cements and other building products as described in the United States Patent Applications incorporated herein by reference.

[0008] In another embodiment, the system and method can be used to precipitate carbonates from saltwater to produce desalinated water as described in U.S. patent application Ser. No. 12/163,205, filed on Jun. 27, 2008, herein incorporated by reference. In various embodiments, the acids produced by the present method can be used to dissolve alkaline earth metal minerals to obtain alkaline earth metal cations for use in sequestering carbon dioxide as described in the United States patent applications incorporated herein by reference.

BRIEF DESCRIPTION OF THE DRAWINGS

[0009] The following figures illustrate by way of examples and not by limitation various embodiments of the present system and method:

[0010] FIG. 1 illustrates an embodiment of the present system.

[0011] FIG. 2 illustrates an embodiment of the present system.

[0012] FIG. 3 illustrates an embodiment of the present system.

[0013] FIG. 4 illustrates an embodiment of the present system.

[0014] FIG. 5 is a flow chart of an embodiment of the present method.

[0015] FIG. 6 is a flow chart of an embodiment of the present method.

DETAILED DESCRIPTION

[0016] In the following detailed description of exemplary embodiments of the system and method where a range of values is specified, each intervening value in the range is encompassed by the invention. Thus, values between the upper and lower limit of the range and any other stated and intervening value in the range are included unless the context clearly dictates otherwise. Also, upper and lower limits of smaller ranges are included in smaller ranges and are encompassed within the scope of the invention, subject to any specifically excluded limit in the stated range.

[0017] Herein, numerical values may be preceded by the term "about." The term "about" is used to provide literal support for the exact number that it precedes, and/or as a number that is near to or approximately the number that it precedes. In determining whether a number is near to or

approximately a specifically recited number, the near and/or approximating unrecited number may be a number that, in the context in which it is presented, provides the substantial equivalent of a specifically recited number.

[0018] Herein, unless otherwise specified, all technical and scientific terms have the same meaning as understood by one of ordinary skill in the art to which this invention pertains. Publications and patents incorporated by reference herein are fully incorporated to disclose their contents as disclosed. A publication, when cited, is cited for its disclosure on its publication date and is not an admission that the present invention is not entitled to antedate such publication by virtue of prior invention. The date of a cited publication may differ from the actual publication date and may need independent confirmation.

[0019] Herein, the singular forms “a,” “an,” and “the” encompass plural forms unless the context clearly dictates otherwise. As will be apparent to one ordinarily skilled in the art, each of the embodiments described and illustrated herein comprises discrete elements that may be separated from, or combined with, other elements without departing from the scope of the claims, e.g., a recited method may be performed in the order of events recited or in another logical order without departing from the scope of the claims.

[0020] Herein, the invention in various embodiments is described for convenience in terms of producing sodium bicarbonate ions, and optionally, hydrochloric acid. However, it will be appreciated by one ordinarily skilled in the art that the present system and method may produce other bicarbonate ions such as, e.g., potassium and calcium bicarbonate ions and other acids such as sulfuric acid, depending on the electrolytes used.

[0021] In various embodiments, the present invention is directed to a low voltage system and method of forming bicarbonate ions by contacting carbon dioxide with an electrolyte salt solution positioned between ion exchange membranes in an electrochemical cell. In one embodiment, on applying a low voltage across a cathode and anode in the cell, bicarbonate ions form in the solution without forming a gas at the electrodes, e.g., without forming chlorine at the anode or hydrogen at the cathode. By the present system and method, bicarbonate ions are formed in the solution on applying a voltage across the anode and cathode of less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6, 1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V, and other low voltages as disclosed herein. In various embodiments, an acid solution is also formed in another electrolyte in contact with an ion exchange membrane, e.g., hydrochloric acid, in the electrochemical cell. Optionally, in some embodiments, the electrolyte in contact with the anode is reused as the electrolyte at the cathode to recover anode material at the cathode.

[0022] Referring to FIG. 1, in one embodiment system 100 comprises first electrolyte 102 and carbon dioxide 104 contained between anion exchange membrane 106A and cation exchange membrane 108A in an electrochemical cell 110. Electrochemical cell 100 includes anode 112 and cathode 114; second electrolyte 116 contacting anion exchange membrane 106A and anode 112; and third electrolyte 118 contacting cation exchange 108A membrane and cathode 114. On applying a voltage across the anode and cathode, the system is capable of forming bicarbonate ions 122 in first electrolyte 102 without forming a gas, e.g., hydrogen at cathode 114 or chlorine at anode 112. In various embodiments the system is

capable of forming bicarbonate ions in first electrolyte 102 when a voltage of 0.2 V or less, 0.3 V or less, 0.4 V or less, 0.5 V or less, 0.6 V or less, 0.7 V or less, or 0.8 V or less is applied across the anode and cathode.

[0023] In various embodiments and with reference to FIGS. 1-4, the system also capable of forming an acid 124 in third electrolyte solution 118 contacting cation exchange membrane 108A as a result of transfer of protons across cation exchange membrane 108A from first electrolyte 102. For example, as is illustrated in FIG. 1, protons transferred from first electrolyte 102 to third electrolyte 118 will result in formation of an acid solution 124 in third electrolyte 118; thus, where third electrolyte contains chloride ions, hydrochloric acid 124 will form in third electrolyte 118.

[0024] Depending on whether a sacrificial anode is used, e.g., tin, copper, iron, zinc, the system in various embodiments is capable of oxidizing the anode to form cations in the electrolyte in contact with the anode e.g., tin ions, in second electrolyte 116. Hence, as is illustrated in FIG. 1, in various embodiments the system is capable of forming a chloride solution in second electrolyte 116 contacting the anode, e.g., where a tin anode is used and tin ions are present in the second electrolyte 116, stannous chloride will form in second electrolyte 116 as a result of transfer of chloride ions across anion exchange membrane 106A from first electrolyte 102. Similarly, as illustrated in FIGS. 2-4, where a tin anode is used, stannous chloride solution will form in electrolyte 116 as a result of ions migrating to or from second electrolyte 116 across the ion exchange membrane in contact with second electrolyte 116 as discussed below. In various embodiments, optionally, the electrolyte solution 116 in contact with anode 112 comprising anode ions can be reused as electrolyte 118 in contact with cathode 114 to recover anode material at the cathode. As will be appreciated by one ordinarily skilled in the art, tin and other sacrificial metal can thus be recovered at the cathode, depending on the material used as the sacrificial anode.

[0025] With reference to FIGS. 1-4, in various embodiments, system 100, system 200, system 300 and system 400 comprise inlet ports 126 A-E (where needed) for introducing substances in to the cell, e.g., for introducing fluids, gases, salts and the like into cells 110, 202, 302, 402; and outlet ports 130A-E (where needed) for removing fluids from the cells. For example, with reference to FIG. 1, system 100 comprises inlet port 126B for introducing carbon dioxide 104 into first electrolyte 102, and inlet port 126C for introducing sodium chloride solution 128 into first electrolyte 102. Similarly, system 100 of FIG. 1 comprises outlet ports 130A for removing acid 124 from third compartment 136, and outlet port 130B for removing bicarbonate ion solution from first compartment 132. As will be appreciated by one ordinarily skilled in the art, the inlet and outlet ports are adaptable for various flow protocols including batch flow, semi-batch flow, or continuous flow. In various embodiments, the system includes voltage regulator 120 for regulating voltages across the electrodes and currents through the electrolytes.

[0026] In an embodiment illustrated in FIG. 1, electrochemical cell 110 comprises first compartment 132, second compartment 134 and third compartment 136 formed by positioning anion exchange membrane 106A and cation exchange membrane 108A in cell 110 such that first electrolyte 102 is separated from second electrolyte 116 and third electrolyte 118. As will be appreciated in the art, the ion exchange membranes are positioned to contact the electrolytes on opposite

surfaces such that ions from one electrolyte will migrate to another electrolyte through the ion exchange membrane without mixing of the electrolytes.

[0027] In various embodiments as illustrated in FIGS. 1-4, the system, depending on its configuration, is initially charged (where appropriate) with first electrolyte **102**, second electrolyte **116**, third electrolyte **118**, fourth electrolyte **206** and fifth electrolyte **404** comprising an aqueous salt solution such as a saltwater, e.g., seawater, brine, brackish water, sodium chloride, conductive fresh water and the like. In an embodiment that produced the results as set forth in Table 1, the system was initially charged with first electrolyte **102** and fifth electrolyte **404** comprising 2 M sodium chloride solution; in another embodiment the system was initially charged with first electrolyte **102** and fifth electrolyte **404** comprising 0.5 M sodium chloride solution. In other specific embodiments the system can be charged initially with a salt solution, e.g., sodium chloride, at a concentration from 0.1 to 4 M, e.g., 0.1 to 2.5 M, or 0.2 to 2.0 M, or 0.1 to 1.0 M, or 0.2 to 1.0 M, or 0.2 to 0.8 M, or 0.3 to 0.7 M, or 0.4 to 0.6 M, or 0.5 to 2.5 M, or 1.0 to 2.5 M, or 1.5 to 2.5 M, or 1.7 to 2.3 M.

[0028] With reference to FIGS. 1-4, anion exchange membranes **106A**, **106B** and cation exchange membranes **108A**, **108B** comprise ionic membranes selectively permeable to one ion or one class of ions, e.g., cation membranes selectively permeable to sodium ions only or hydrogen ions only, or to cations generally; or anion membranes selectively permeable to chloride ions only or to anions generally, can be used. In various embodiments, anion exchange membranes **106A**, **106B** and cation exchange membranes **108A**, **108B** may comprise membranes that will function in an acid and/or basic electrolytic at pH from 1 to 14; also, the membranes may be selected to function with electrolytes wherein the temperatures ranges from about 0° C. to 100° C. or higher. Such ion exchange membranes are commercially available, e.g., PCA GmbH of Germany supplies a suitable anion exchange membrane permeable to chloride ions and identified as PCSA-250-250; and a cation exchange membrane permeable to sodium ions and identified as PCSK 250-250.

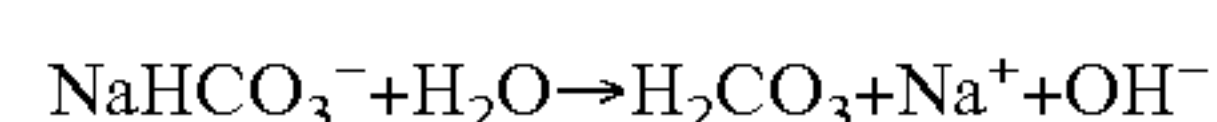
[0029] With reference to FIGS. 1-4, in various embodiments anode **112** comprises a sacrificial anode, e.g., tin, copper, iron, zinc. Where a sacrificial anode such as tin is used, cations such as Sn^{2+} will form in second electrolyte **116** in contact with anode **112**. Optionally, as will be appreciated by one ordinarily skilled in the art, cations in electrolyte **116** in contact with anode **112** can be recovered by plating out the cations at the cathode **114**, e.g., using electrolyte **116** from the anode as the electrolyte at the cathode. Thus, the anode material can be recovered at the cathode by switching electrolyte **116** in contact anode **112** with the electrolyte in contact with the cathode **114** when a sufficient concentration of Sn^{2+} has accumulated in the electrolyte **116**, and allowing the cations to plate out at the cathode. It will also be appreciated that when sacrificial anode **112** is diminished and cathode **114** is augmented sufficiently, these electrodes may be switched so that anode **112** is transferred to replace cathode **114** and vice versa.

[0030] As is illustrated in FIGS. 1-4, the voltage across anode **112** and cathode **114** can be regulated to form bicarbonate ions **122** in first electrolyte **102** without forming a gas, e.g., chlorine at anode **112** or hydrogen at cathode **114**. In various embodiments, bicarbonate ions **122** are formed when the voltage applied across anode **112** and cathode **114** is less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6,

1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0031] In various embodiments as illustrated in FIGS. 1-4, a protonated solution, e.g., hydrochloric acid **124** is formed in third electrolyte **118**. For example, with reference to FIG. 1, on transfer of protons from first electrolyte **102** to third electrolyte **118** through cation exchange membrane **108A**, the pH of the third electrolyte **118** will adjust, e.g., become more acid if protons accumulate in the electrolyte. The acid formed will depend on the electrolytes used, e.g., as illustrated in FIG. 1, where third electrolyte **118** comprises chloride ions, hydrochloric acid will form in third electrolyte **118**. With the accumulation of protons in third electrolyte **118**, the pH of this electrolyte will decrease; it will be appreciated, however, that the pH of third electrolyte may increase, decrease or remain constant depending on the rate of removal of third electrolyte from the system.

[0032] Also as will be appreciated by one skilled in the art and as is illustrated, e.g., in FIG. 1, where first electrolyte **102** initially comprises sodium chloride solution **128**, sodium bicarbonate **122** will form in first electrolyte **102** as a consequence of the migration of protons and chloride ions from first electrolyte **102**. Further, as sodium bicarbonate is an amphoteric salt that forms a mildly alkaline solution in water, with the formation of sodium bicarbonate in first electrolyte **102** the pH of the first electrolyte will increase (assuming that first electrolyte **102** is not removed from the system) due to formation of hydroxyl ions (OH^-) in accordance with the following reaction:

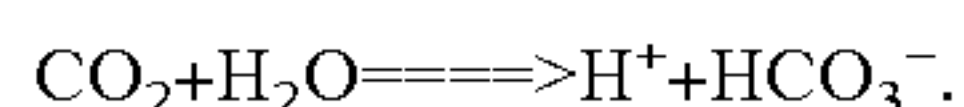


[0033] In various embodiments of the invention as illustrated in FIGS. 1-6, carbon dioxide **104** from any convenient source can be used. Such sources include carbon dioxide dissolved in a liquid, solid carbon dioxide, e.g., dry ice, or gaseous carbon dioxide. In various embodiments, carbon dioxide in post-combustion effluent stacks of industrial plants such as power plants, cement plants and coal processing plants can be used. In various embodiments carbon dioxide **104** may comprise substantially pure carbon dioxide or a

multi-component gaseous stream comprising carbon dioxide and one or more additional gases. Additional gases and other components may include CO, SO_x (e.g., SO₂), NO_x, mercury and other heavy metals and dust particles e.g., from calcining and combustion processes. In various embodiments, one or more of these additional components can be precipitated by contacting first electrolyte **102** with a solution of alkaline earth metal ions, e.g., where SO₂ is contained in the gas stream, sulfates and sulfides of calcium and magnesium can be precipitated.

[0034] Multi-component gaseous streams include reducing condition streams, e.g., syngas, shifted syngas, natural gas, and hydrogen and the like, and oxidizing condition streams, e.g., flue gases from combustion. Such gaseous streams include oxygen-containing flue gas, e.g., from a coal fired power plant, a cement plant, or a natural gas power plant; turbo charged boiler product gas; coal gasification product gas; shifted coal gasification product gas; anaerobic digester product gas; wellhead natural gas; reformed natural gas or methane hydrates; and the like. In various embodiments, gases that are not absorbed in first electrolyte **102**, e.g., nitrogen, in one embodiment are vented from the system; in other embodiments, the gases are collected for other uses.

[0035] As will be appreciated by one skilled in the art and with reference to FIGS. 1-6, without being bound by any theory it is believed that bicarbonate ions (HCO₃⁻) form in first electrolyte **102** as a result of carbon dioxide contacting water in the first electrolyte **102**, as follows:



Thus, in accordance with the present invention and with reference to FIG. 1, where first electrolyte **102** comprise Na⁺ and Cl⁻ ions from added sodium chloride **128**, by placing first electrolyte **102** between cation exchange membrane **108A** selective to transferring H⁺ ions, and an anion exchange membrane **106A** selective to transferring of Cl⁻ ions, and applying a voltage across the electrodes, H⁺ will migrate through the cation exchange membrane **108A** to adjacent third electrolyte **118**. Similarly, Cl⁻ will migrate from first electrolyte **102** through the anion exchange membrane **106A** to adjacent second electrolyte **116**. Consequently, in first electrolyte **102**, a solution comprising sodium bicarbonate will form. Depending on the rate of introduction and/or removal of first electrolyte from the system and the voltage applied across electrodes **112**, **114**, the concentration of bicarbonate ions in first electrolyte **102** will adjust, e.g., increase, decrease or will not change.

[0036] Also, with reference to FIG. 1, as H⁺ migrate from first electrolyte **102** through cation exchange membrane **108A** to adjacent electrolyte **118**, the pH of adjacent third electrolyte **118** will adjust depending on rate of introduction and/or removal of first electrolyte **102** from the system. Similarly, as chloride ions migrate from the first electrolyte to adjacent second electrolyte **114** across the anion exchange membrane **106A**, the chloride in second electrolyte **114** will adjust, e.g., increase, decrease or does not change. Hence, as illustrated in FIGS. 1-6, in various embodiments of the system and method, a solution of bicarbonate ions **122**, e.g., sodium bicarbonate, is obtained in first electrolyte **102**, an acid solution **124**, e.g., hydrochloric acid, is obtained in third electrolyte **118**, and a chloride solution is obtained in second electrolyte **116**.

[0037] In an embodiment of system **200** as illustrated in FIG. 2, first electrolyte **102** and carbon dioxide **104** are con-

tained between first anion exchange membrane **106A** and first cation exchange membrane **108A** in an electrochemical cell **202** comprising anode **112** and cathode **114**. In the system, second electrolyte **116** contacts first anion exchange membrane **106A** and anode **112**; third electrolyte **118** is contained between first cation exchange membrane **108A** and second anion exchange membrane **106B**; and fourth electrolyte **206** contacts second anion exchange membrane **106B** and cathode **114**, wherein on applying a voltage **130** across cathode **114** and anode **112**, the system forms bicarbonate ions **122** in first electrolyte **102** without forming a gas at the cathode or anode. In various embodiments, the system forms bicarbonate ions in first electrolyte **102** when a voltage of 0.4 V or less, or 0.6 V or less, or 0.8 V or less is applied across the anode and cathode. In various embodiments, bicarbonate ions **122** are formed when the voltage applied across anode **112** and cathode **114** is less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6, 1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0038] System **200** in various embodiments will form an acid, e.g., hydrochloric acid **124**, depending on the electrolytes used. As is illustrated in FIG. 2, the system will form a protonated solution (acid solution) in third electrolyte **118**, e.g., hydrochloric acid as a result of transfer of protons across cation exchange membrane **108A** from first electrolyte **102**; and an ionic solution, e.g., stannous chloride will form in second electrolyte **116** as a result of chloride ions transferring across anion exchange membrane **106** from first electrolyte **102**, assuming tin ions are present in the second electrolyte **116** from oxidation of anode **112** comprising tin. In various embodiments, optionally, electrolyte **116** in contact with anode **112** is reused as electrolyte **118** in contact with cathode **114** to recover anodic metal that may have oxidized into second electrolyte **116** at anode **112**. Likewise, electrolyte **206** in contact with cathode **114** may be reused as electrolyte **116** in contact with anode **112**. It will be appreciated that when sacrificial anode **112** is diminished and cathode **114** is

augmented sufficiently, these electrodes may be switched so that anode **112** is transferred to replace cathode **114** and vice versa.

[0039] As is illustrated in FIG. 2, system **200** includes inlet ports **126 A-E** adapted for introducing materials into cell **202**, e.g., for introducing carbon dioxide **104**, sodium chloride solution **126** and other electrolytes into cell **202**; and outlet ports **130 A-D** for removing materials from the cell, e.g., removing bicarbonate solution **122** and acid **124** from the cell. As will be appreciated by one ordinarily skilled in the art, the inlet and outlet ports are adaptable for various flow protocols including batch flow, semi-batch flow, or continuous flow. In various embodiments, the system includes voltage/current regulator **120** for regulating currents and voltages across the anode, cathode and the electrolytes.

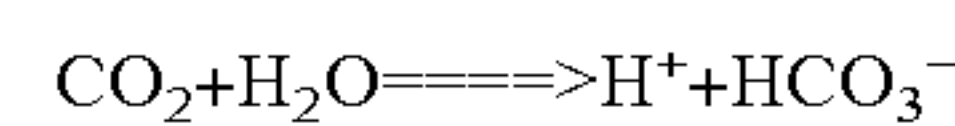
[0040] In the system illustrated in FIG. 2, electrochemical cell **202** comprises first compartment **132**, second compartment **134**, third compartment **136** and fourth compartment **138** formed by positioning first anion exchange membrane **106A** and first cation exchange membrane **108A** to separate first electrolyte **102** from second electrolyte **116** and third electrolyte **118**, and by positioning second anion exchange membrane **106B** to separate third electrolyte **118** from fourth electrolyte **206**. As will be appreciated in the art, the ion exchange membranes in various embodiments are positioned to contact the electrolytes at opposite surfaces to allow movement of ions from one electrolyte to another electrolyte through the ion exchange membranes without mixing of the electrolytes.

[0041] As with the system of FIG. 1, in the system of FIG. 2, first **102**, second **116** third **118** and fourth **206** electrolytes initially may comprise an aqueous salt solution such as a saltwater, e.g., seawater, brine, brackish water, conductive fresh water and the like. In results obtained from one embodiment as set forth in Table 1, first electrolyte **102** initially comprised 2 M sodium chloride solution; in another embodiment first electrolyte **102** comprised 0.5 M sodium chloride solution. In specific embodiments the system may be charged initially with a salt solution, e.g., sodium chloride, at a concentration from 0.1 to 4 M, e.g., 0.1 to 2.5 M, or 0.2 to 2.0 M, or 0.1 to 1.0 M, or 0.2 to 1.0 M, or 0.2 to 0.8 M, or 0.3 to 0.7 M, or 0.4 to 0.6 M, or 0.5 to 2.5 M, or 1.0 to 2.5 M, or 1.5 to 2.5 M, or 1.7 to 2.3 M.

[0042] With reference to FIG. 2, the voltage across anode **112** and cathode **114** can be regulated to form bicarbonate ions **122** in first electrolyte **102** without forming a gas, e.g., chlorine at anode **112** or hydrogen at cathode **114**. Similarly, by regulating a voltage across cathode **114** and anode **112** as with the system of FIG. 1, a protonated (acid) solution **124** is formed in third electrolyte **118** in contact with cation exchange membrane **108A** by protons transferred from first electrolyte **102**. The acid solution formed will depend on the electrolytes used, e.g., as illustrated in FIG. 2, where the first electrolyte **102** comprises sodium chloride, the acid solution formed will comprise hydrochloric acid. An acid solution is formed, in various embodiments, when the voltage applied across anode **112** and cathode **114** is less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6, 1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at

the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, an acid solution is formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0043] In the system of FIG. 2 as with the system of FIG. 1, without being bound by any theory, it is believed that bicarbonate ions are formed in first electrolyte **102** by carbon dioxide contacting water in the first electrolyte, as follows:



Thus, in accordance with the present invention and with reference to FIG. 2, where first electrolyte **102** comprise Na^+ and Cl^- ions from added sodium chloride **128**, by placing first electrolyte **102** between cation exchange membrane **108A** selective to transferring H^+ ions, and anion exchange membrane **106A** selective to transferring of Cl^- ions, and applying a voltage across the electrodes, H^+ will migrate through cation exchange membrane **108A** to adjacent third electrolyte **118**. Similarly, Cl^- will migrate from the first electrolyte through the anion exchange membrane **106A** to adjacent second electrolyte **116**. Consequently, in first electrolyte **102**, a solution comprising sodium bicarbonate **122** will form. Depending on the rate of introduction and/or removal of first electrolyte from the cell and the voltage applied across electrodes **112**, **114**, the concentration of sodium bicarbonate **122** in first electrolyte **102** will be adjusted, e.g., increase, decrease or does not change.

[0044] Also with reference to FIG. 2, as H^+ migrate from first electrolyte **102** through cation exchange membrane **108A** to adjacent third electrolyte **118**, the acidity of adjacent third electrolyte **118** will adjust depending on rate of introduction and/or removal of third electrolyte **118** from the system. Similarly, as chloride ions migrate from fourth electrolyte **206** to adjacent third electrolyte **118** across second anion exchange membrane **106B** the chloride ions concentration in adjacent third electrolytes **118** and fourth electrolyte **206** will adjust.

[0045] Hence, as illustrated in FIG. 2, in various embodiments a solution of bicarbonate ions **122**, e.g., sodium bicarbonate is obtained in first electrolyte **102**; an acid solution **124**, e.g., hydrochloric acid is obtained in third electrolyte **118**; a chloride solution, e.g., tin chloride is obtained in second electrolyte **116** where a tin anode is used; and the fourth electrolyte **206** is depleted of chloride ions and cations, e.g., the electrolyte is depleted of Sn^{2+} where the fourth electrolyte was initially charged with a tin salt, e.g., stannous chloride.

[0046] Referring to FIG. 3, system **300** comprises first electrolyte **102** contained between first cation exchange mem-

brane **108A** and second cation exchange membrane **1088** and to which carbon dioxide **104** is added in an electrochemical cell **302** comprising anode **112** and cathode **114**; second electrolyte **116** contacting second cation exchange membrane **1088** and anode **112**; third electrolyte **118** contained between first cation exchange membrane **108A** and anion exchange membrane **106B** in electrochemical cell **302**; and fourth electrolyte **206** contacting anion exchange membrane **106B** and cathode **114**, wherein on applying a voltage **120** across the cathode and anode the system is capable of forming bicarbonate ions **122** in first electrolyte **102** without forming a gas at cathode **114** or anode **112**.

[0047] In system **300** illustrated in FIG. 3, electrochemical cell **302** comprises first compartment **132**, second compartment **134**, third compartment **136** and fourth compartment **138** formed by positioning first cation exchange membrane **108A** and second cation exchange membrane **1088** to separate first electrolyte **102** from second electrolyte **116** and from third electrolyte **118**; and by positioning second anion exchange membrane **1068** to separate third electrolyte **118** from fourth electrolyte **206**. As will be appreciated, the ion exchange membranes in various embodiments are positioned to contact the electrolytes at opposite surfaces to allow for movement of ions from one electrolyte to another electrolyte through the ion exchange membranes without mixing of the electrolytes.

[0048] As with the system of FIGS. 1 and 2, in the system of FIG. 3, first **102**, second **116** third **118** and fourth **206** electrolytes may initially comprise an aqueous salt solution, e.g., seawater, brine, brackish water, conductive fresh water and the like. With results achieved in one embodiment as set forth in Table 1, first electrolyte **102** initially comprised 2 M solution of sodium chloride; in another embodiment the first electrolyte **102** initially comprised 0.5 M solution of sodium chloride. In specific embodiments, electrolytes in the system may be charged initially with a salt solution, e.g., sodium chloride, at a concentration from 0.1 to 4 M, e.g., 0.1 to 2.5 M, or 0.2 to 2.0 M, or 0.1 to 1.0 M, or 0.2 to 1.0 M, or 0.2 to 0.8 M, or 0.3 to 0.7 M, or 0.4 to 0.6 M, or 0.5 to 2.5 M, or 1.0 to 2.5 M, or 1.5 to 2.5 M, or 1.7 to 2.3 M.

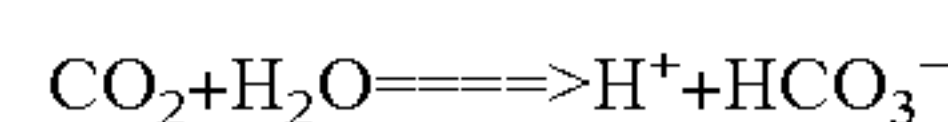
[0049] With reference to FIG. 3, the voltage across the anode **112** and cathode **114** can be regulated to form bicarbonate ions **122** in the first electrolyte **102** without forming a gas, e.g., chlorine at anode **112** or hydrogen at cathode **114**. In various embodiments, bicarbonate ions **122** are formed in first electrolyte **102** when the voltage applied across anode **112** and cathode **114** is less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6, 1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V. Similarly, on applying a voltage across the cathode and anode acid solution **124** is formed in third electrolyte **118** in contact with cation exchange membrane **108A** as a result of protons transferring from first electrolyte **102**. The acid formed depends on the electrolytes used, e.g., as illustrated in FIG. 3, where the first electrolyte **102** comprises sodium chloride, the acid formed comprises hydrochloric acid **124**.

[0050] In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied

across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0051] As is illustrated in FIG. 3, system **300** includes inlet ports **126 A-D** adapted for introducing materials into cell **302**, e.g., for introducing carbon dioxide **104**, sodium chloride solution **126** and other electrolytes into cell **302**; and outlet ports **130 A-D** for removing materials from the cell, e.g., removing bicarbonate solution **122** and acid **124** from the cell. As will be appreciated by one ordinarily skilled in the art, the inlet and outlet ports are adaptable for various flow protocols including batch flow, semi-batch flow, or continuous flow. In various embodiments, the system includes voltage/current regulator **120** for regulating voltages across the anode and cathode and currents through the electrolytes.

[0052] In the system of FIG. 3, as with the systems of FIGS. 1 and 2, without being bound by any theory, it is believed that bicarbonate ions (HCO_3^-) are formed in first electrolyte **102** by carbon dioxide contacting water in the first electrolyte, as follows:



Thus, in accordance with the present invention and with reference to FIG. 3, where second electrolyte **116** comprise Na^+ and Cl^- ions from added sodium chloride **128**, by placing first electrolyte **102** between cation exchange membrane **108A** selective to transferring H^+ ions, and second cation exchange membrane **108B** selective to transferring of cations, e.g., Na^+ ions, and on applying a voltage across the electrodes, H^+ will migrate through first cation exchange membrane **108A** to adjacent third electrolyte **118**. Similarly, Na^+ will migrate from second electrolyte **116** through second cation exchange membrane **1088** to adjacent first electrolyte **102**. Consequently, in first electrolyte **102**, a solution comprising sodium bicarbonate **122** will form. Depending on the rate of introduction and/or removal of first electrolyte **102** from the cell and the voltage applied across the electrodes **112**, **114**, the concentration of sodium bicarbonate in first electrolyte **102** will adjust, e.g., increase, decrease or will not change.

[0053] Also with reference to FIG. 3, as H^+ migrate from first electrolyte **102** through first cation exchange membrane **108A** to adjacent third electrolyte **118**, the acidity of adjacent third electrolyte **118** will adjust depending on the rate of introduction and/or removal of first electrolyte **102** from the system. Similarly, as chloride ions migrate from fourth electrolyte **206** to adjacent third electrolyte **118** across second

anion exchange membrane **106B** the chloride ion concentration in adjacent electrolytes **118** and **206** will adjust.

[0054] Hence, as illustrated in FIG. 3, in various embodiments a solution of bicarbonate ions **122**, e.g., sodium bicarbonate is obtained in first electrolyte **102**; an acid solution **124**, e.g., hydrochloric acid is obtained in third electrolyte **118**; a chloride solution, e.g., tin chloride is obtained in second electrolyte **116** where a tin anode is used; fourth electrolyte **206** is depleted of chloride ions due to chloride transfer across anion exchange membrane **106B**; and fourth electrolyte **204** is also depleted of cations by a reduction reaction at the cathode, e.g., fourth electrolyte **206** is depleted of Sn^{2+} where the fourth electrolyte was initially charged with, e.g., stannous chloride.

[0055] Optionally, as will be appreciated by one ordinarily skilled in the art, cations in electrolyte **116** in contact with anode **112** can be recovered by plating out the cations at the cathode **114**, e.g., using electrolyte **116** from the anode as the electrolyte at the cathode. Thus, the anode material can be recovered at the cathode by switching electrolyte **116** in contact anode **112** with the electrolyte in contact with the cathode **114** when a sufficient concentration of Sn^{2+} has accumulated in the electrolyte **116**, and allowing the cations to plate out at the cathode. Similarly, it will be appreciated that when sacrificial anode **112** is diminished and cathode **114** is augmented sufficiently, these electrodes may be switched so that anode **112** is transferred to replace cathode **114** and vice versa.

[0056] In an embodiment as illustrated in FIG. 4, system **400** comprises first electrolyte **102** contained between first cation exchange membrane **108A** and second cation exchange membrane **108B**; second electrolyte **116** contacting anode **112** and separated from fifth electrolyte **404** by first anion exchange membrane **106A**; third electrolyte **118** contained between first cation exchange membrane **108A** and second anion exchange membrane **106B**; fourth electrolyte **206** contacting second anion exchange membrane **106B** and cathode **114**; and fifth electrolyte **404** comprising an electrolyte containing, e.g., sodium chloride solution **128**, and contained between first anion exchange membrane **106A** and second cation exchange membrane **108B**, wherein on applying a voltage across the cathode **114** and anode **112** and adding carbon dioxide **104** to first electrolyte **102**, the system is capable of forming bicarbonate ions **122** in first electrolyte **102** without forming a gas at cathode **114** and anode **112**.

[0057] Referring to system **400** of FIG. 4, electrochemical cell **402** comprises first compartment **132**, second compartment **134**, third compartment **136**, fourth compartment **138**, and fifth compartment **140** formed by positioning first cation exchange membrane **108A** and second cation exchange membrane **108B** to separate first electrolyte **102** from fifth electrolyte **404** and from third electrolyte **118**. In the system, second anion exchange membrane **106B** is positioned to separate third electrolyte **118** from fourth electrolyte **206**; and first anion exchange membrane **106A** is positioned to separate second electrolyte **116** in contact with anode **112** from fifth electrolyte **404** comprising sodium chloride solution **128**.

[0058] As is illustrated in FIG. 4, in various embodiments initially sodium chloride **128** is added to fifth compartment **140** and carbon dioxide is added to first electrolyte **102** in compartment **132**. As will be appreciated in the art, the ion exchange membranes in various embodiments are positioned to contact the electrolytes at opposite surfaces to allow for transfer of ions from one electrolyte to another electrolyte

through the ion exchange membranes without mixing of the electrolytes. In various embodiments system **400** is capable of forming bicarbonate ions in first electrolyte **102** when a voltage of 0.4 V or less, or 0.6 V or less, or 0.8 V or less is applied across the anode **112** and cathode **114**.

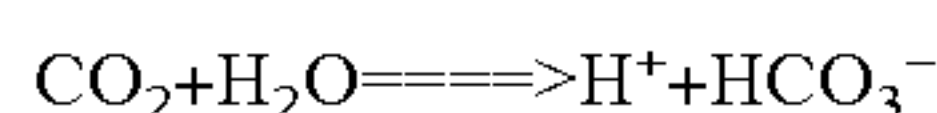
[0059] In the system of FIG. 4, first electrolyte **102**, second electrolyte **116** third electrolyte **118**, and fourth electrolyte **206** initially may comprise an aqueous salt solution such as a saltwater, e.g., sodium chloride, stannous chloride, seawater, brine, brackish water, conductive fresh water and the like. As indicated by the results achieved with one embodiment as set forth in Table 1, initially a 2 M solution of sodium chloride sodium chloride solution **128** was added to fourth compartment **140** to form fifth electrolyte **404**; in another embodiment, initially fifth electrolyte **404** comprised 0.5 M solution of sodium chloride. In specific embodiments, the fifth electrolyte **404** may be charged initially with a salt solution, e.g., sodium chloride, at a concentration from 0.1 to 4 M, e.g., 0.1 to 2.5 M, or 0.2 to 2.0 M, or 0.1 to 1.0 M, or 0.2 to 1.0 M, or 0.2 to 0.8 M, or 0.3 to 0.7 M, or 0.4 to 0.6 M, or 0.5 to 2.5 M, or 1.0 to 2.5 M, or 1.5 to 2.5 M, or 1.7 to 2.3 M.

[0060] Referring to FIG. 4, the voltage across anode **112** and cathode **114** can be regulated to form bicarbonate ions **122** in first electrolyte **102** without forming a gas, e.g., chlorine at anode **112** or hydrogen at cathode **114**. In various embodiments, bicarbonate ions **122** are formed in first electrolyte **102** where the voltage applied across anode **112** and cathode **114** is less than 2.8, 2.7, 2.6, 2.5, 2.4, 2.3, 2.2, 2.1, 2.0, 1.9, 1.8, 1.7, 1.6, 1.5, 1.4, 1.3, 1.2, 1.1, 1.0, 0.9, 0.8, 0.7, 0.6, 0.5, 0.4, 0.3, 0.2, or 0.1 V. Similarly, on applying a voltage across the cathode **114** and anode **112** an acid solution **124** is formed in third electrolyte **118** in contact with first cation exchange membrane **108A** as a result of protons transfer through first cation exchange membrane **108A** from first electrolyte **102**. The acid formed depends on the electrolytes used, e.g., as illustrated in FIG. 4, where first electrolyte **102** comprises sodium chloride, the acid formed in third electrolyte **118** comprises hydrochloric acid. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions and the acid solution are formed

when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0061] As is illustrated in FIG. 4, system 400 includes inlet ports 126 A-E adapted for introducing substances into cell 402, e.g., for introducing carbon dioxide 104, sodium chloride solution 128 and other electrolytes into cell 402; and outlet ports 130 A-E for removing substances from the cell, e.g., removing bicarbonate solution 122 and acid 124 from the cell. As will be appreciated by one ordinarily skilled in the art, the inlet and outlet ports are adaptable for various flow protocols including batch flow, semi-batch flow, or continuous flow. In various embodiments, the system includes voltage regulator 120 for regulating voltages across the anode and cathode and current through the electrolytes.

[0062] In system 400 of FIG. 4, as with the systems of FIGS. 1, 2 and 3, without being bound by any theory, it is believed that bicarbonate ions (HCO_3^-) are formed in first electrolyte 102 by carbon dioxide contacting water in the first electrolyte 102, as follows:



Thus, in accordance with the present invention and with reference to FIG. 4, where fifth electrolyte 404 comprise Na^+ and Cl^- ions from added sodium chloride 128, by placing first electrolyte 102 between cation exchange membrane 108A selective to transferring H^+ ions, and second cation exchange membrane 108B selective to transferring of cations, e.g., Na^+ ions, and on applying a voltage across the electrodes, H^+ will migrate through first cation exchange membrane 108A to adjacent third electrolyte 118. Similarly, Na^+ will migrate from fifth electrolyte 404 through second cation exchange membrane 108B to first electrolyte 102. Consequently, in first electrolyte 102, a solution comprising sodium bicarbonate 122 will form. Depending on the rate of introduction and/or removal of first electrolyte from the cell and the voltage applied across the electrodes 112, 114, the concentration of sodium bicarbonate 122 in first electrolyte 102 will adjust, e.g., increase, decrease or will not change.

[0063] Also with reference to FIG. 4, as H^+ migrate from first electrolyte 102 through first cation exchange membrane 108A to adjacent third electrolyte 118, the acidity of adjacent third electrolyte 118 will adjust depending on the rate of introduction and/or removal of hydrochloric acid 124 from the system. Similarly, as chloride ions migrate from fourth electrolyte 206 to adjacent third electrolyte 118 across second anion exchange membrane 108B, the chloride ion concentration in adjacent electrolytes 118 and 206 will adjust. Additionally, as chloride ions migrate from sodium chloride solution 128 in fifth electrolyte 404 to the second electrolyte 116 across first anion exchange membrane 106A, fifth electrolyte 404 will be depleted of chloride ions; consequently, fifth electrolyte will be depleted of sodium chloride, and correspondingly, the chloride ion content of the second electrolyte 116 will adjust, e.g., increase, decrease or remain constant depending on the flow of second electrolyte 116 from the system.

[0064] Hence, as illustrated in FIG. 4, in various embodiments a solution of bicarbonate ions 122, e.g., sodium bicarbonate is obtained in first electrolyte 102; an acid solution 124, e.g., hydrochloric acid is obtained in third electrolyte 118; a chloride solution, e.g., stannous chloride, is obtained in second electrolyte 116; fourth electrolyte 206 will be

depleted of chloride ions; and fifth electrolyte 404 initially comprising sodium chloride solution 128 will be depleted of sodium and chloride ions.

[0065] Optionally, as will be appreciated by one ordinarily skilled in the art, cations in electrolyte 116 in contact with anode 112 can be recovered by plating out the cations at the cathode 114, e.g., using electrolyte 116 from the anode as the electrolyte at the cathode. Thus, the anode material can be recovered at the cathode by switching electrolyte 116 in contact anode 112 with the electrolyte in contact with the cathode 114 when a sufficient concentration of Sn^{2+} has accumulated in the electrolyte 116, and allowing the cations to plate out at the cathode. Similarly, it will be appreciated that when sacrificial anode 112 is diminished and cathode 114 is augmented sufficiently, these electrodes may be switched so that anode 112 is transferred to replace cathode 114 and vice versa.

[0066] In an embodiment as illustrated in FIG. 5 and with reference to FIGS. 1-4, present method 500 comprises step 502 of applying a voltage 120 across an anode 112 and a cathode 114 through a first electrolyte 102 comprising carbon dioxide 104 to form bicarbonate ions 122 in the first electrolyte without forming a gas at the cathode or the anode. In accordance with the method and with reference to FIG. 1, first electrolyte 102 is contained between first anion exchange membrane 106A and first cation exchange membrane 108A in electrochemical cell 100; the anion exchange membrane contacts the anode 112 through second electrolyte 116; and the cation exchange membrane contacts cathode 112 through third electrolyte 118. In various embodiments, the method forms bicarbonate ions 122 in first electrolyte 102 when a voltage, e.g., 0.4 V or less, or 0.6 V or less, or 0.8 V or less is applied across the anode and cathode. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0067] As described with reference to the systems of FIGS. 1-4 above, method 500 forms a protonated solution in third electrolyte 118, e.g., hydrochloric acid 124 as a result of transfer of protons across first cation exchange membrane 108A from first electrolyte 102; and an ionic solution, e.g., stannous chloride in second electrolyte 116, as a result of chloride ions transferring across first anion exchange mem-

brane **106A** from first electrolyte **102**, and tin ions forming in second electrolyte **116** by oxidation of anode **112** comprising tin.

[0068] In various embodiments of method **500**, optionally, where anode **112** comprises a sacrificial anode, e.g., tin, copper, iron, zinc, cations such as Sn^{2+} will form in second electrolyte **116** in contact with anode **112**. Optionally, as described above with reference to FIGS. 1-4, cations in electrolyte **116** in contact with anode **112** can be recovered by plating out the cations at the cathode **114**, e.g., using electrolyte **116** from the anode as the electrolyte at the cathode. Thus, anode material can be recovered at the cathode **114** by switching electrolyte **116** in contact anode **112** with the electrolyte in contact with the cathode **114** when a sufficient concentration of Sn^{2+} has accumulated in the electrolyte **116**, and allowing the cations to plate out at the cathode.

[0069] In another embodiment and with reference to FIGS. 1-4, method **600** comprises step **602** of applying a voltage **120** of less than 2.0 V, less than 1.5 V, less than 1.0 V, less than 0.5 V, less than 0.1 V or less than 0.05 V across an anode **112** and a cathode **114** through first electrolyte **102** comprising carbon dioxide **104** to form bicarbonate ions **122** in the first electrolyte. In various embodiments the method forms bicarbonate ions in first electrolyte when a voltage, e.g., 0.4 V or less, or 0.6 V or less, or 0.8 V or less is applied across the anode and cathode. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0070] In accordance with method **600**, and with reference to FIG. 1, first electrolyte **102** is contained between first anion exchange membrane **106A** and first cation exchange membrane **108A** in electrochemical cell **302**; first anion exchange membrane **106A** contacts anode **112** through second electrolyte **116**; and first cation exchange membrane **108A** contacts the cathode through third electrolyte **118**. In another embodiment of method **600**, and with reference to FIG. 3, first electrolyte **102** is contained between first cation exchange membrane **108A** and second cation exchange membrane **108B** in electrochemical cell **303**; second cation exchange membrane **108B** contacts anode **112** through second electrolyte **116**; first cation exchange membrane **108A** separates first electrolyte

102 from third electrolyte **118**; second anion exchange membrane **106B** separates third electrolyte **118** from fourth electrolyte **206**; and fourth electrolyte **206** is in contact with cathode **114**. In various embodiments, method **600** forms bicarbonate ions **122** in first electrolyte when a voltage, e.g., 0.4 V or less, or 0.6 V or less, or 0.8 V or less is applied across the anode and cathode.

[0071] As disclosed with reference to the system of FIGS. 1-4 above, method **600** forms an acid, e.g., hydrochloric acid **124**, depending on the electrolytes used. As is illustrated in FIG. 1, the method forms a protonated solution in third electrolyte **118**, e.g., hydrochloric acid as a result of transfer of protons across first cation exchange membrane **108A** from first electrolyte **102**; and an ionic solution, e.g., stannous chloride in second electrolyte **116**, as a result of chlorine ions transferring across first anion exchange membrane **106A** from first electrolyte **102**, and tin ions forming by oxidation of the anode **112**.

[0072] In various embodiments of method **600**, optionally, where anode **112** comprises a sacrificial anode, e.g., tin, copper, iron, zinc, cations such as Sn^{2+} will form in second electrolyte **116** in contact with anode **112**. Optionally, as described above with reference to FIGS. 1-4, cations in electrolyte **116** in contact with anode **112** can be recovered by plating out the cations at the cathode **114**, e.g., using electrolyte **116** from the anode as the electrolyte at the cathode. Thus, anode material can be recovered at the cathode **114** by switching electrolyte **116** in contact anode **112** with the electrolyte in contact with the cathode **114** when a sufficient concentration of Sn^{2+} has accumulated in the electrolyte **116**, and allowing the cations to plate out at the cathode. Similarly, it will be appreciated that when sacrificial anode **112** is diminished and cathode **114** is augmented sufficiently, these electrodes may be switched so that anode **112** is transferred to replace cathode **114** and vice versa.

[0073] Exemplary results achieved in one embodiment of the present system and method are set forth in Table 1.

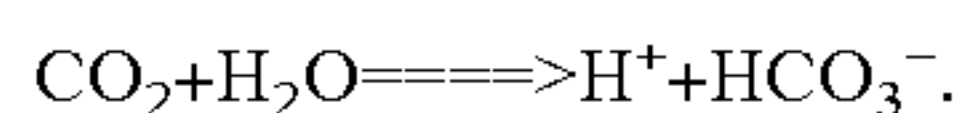
TABLE 1

LOW ENERGY ELECTROCHEMICAL BICARBONATE ION SOLUTIONS			
Voltage across Anode and Cathode (V)		pH of electrolyte solution in Compartment 136	pH of Bicarbonate ion solution in Compartment 132
0.4	Initial	6.163	4.229
	pH		
	Final	4.367	5.950
0.6	pH		
	Initial	5.846	4.447
	pH		
0.8	Final	4.408	5.824
	pH		
	Initial	8.502	4.306
	pH		
	Final	4.353	6.642
	pH		

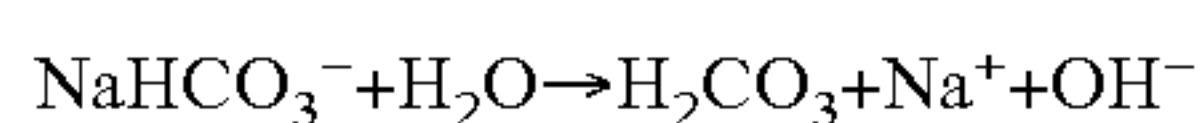
[0074] In this example, based on system **100** of FIG. 1 and method **500** of FIG. 5 and method **600** of FIG. 6, first electrolyte **102**, contained in compartment **132**, was charged with a 2 M sodium chloride solution **128** to which carbon dioxide gas **104** was added. Third electrolyte **118** comprising saltwater, e.g., stannous chloride was contained in compartment **136**. First anion exchange membrane **106A** separated first electrolyte **102** from second electrolyte **116**; first cation

exchange membrane **108A** separated first electrolyte **102** from third electrolyte **118**; anode **112** formed of tin foil were placed in contact with second electrolyte **116**, and cathode **114** formed of tin foil was placed in contact with third electrolyte **118**. Voltages of 0.4 V, 0.6 V and 0.8 V were applied across anode **112** and cathode **114** in a batch mode operation for one hour. As set forth in Table 1, the pH of first electrolyte **102** in compartment **132** increased (correlating to an increase of hydroxide ion concentration in first electrolyte **102** as described above), while the pH of third electrolyte **118** in compartment **136** decreased (correlating to an increase in protons in third electrolyte **118** as described above), without the formation of a gas, e.g., chlorine at anode **112** or hydrogen at cathode **114**.

[0075] As discussed above with reference to FIGS. 1-4, without being bound by any theory it is believed that in first electrolyte **102** bicarbonate ions formed as a result of carbon dioxide contacting water in first electrolyte **102**, as follows:



In first electrolyte **102**, Na^+ and Cl^- ions are present from the sodium chloride **128**. Thus, in accordance with the present invention, by placing first electrolyte **102** between first cation exchange membrane **108A** selective to transferring H^+ ions, and first anion exchange membrane **106A** selective to transferring of Cl^- ions, and on applying a voltage across the electrodes, H^+ migrated through cation exchange membrane **108A** to adjacent third electrolyte **118**. Similarly, Cl^- migrated from first electrolyte **102** through anion **106A** exchange membrane to adjacent second electrolyte **116**. Consequently, a solution comprising sodium bicarbonate **122** formed in first electrolyte **102**. With the formation of sodium bicarbonate in first electrolyte **102** the pH of the first electrolyte increased in accordance with the following reaction:



As discussed with reference to FIGS. 1-4, as H^+ migrated from the first electrolyte **102** through first cation exchange membrane **108A** to adjacent third electrolyte **118** in third compartment **136**, the acidity of adjacent third electrolyte **118** increased as indicated by the decrease in pH in third compartment **136** as set forth in Table 1.

[0076] As will be appreciated by one ordinary skilled in the art, the voltages may be adjusted up or down from these exemplary voltages; a minimum theoretical voltages 0 V or very close to 0 V, but to achieve a useful rate of production of bicarbonate ions, a practical lower limit may be in some embodiments 0.001 V or 0.01 V, or 0.1 V, depending on the desired time for bicarbonate ion production and/or pH adjustment, volume of first electrolyte solution **102**, and other factors apparent to those of ordinary skill; e.g., in some embodiments system **100**, system **200**, system **300** and system **400** and method **500** and method **600** are capable of producing bicarbonate ions at voltages as low as 0.001 V, or 0.01 V, or 0.1V, and can also produce bicarbonate ions at higher voltages if more rapid production is desired, e.g., at 0.2-2.0 V; in various embodiments the bicarbonate ions are produced with no gas formation at the anode or cathode, e.g., no formation of hydrogen or chlorine at the electrodes.

[0077] In these examples, and in various embodiments of the invention, a pH difference of more than 0.5, 1.0, 1.5, 2.0, 2.5, 3.0, 3.5, 4.0, 4.5, 5.0, 5.5, 6, 6.5, 7.0, 7.5, 8.0, 8.5, 9.0, 9.5, 10.0, 10.5, 11.0, 11.5, or 12.0 pH units may be produced in a first electrolyte solution **102** and in third electrolyte solution **118** when a voltage of 1.0 V or less, or 0.9 V or less, or 0.8 V

or less, or 0.7 V or less, or 0.6V or less, or 0.5 V or less, or 0.4 V or less, or 0.3 V or less, or 0.2 V or less, or 0.1 V or less, or 0.05 V or less is applied across the anode and cathode. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.8 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.7 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.6 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.5 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.4 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.3 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.2 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.1 V without forming a gas at the electrodes. In certain embodiments, bicarbonate ions are formed when the voltage applied across the anode and cathode is less than 0.05 V without forming a gas at the electrodes.

[0078] As will be appreciated, in particular embodiments, the present invention provides a system and method capable of producing a pH difference of more than 0.5 pH units in first electrolyte **102** and third electrolyte **118** when a voltage of 0.05 V or less is applied across the anode and cathode. In some embodiments, the invention provides a system and method capable of producing a pH difference of more than 1.0 pH units between first electrolyte **102** and third electrolyte **118** when a voltage of 0.1V or less is applied across the anode and cathode. In some embodiments, the invention provides a system and method capable of producing a pH difference of more than 2.0 pH units between a first electrolyte and third electrolyte when a voltage of 0.2 V or less is applied across the anode and cathode.

[0079] In some embodiments, the invention provides a system and method capable of producing bicarbonate ions in first electrolyte **102** when a voltage of 0.4V or less is applied across the anode and cathode. In some embodiments, the invention provides a system and method capable of producing bicarbonates ions **122** when a voltage of 0.6V or less is applied across the anode and cathode. In some embodiments, the invention provides a system and method capable of producing bicarbonate ions **122** when a voltage of 0.8V or less is applied across the anode and cathode. In particular embodiments, the invention provides a system capable of producing bicarbonate ions **122** when a voltage of 1.0 V or less is applied across the anode and cathode. In some embodiments the invention provides a system capable of producing bicarbonate ions **122** in first electrolyte **102** when a voltage of 1.2 V or less is applied across the anode and cathode.

[0080] It will be appreciated that the voltage need not be kept constant and that the voltage applied across the anode and the cathode may be very low, e.g., 0.05V or less and that the voltage may be increased as needed as the concentration of bicarbonate ions in the solution **102** increases. In this manner, a desired bicarbonate ion concentration may be

achieved with the minimum average voltage, without generating a gas at the electrodes. Thus, in some embodiments as described in the previous paragraph, the average voltage may be less than 80%, 70%, 60%, or less than 50% of the voltages noted in the previous paragraph for particular embodiments.

[0081] In some embodiments, one or more of the initial electrolytes charged into the system may be depleted of divalent cations, e.g., the electrolytes are depleted of magnesium or calcium ion as for example where the electrolytes are taken form an ion exchange process. Thus, in some embodiments the total concentration of divalent cations in the electrolyte solutions in contact with the ion exchange membrane or membranes is less than 0.06 mol/kg solution, or less than 0.05 mol/kg solution, or less than 0.04 mol/kg solution, or less than 0.02 mol/kg solution, or less than 0.01 mol/kg solution, or less than 0.005 mol/kg solution, or less than 0.001 mol/kg solution, or less than 0.0005 mol/kg solution, or less than 0.0001 mol/kg solution, or less than 0.00005 mol/kg solution.

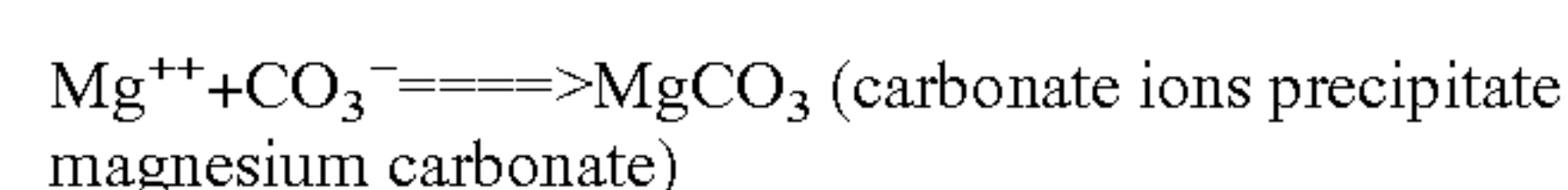
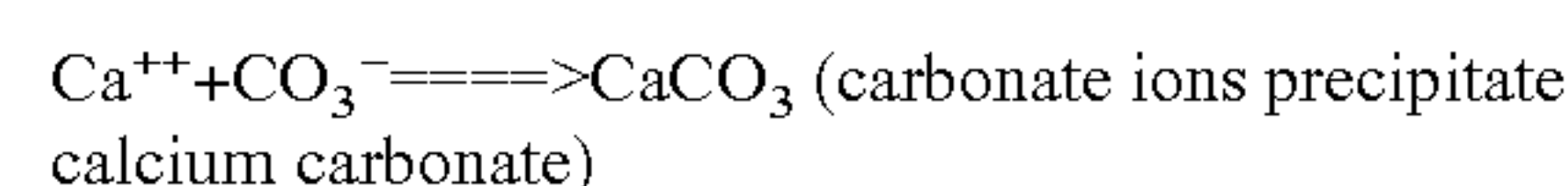
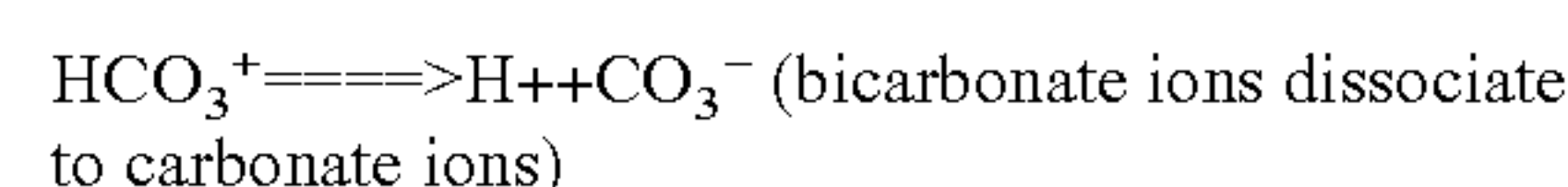
[0082] As discussed in various embodiments herein, the carbon dioxide that contacts first electrolyte **102** may initially form bicarbonate ions **122** in the first electrolyte. As bicarbonate ions are removed from first electrolyte **102** more carbon dioxide may dissolve in the electrolyte to form bicarbonate and/or carbonate ions. Depending on the pH of the first electrolyte, the balance is shifted toward bicarbonate or toward carbonate formation, as is well understood in the art. In these embodiments the pH of the first electrolyte may decrease, remain the same, or increase, depending on the rate of removal of bicarbonate and/or carbonate ions compared to rate of introduction of carbon dioxide. It will be appreciated that no bicarbonate ions need form in these embodiments, or that bicarbonate ions may not form during one period but form during another period.

[0083] Optionally, the present system is used to produce bicarbonate ions **122**, which, when included in a solution comprising alkaline earth cations and hydroxide ions causes precipitation of carbonate and/or bicarbonate compounds such as calcium carbonate or magnesium carbonate and/or their bicarbonates. In various embodiments, divalent cations such as magnesium and/or calcium are present in the solutions used in the process, and/or are added. The precipitated carbonate compound can be used as cements and other building and construction material such as aggregates and the like as described in U.S. patent application Ser. No. 12/126,776, filed on May 23, 2008, incorporated herein by reference.

[0084] In an optional step, the acidified electrolyte solution **118** illustrated in FIGS. 1-4 is utilized to dissolve a calcium and/or magnesium rich mineral, such as mafic mineral including serpentine or olivine, to form a solution for precipitating carbonates and bicarbonates as described in the United States patent applications incorporated herein by reference. For example, acidified stream **118** can be used to dissolve calcium and/or magnesium rich minerals such as serpentine and olivine to from an electrolyte solution that can be charged with bicarbonate ions **122** and then made sufficiently basic to precipitate carbonate compounds. Such precipitation reactions and the use of the precipitates, e.g., as in cements are described in the U.S. patent application Ser. No. 12/126,776, filed on May 23, 2008 and incorporated herein by reference.

[0085] In an other optional embodiment, the bicarbonate ion solutions of the present invention can be utilized to desalinate saltwater by removing divalent cations as insoluble carbonates, e.g., removing calcium and magnesium ions from a saltwater e.g., seawater based on the following reactions and

as described in U.S. patent application Ser. No. 12/163,205, filed on Jun. 27, 2008, herein incorporated by reference:



[0086] Although the foregoing invention has been described in some detail by way of illustration and example for purposes of clarity of understanding, it will be readily apparent to those of ordinary skill in the art in light of the teachings of this invention that certain changes and modifications may be made thereto without departing from the spirit or scope of the appended claims.

[0087] Accordingly, the preceding merely illustrates the principles of the invention. It will be appreciated that those skilled in the art will be able to devise various arrangements, which, although not explicitly described or shown herein, embody the principles of the invention, and are included within its spirit and scope. Also, examples and conditional language recited herein are principally intended to aid the reader in understanding the principles of the invention and the concepts contributed by the inventors to furthering the art, and are to be construed as being without limitation to such specifically recited examples and conditions. Moreover, all statements herein reciting principles, aspects, and embodiments of the invention as well as specific examples thereof, are intended to encompass both structural and functional equivalents thereof. Additionally, it is intended that such equivalents include both currently known equivalents and equivalents developed in the future, i.e., any elements developed that perform the same function, regardless of structure. The scope of the present invention, therefore, is not intended to be limited to the exemplary embodiments shown and described herein. Rather, the scope and spirit of present invention is embodied by the appended claims.

1. An electrochemical system comprising:
an anode, a cathode and a first electrolyte disposed between the anode and cathode, wherein the system is configured to form bicarbonate ions in the first electrolyte without forming a gas at the cathode or anode on applying a voltage across the anode and cathode and contacting the first electrolyte with carbon dioxide.
2. (canceled)
3. The system of claim 1 comprising:
a first anion exchange membrane and a cation exchange membrane between which is contained the first electrolyte.
4. (canceled)
5. The system of claim 3, further comprising:
a second electrolyte contacting the first anion exchange membrane and the anode;
a third electrolyte contained between the cation exchange membrane and a second anion exchange membrane; and
a fourth electrolyte contacting the second anion exchange membrane and the cathode.
6. (canceled)
7. (canceled)
8. The system of claim 5, further comprising:
a fifth electrolyte contained between the second cation exchange membrane and a second anion exchange

membrane, and wherein the second electrolyte contacts the second anion exchange membrane and the anode.

9. (canceled)

10. (canceled)

11. (canceled)

12. (canceled)

13. The system of claim 8, wherein the first electrolyte comprises sodium chloride and carbon dioxide.

14. (canceled)

15. (canceled)

16. (canceled)

17. (canceled)

18. (canceled)

19. (canceled)

20. (canceled)

21. (canceled)

22. (canceled)

23. The system of claim 1, configured to form sodium bicarbonate in the first electrolyte.

24. The system of claim 13, configured to form hydrochloric acid in the third electrolyte.

25. The system of claim 24, wherein the anode and cathode are selected from tin, nickel, cobalt or copper.

26. The system of claim 25, configured to oxidize the anode to tin ions in the second electrolyte and reducing tin ions to tin at the cathode.

27. (canceled)

28. The system of claim 26, configured to form sodium bicarbonate in the first electrolyte.

29. (canceled)

30. (canceled)

31. The system of claim 28, configured to form tin ions in the electrolyte at the anode and reducing tin ions to tin from the electrolyte at the cathode.

32. (canceled)

33. (canceled)

34. (canceled)

35. (canceled)

36. (canceled)

37. (canceled)

38. (canceled)

39. (canceled)

40. (canceled)

41. An electrochemical method comprising:
applying a voltage across an anode and a cathode through a first electrolyte comprising added carbon dioxide to form bicarbonate ions in the first electrolyte without forming a gas at the cathode or the anode, wherein the first electrolyte is located between the anode and cathode.

42. (canceled)

43. The method of claim 41, wherein the first electrolyte is contained between an anion exchange membrane and a cation exchange membrane.

44. The method of claim 43, wherein the anion exchange membrane contacts the anode through the second electrolyte; and
the cation exchange membrane contacts the cathode through a third electrolyte.

45. (canceled)

46. The method of claim 44, wherein
the first cation exchange membrane separates the first electrolyte from the third electrolyte;

the second cation exchange membrane contacts the anode through the second electrolyte;
a first anion exchange membrane separates the third electrolyte from the fourth electrolyte; and
the fourth electrolyte contacts the cathode.

47. The method of claim 44, wherein
the first cation exchange membrane separates the first electrolyte from the third electrolyte;
the first anion exchange membrane separates the third electrolyte from the fourth electrolyte;
the fourth electrolyte contacts the cathode;
a fifth electrolyte is contained between the second cation exchange membrane and a second anion exchange membrane; and
the second anion exchange membrane contacts the anode through the second electrolyte.

48. The method of claim 44, wherein the first electrolyte comprises sodium chloride and carbon dioxide.

49. The method of claim 44, further comprising transferring anions across the anion exchange membrane from the first electrolyte to the second electrolyte.

50. (canceled)

51. The method of claim 44, further comprising transferring cations across the cation exchange membrane from the first electrolyte to the third electrolyte.

52. The method of claim 44, further comprising transferring protons across the first cation ion exchange membrane from the first electrolyte to the third electrolyte.

53. (canceled)

54. (canceled)

55. The method of claim 44, comprising precipitating alkaline metal carbonates utilizing the first electrolyte.

56. (canceled)

57. The method of claim 44, comprising forming sodium bicarbonate in the first electrolyte.

58. The method of claim 44, further comprising separating the cathode from the third electrolyte utilizing second anion exchange membrane whereby the cathode is electrically connected to second anion exchange membrane through fourth electrolyte.

59. The method of claim 58, further comprising transferring chloride ions across the second anion exchange membrane from the fourth electrolyte to the third electrolyte.

60. (canceled)

61. The method of claim 59, comprising forming hydrochloric acid in the third electrolyte.

62. (canceled)

63. (canceled)

64. (canceled)

65. (canceled)

66. The method of claim 44 comprising forming hydrochloric acid in the third electrolyte.

67. (canceled)

68. (canceled)

69. (canceled)

70. (canceled)

71. (canceled)

72. (canceled)

73. (canceled)

74. (canceled)

75. (canceled)